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Fully quantitative description of hybrid TiO₂ nanoparticles by means of solid state ³¹P NMR

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ABSTRACT: For the first time, an absolute quantification of hybrid materials obtained from the reaction of phenylphosphonic acid (PPA) with TiO₂ nanoparticles under different reaction conditions is reported. Next to the amount of PPA involved in grafting to the TiO₂ nanoparticles, also the PPA included in titaniumphenylphosphonate crystallites is described quantitatively. The quantitative analysis is based on solid state ³¹P MAS NMR and is further applied to evaluate the stability of the resulting hybrid materials towards hydrolysis and organic solvent exposure.

INTRODUCTION

In the past 30 years, the development of organic-inorganic hybrid materials attracted the attention of many scientists^{1,2,3,4} since the combination of inorganic and organic properties revealed to be functional in many applications⁵ such as coatings⁶, gas storage⁷, separation processes^{8,9}, catalysis¹⁰ and optical device production¹¹. In this framework, the scientific community focused its efforts on the development of synthesis pathways dedicated to form stable bonds between the organic and the inorganic species^{12,13,14}. Surface grafting of inorganic substrates by incorporating organic monolayers revealed to be one of the most studied pathways for producing hybrid materials^{14,15,16}. Many examples of such materials, composed of metal oxide surfaces modified by organic coupling molecules such as organosilanes or organophosphorous compounds are reported in literature^{12,17,18,19,20}. Surface grafting of

metal oxide supports anchoring organosilanes (i.e. organoalkoxysilanes, organohalosilanes or organoaminosilanes) can only by formation of M-O-Si-R bonds (M = transition metal)^{21,22}. However, a limited stability of such organic layers is observed, due to the vulnerability of M-O-Si-R bonds towards hydrolysis²³. Also phosphonic acids have been used to introduce new properties at the surface of metal oxides via heterocondensation between the metal oxide hydroxyl groups and the P-OH groups of the phosphonic acid^{14,24,25,26}. This reaction results in more stable M-O-P-R bonds²³. A stable layer can be formed by the reaction of TiO₂ with phosphonic acid in monodentate or bidentate binding modes¹⁹. The phosphonic acid grafting methodology has been widely applied to TiO₂ nanoparticles²⁷ in order to produce multifunctional nanomaterials^{28,29}. Guerrero et al. reported that the formation of a stable organic layer anchored at the TiO₂ surface can be obtained at room temperature^{19,24}. It was also reported that the surface modification of TiO₂ at higher reaction temperatures also includes the formation of titaniumphosphonate resulting from the partial hydrolysis of the TiO₂ matrix^{19,24}. ATR-FTIR, XRD, TEM and solid state ³¹P, ¹³C and ¹⁷O NMR spectroscopy have been used to characterize the obtained titanium dioxide modified materials^{19,24,30}. Solid state NMR in general is a strong technique for the structural characterization of materials^{31,32,33}. In particular, ³¹P solid state NMR revealed to be very important to differentiate between phosphonic acid involved in the surface modification of TiO₂ nanoparticles and the phosphonic acid involved in the formation of titaniumphosphonate since, depending on the reaction temperature, both species can be present together in the modified material³⁰. From literature, it is clear that numerous papers deal with the qualitative characterization of phosphonic acid modified metal oxides^{14,15,16,26}. However, very little attention has been paid to obtain quantitative data regarding the amount of PPA grafted to TiO₂ nanoparticles and the amount included in titaniumphosphonate. This paper introduces for the first time a method that allows a fully and absolute quantification of phosphonic acid modified TiO₂ nanoparticles, including the amount of PPA involved in TiO₂ surface grafting next to the amount involved in titaniumphenylphosphonate structures. Moreover, the stability in strong acid conditions and organic solvents was evaluated quantitatively. Last

but not least, the quantitative description of these modified TiO₂ materials by solid state ³¹P-NMR might pave the way towards the design and production of property-tuned TiO₂ hybrid materials.

EXPERIMENTAL

Hydrophilic fumed P25 TiO₂ powder (Aeroxide), with a specific surface area of 50 m²/g and an average primary particle size of 21 nm was used as substrate material. Phenylphosphonic acid (PhPO₃H₂, PPA, 98%, 158.09 MW) and octylphosphonic acid (C₈PO₃H₁₉, OPA, 97%, 194.21 MW) were obtained from Aldrich, analyzed for structure and purity by solid state magic angle spinning (MAS) ³¹P-NMR and used without further purification. MilliQ water was used for all solutions.

TiO₂/PPA-20°C: 1 g (8.7 mmol) of TiO₂ was dispersed in 20 ml of a 0.1 M solution of PPA in water (316 mg/20 ml, 2 mmol). After stirring the reaction mixture for 24 h at 20°C, the solid was recovered by membrane filtration (VitraPOR Borosilicate 3.3) and washed with 200 ml water in order to remove remaining free PPA. Finally the material was dried at 120°C under vacuum for 24h.

TiO₂/PPA-45°C, TiO₂/PPA-90°C, TiO₂/PPA-120°C and TiO₂/PPA-150°C: 750 mg (6.5 mmol) of TiO₂ was dispersed in 15 ml of a 0.1 M solution of PPA in water (237 mg/15ml, 1.5 mmol). The reaction was carried out for 3 hours in a microwave reactor. The solid was recovered and dried as described above.

TiO₂/PPA-150°C-HCl-treated: 1g of TiO₂/PPA-150°C was dispersed in 15 ml HCl 1M at room temperature under stirring. After 12 hours the material was recovered by membrane filtration (VitraPOR Borosilicate 3.3) and washed with 800 ml of water. Finally the material was dried for 24 hours under vacuum.

TiO₂/OPA: 6 mg of OPA was dissolved in 1.5 ml acetone and dispersed with 114 mg of TiO₂. Subsequently, the acetone was evaporated at 60°C and the sample was analyzed by means of solid state ³¹P MAS NMR.

TiO₂/PPA: 10 mg of PPA was dissolved in 2 ml acetone and dispersed with 100 mg of TiO₂. Subsequently, the acetone was evaporated at 60°C and the sample was analyzed by means of solid state ³¹P MAS NMR.

Preparation of TiO₂/PPA-OPA samples for absolute quantification. In order to absolutely quantify the TiO₂/PPA samples, octylphosphonic acid was added as internal standard. The samples are coded TiO₂/PPA-x°C-OPA, with x°C the reaction temperature at which the TiO₂ was modified with PPA before. **TiO₂/PPA-20°C-OPA:** 6 mg of OPA was dissolved in 1.5 ml acetone and dispersed with 114 mg of TiO₂/PPA-20°C. Subsequently, the acetone was evaporated at 60°C, after which a known amount (in the order of 50 mg) was placed in a ceramic rotor for solid state MAS NMR analysis. The exact amount of material in the rotor was obtained from the weight difference between an empty and a filled rotor. **TiO₂/PPA-45°C-OPA** was prepared in the same way and 48 mg was put in the rotor for SS-NMR. **TiO₂/PPA-90°C-OPA:** 10 mg of OPA was dissolved in 1.5 ml of acetone and dispersed with 110 mg of TiO₂/PPA-90°C. Subsequently, the acetone was dried for 5 minutes at 60°C and 45 mg was used for SS-NMR. The samples **TiO₂/PPA-120°C-OPA, TiO₂/PPA-150°C-OPA, TiO₂/PPA-150°C-HCl-treated-OPA and TiO₂/PPA-150°C-post-SPE-OPA** were prepared in the same way as described for TiO₂/PPA-90°C-OPA.

Solid phase extraction (SPE) with TiO₂/PPA-150°C as stationary phase: 1.8 g of TiO₂/PPA-150°C material was packed in a solid phase extraction column and solid phase extraction was carried out following the CLWE procedure³⁴. Conditioning of the material was carried out with 2 ml of water followed by loading a mixture composed of 1 ml methanol and 1 ml toluene. After recovery of the flow-

through, the solid phase was washed with 2 ml of water, followed by a final elution with 4 ml of cyclohexane. All operations were accomplished in a glass box (see Figure S1) under vacuum at a fixed pressure of 15 psi. After SPE, the stationary phase TiO₂/PPA-150°C material was recovered and washed several times with acetone and water, followed by drying at 120°C under vacuum for 24h. Regarding the eluted fractions, all were analyzed by liquid state ¹H NMR spectroscopy. ¹H liquid state NMR spectra were recorded at room temperature on a Varian Inova 300 spectrometer in a 5mm four-nucleus PFG probe. The chemical shift (δ) scale in ppm was calibrated relative to TMS (0 ppm). Other NMR parameters were: a 90° pulse of 4.0 μs, a spectral width of 5 kHz, an acquisition time of 4 s, a preparation delay of 12 s and 16 accumulations. A line-broadening factor of 0.2 Hz was applied before Fourier transformation to the frequency domain.

Solid state ³¹P MAS NMR spectra were acquired at ambient temperature on an Agilent VNMRS DirectDrive 400MHz spectrometer (9.4 T wide bore magnet) equipped with a T3HX 3.2 mm probe dedicated for small sample volumes and high decoupling powers. Magic angle spinning (MAS) was performed at 10 kHz using ceramic zirconia rotors of 3.2 mm in diameter (22 μl rotors). The phosphorus chemical shift scale was calibrated to orthophosphoric acid (H₃PO₄) at 0 ppm. Other acquisition parameters used were: a spectral width of 50 kHz, a 90° pulse length of 3.6 μs, an acquisition time of 20 ms and 200-3000 accumulations. High power proton dipolar decoupling was set to 100 kHz during the acquisition time. Since the longest spin-lattice relaxation time (T₁) registered was 250 s (for PPA physically adsorbed at the TiO₂ surface), a recycle delay of 1250 s was used in order to satisfy the condition to obtain fully quantitative results, i.e. using a recycle delay of 5x the longest T₁ decay time.

For completeness, the T₁ values measured for TiO₂/PPA prepared at 20°C, 45°C, 90°C, 120°C and 150°C are reported in Table 1. The T₁ decay time of TiO₂/OPA was 106 s.

Table 1: T1 values of the TiO₂/PPA materials.

Sample	T1 (s) of surface grafted PPA (23-2 ppm region)	T1 (s) of PPA in titaniumphosphonate (-4 ppm signal)
TiO ₂ /PPA-20°C	11	/
TiO ₂ /PPA-45°C	13	*ND
TiO ₂ /PPA-90°C	9	54
TiO ₂ /PPA-120°C	11	114
TiO ₂ /PPA-150°C	11	226

*ND: not determined as the signal is too small.

RESULTS AND DISCUSSION

Evaluation of OPA as internal standard for quantification. After reacting TiO₂ nanoparticles with phenylphosphonic acid (PPA) at different reaction temperatures, the PPA on the TiO₂ was characterized and quantified using solid state ³¹P MAS NMR. For the quantification, octylphosphonic acid (OPA) was used as an internal standard. Hereto, OPA was mixed with the respective TiO₂/PPA-x°C powder in acetone (with x°C the reaction temperature at which the TiO₂ was modified with PPA), resulting in the final TiO₂/PPA-x°C-OPA powders after drying. Figure 1a shows the ³¹P MAS spectrum of a dried TiO₂/OPA dispersion to demonstrate that the OPA signals do not interfere with the signals of free PPA at 21 ppm³⁰ (figure not shown) and the signal of a dried TiO₂/PPA dispersion (Figure 1b) which also shows a signal at 21 ppm. For all TiO₂/PPA species, the most downfield ³¹P chemical shift is situated at 21 ppm, and so more upfield as compared to the OPA signals³⁰. Remark that the signals of the dried TiO₂/OPA dispersion cover a region between 42 and 23 ppm due to the formation of covalent P-O-Ti bonds between OPA and TiO₂ in the dispersion. This is in agreement with literature, assigning the 42-23 ppm region to signals of OPA bonded to TiO₂¹⁴.

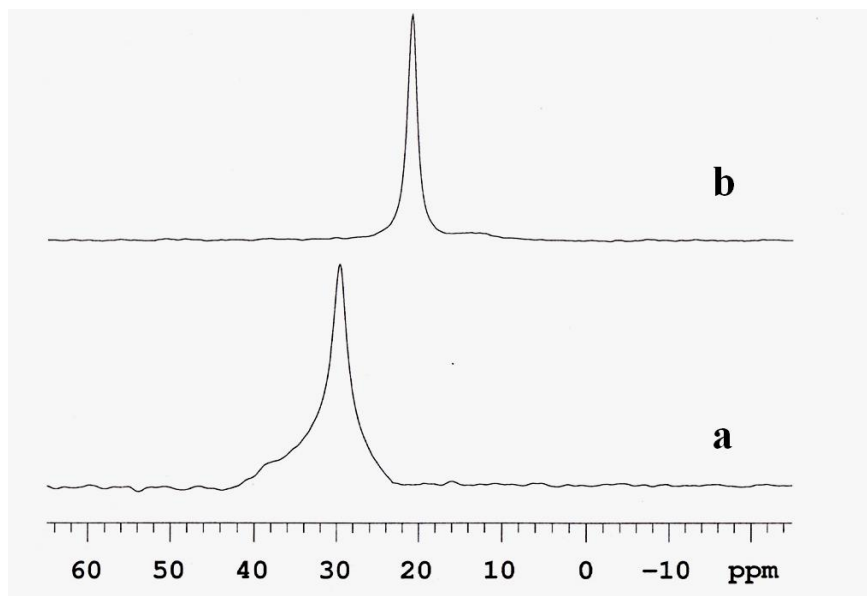


Figure 1: Solid state ^{31}P MAS NMR spectrum of a (a) dried dispersion of TiO_2/OPA and (b) dried dispersion of TiO_2/PPA .

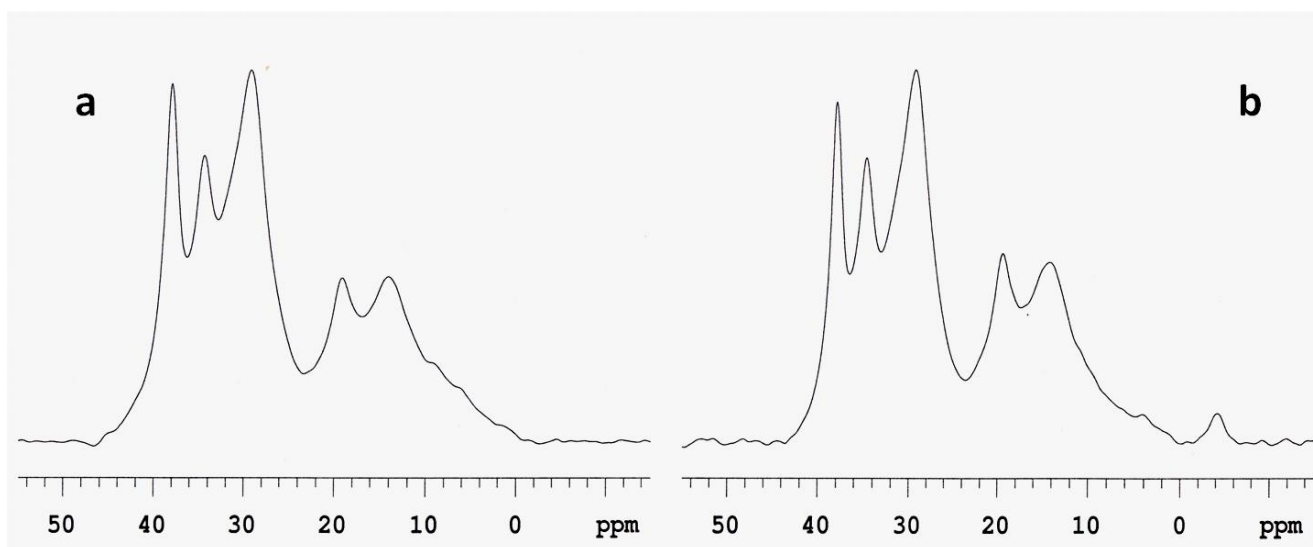


Figure 2: Solid state ^{31}P MAS NMR spectrum of (a) $\text{TiO}_2/\text{PPA-20}^\circ\text{C-OPA}$ and (b) $\text{TiO}_2/\text{PPA-45}^\circ\text{C-OPA}$.

Figures 2-4 show ^{31}P MAS NMR spectra of different $\text{TiO}_2/\text{PPA-x}^\circ\text{C-OPA}$ materials. Figure 2 illustrates the ^{31}P MAS spectra of $\text{TiO}_2/\text{PPA-20}^\circ\text{C-OPA}$ (Figure 2a) and $\text{TiO}_2/\text{PPA-45}^\circ\text{C-OPA}$ (Figure 2b). Both spectra display three signals around 38, 34 and 28 ppm arising from OPA (42 ppm to 23 ppm region). Whereas the resonance at 38 ppm arises from OPA physically adsorbed at the TiO_2 surface, the

resonances at 34 and 29 ppm arise from OPA which is covalently bonded to the metal oxide via P-O-Ti bonds¹⁴. The presence of covalently bonded OPA at the TiO₂ surface indicates that there are still Ti-OH groups available after the reaction of PPA with TiO₂ at these temperatures. The PPA region (23-2 ppm) of TiO₂/PPA-20°C-OPA and TiO₂/PPA-45°C-OPA shows two bands at around 19 ppm and 14 ppm, arising from PPA grafted at the surface of the TiO₂ nanoparticles. The ³¹P MAS spectrum of TiO₂/PPA-45°C-OPA shows a weak signal at -4 ppm in addition, indicating the presence of a minor amount of titaniumphenylphosphonate^{26,35}.

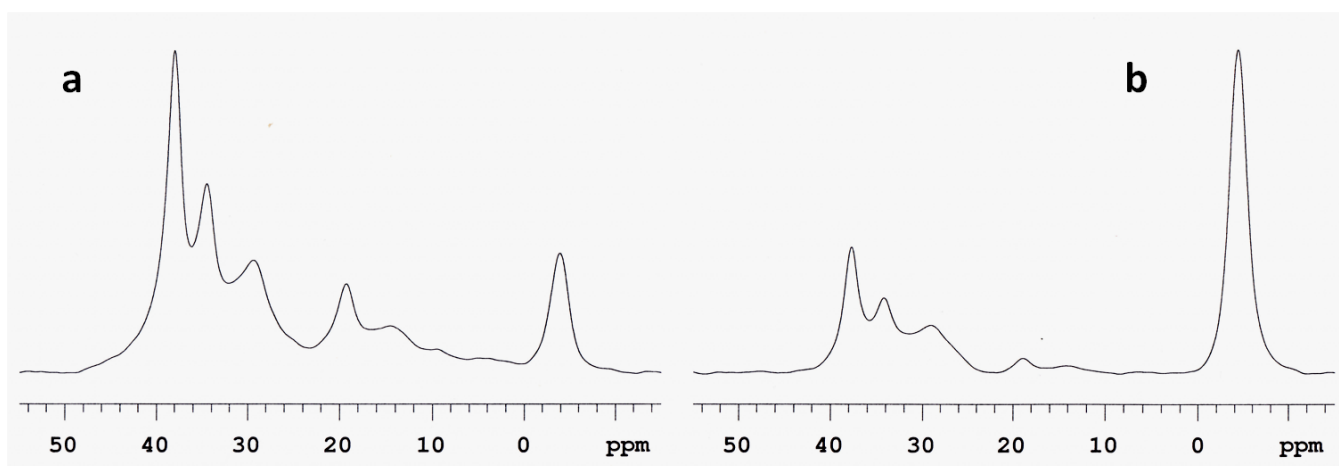


Figure 3: Solid state ³¹P MAS NMR spectrum of (a) TiO₂/PPA-90°C-OPA and (b) TiO₂/PPA-120°C-OPA.

Figure 3 shows the ³¹P MAS NMR spectra of TiO₂/PPA-90°C-OPA (Figure 3a) and TiO₂/PPA-120°C-OPA (Figure 3b). Remark that a somewhat larger amount of OPA had to be used since the amount of titaniumphenylphosphonate formed (signal at -4 ppm) starts to increase strongly as a function of the reaction temperature. This explains the relative change in the intensities of the three OPA signals in the region between 42 ppm and 23 ppm. In particular, the most significant difference is related to the increased intensity of the 38 ppm signal of OPA physically adsorbed at the TiO₂ surface. This can easily be explained by a larger amount of OPA standard that is physically adsorbed at the TiO₂ surface. The spectrum of TiO₂/PPA-90°C-OPA (Figure 3a) further presents the 19 and 14 ppm signals from PPA grafted at the TiO₂ nanoparticles, but compared to the spectra of TiO₂/PPA-20°C-OPA and TiO₂/PPA-45°C-OPA (Figure 2), the intensity ratio of the 19 ppm/14 ppm signals is increased. This results from a

changing distribution of the different PPA-TiO₂ binding modes. Furthermore, the intensity of the -4 ppm resonance, assigned to titaniumphenylphosphonate, is increased noticeably. The ³¹P MAS spectrum of TiO₂/PPA-120°C-OPA (Figure 3b) shows a severe decrease of the 19 and 14 ppm signals, relatively to the strong increase of the -4 ppm signal. This indicates that the formation of titaniumphosphonate becomes predominant as compared to PPA grafting at the TiO₂ nanoparticles.

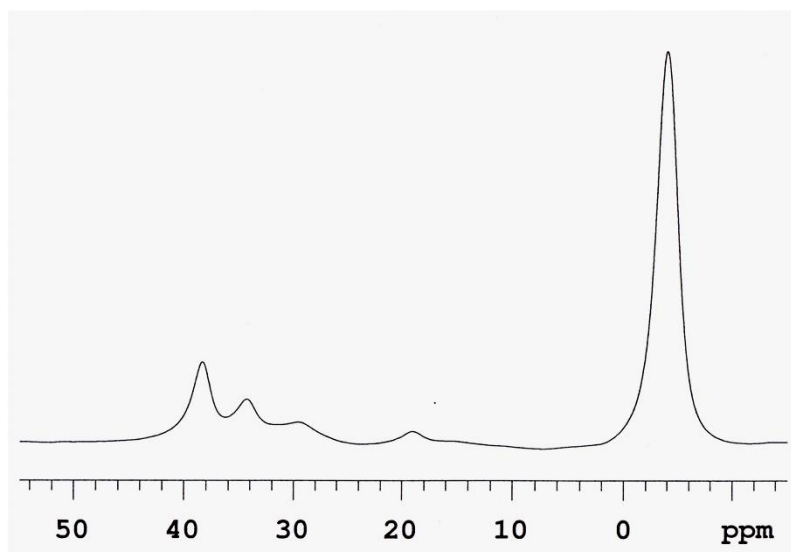


Figure 4: Solid state ³¹P MAS NMR spectrum of TiO₂/PPA-150°C-OPA.

This trend is even more enhanced for TiO₂/PPA-150°C-OPA as shown in the ³¹P MAS NMR spectrum (Figure 4). The spectrum is clearly dominated by the signal that represents titaniumphenylphosphonate at -4 ppm. Relatively to the titaniumphenylphosphonate signal, the signals around 19 and 14 ppm, characteristic for PPA grafted at the TiO₂ nanoparticles, become hardly visible.

Table 2: Quantification of PPA grafted at the TiO₂ nanoparticles surface for TiO₂/PPA-20°C, TiO₂/PPA-45°C, TiO₂/PPA-90°C, TiO₂/PPA-120°C and TiO₂/PPA-150°C.

Specimen	PPA grafted at the TiO ₂ surface in mmol/g *	PPA grafted at the TiO ₂ surface in mg/g
TiO ₂ /PPA-20°C	0.10	15.8
TiO ₂ /PPA-45°C	0.13	20.5
TiO ₂ /PPA-90°C	0.12	19.0

TiO ₂ /PPA-120°C	0.11	17.4
TiO ₂ /PPA-150°C	0.08	13.2

*: the error is in the order of 0.01 mmol/g

Quantification of the amount of PPA grafted at the TiO₂ nanoparticle surface (see supporting information for more details). By integrating the quantitative ³¹P MAS spectra (**Table S4**), the amount of PPA grafted at the TiO₂ surface can be determined for all specimens obtained from the reaction of PPA with TiO₂ at 20°C, 45°C, 90°C, 120°C and 150°C. The results are reported in **Table 2** from which it can be observed that the amount of PPA grafted at the TiO₂ nanoparticle surface only shows limited variation, indicating that the reaction temperature only has a minor effect on the amount of surface grafting. It further indicates that the amount of surface grafting under the applied reaction conditions (0.1 M PPA in water as solvent) is rather limited. The authors presume that this limited amount originates from the use of an excess of PPA which, once it is physically adsorbed to the TiO₂ surface, reduces the accessibility of the Ti-OH groups.

Table 3: Quantification of PPA involved in the titaniumphenylphosphonate structure for TiO₂/PPA-20°C, TiO₂/PPA-45°C, TiO₂/PPA-90°C, TiO₂/PPA-120°C and TiO₂/PPA-150°C.

Specimen	PPA in titaniumphenylphosphonate in mmol/g	PPA in titaniumphenylphosphonate in mg/g
TiO ₂ /PPA-20°C	/	/
TiO ₂ /PPA-45°C	0.004	0.7
TiO ₂ /PPA-90°C	0.08	12.1
TiO ₂ /PPA-120°C	0.5	84.6
TiO ₂ /PPA-150°C	1.3	206.4

*: the error is in the order of 0.01 mmol/g

Quantification of the amount of PPA involved in titaniumphenylphosphonate structures. The quantification results regarding PPA involved in the titaniumphenylphosphonate formation are reported in **Table 3**. While no titaniumphenylphosphonate is observed in TiO₂/PPA-20°C, an increasing amount

is noticed as a function of the reaction temperature. A 325 fold increase of the molar amount of titaniumphenylphosphonate is obtained by increasing the reaction temperature from 45°C to 150°C. The quantification of TiO₂/PPA-90°C shows that for a reaction temperature of 90°C, the molar amount of PPA grafted at the TiO₂ nanoparticles (**Table 2**) is almost equal to the amount of PPA involved in titaniumphenylphosphonate (**Table 3**), i.e. the difference is only 0.04 mmol/g. For TiO₂/PPA-150°C, the amount of PPA in titaniumphenylphosphonate is even sixteen times higher than the amount of PPA grafted at the TiO₂ nanoparticles.

In conclusion it can be stated that the amount of PPA grafted at the TiO₂ nanoparticle surfaces is almost independent on the synthesis reaction temperature and so it will not influence the ratio PPA-grafted/titaniumphenylphosphonate. Conversely, the formation of titaniumphenylphosphonate becomes predominant in case of enhanced temperatures (T > 100°C).

Quantification of TiO₂/PPA-150°C-HCl-treated. In order to investigate the stability of the P-O-Ti bonds towards acid hydrolysis, TiO₂/PPA-150°C was exposed to a 1M solution of concentrated hydrochloric acid for 12 hours (TiO₂/PPA-150°C-HCl-treated). The ³¹P MAS NMR spectrum of TiO₂/PPA-150°C-HCl-treated in the presence of the OPA standard (TiO₂/PPA-150°C-HCl-treated-OPA) is shown in the supporting information (Fig S2). **Table 4** shows the quantitative results of a comparison with the untreated TiO₂/PPA-150°C.

Table 4: Quantification of PPA in TiO₂/PPA-150°C and TiO₂/PPA-150°C–HCl-treated.

Sample	PPA grafted at the TiO ₂ nanoparticles surface mmol/g	PPA grafted at the TiO ₂ nanoparticles surface mg/g	PPA in titaniumphenylphosphonate mmol/g	PPA in titaniumphenylphosphonate mg/g
TiO ₂ /PPA-150°C	0.0830	13.2	1.30	206
TiO ₂ /PPA-150°C-HCl-treated	0.0012	0.2	1.29	204

*: the error is in the order of 0.01 mmol/g

The quantification of TiO₂/PPA-150°C-HCl-treated revealed that the amount of PPA involved in the titaniumphenylphosphonate structure remains unaffected. Prolonged acid treatment did not initiate hydrolysis of the Ti(C₆H₅PO₃)₂ units³⁶. In contrast, **Table 4** illustrates that the amount of PPA grafted at the TiO₂ nanoparticles dramatically decreases as a result of the corrosive treatment. In fact, the PPA grafted at the TiO₂ surfaces in TiO₂/PPA-150°C-HCl-treated decreased with 98.5% as compared to the untreated TiO₂/PPA-150°C material, revealing that the surface grafted PPA is not resisting extremely hydrolytic conditions. This large difference in stability of P-O-Ti bonds in titaniumphenylphosphonate versus in grafted PPA can be explained by the self-assembly of the Ti(C₆H₅PO₃)₂ units into crystal layers with hydrophobic character, shielding the P-O-Ti bonds from the acid catalyzed hydrolysis. Conversely, the PPA grafted at the TiO₂ surfaces is not shielded and enhanced hydrolysis of these P-O-Ti bonds takes place.

Quantification of post solvent exposed TiO₂/PPA-150°C, i.e. after use as stationary phase in SPE experiments (TiO₂/PPA-150°C-post-SPE). SPE experiments were carried out in order to evaluate the stability of TiO₂/PPA-150°C towards organic solvents exposure. The stability of this class of hybrid materials in organic solvents is crucial since many dedicated and chromatography-related applications require materials operating under such conditions³⁷. In addition, this experiment was also performed in order to evaluate the intrinsic hydrophobicity of the TiO₂/PPA-150°C material by using it as stationary phase in the SPE column for the separation of toluene from a 1/1 mixture of methanol/toluene (**Figure S1**). The hydrophobicity of the TiO₂/PPA-150°C is related to the amount of toluene kept in the extraction column after the loading step. It was measured that TiO₂/PPA-150°C was able to separate 91% of the toluene present in the methanol/toluene mixture. This high separation efficiency can be explained by the strong hydrophobic and Van der waals interactions between toluene and the phenyl ring of PPA at the outside of the titaniumphenylphosphonate crystals³⁵ and at the nanoparticles surface. After the chromatography, the structural composition of the resulting material was evaluated by quantitative ³¹P MAS NMR using the OPA standard. The spectrum of TiO₂/PPA-150°C-post-SPE-OPA

is shown in the supporting information (Fig S3). **Table 5** reveals the characteristics of TiO₂/PPA-150°C before and after SPE.

Table 5: Quantification of PPA in TiO₂/PPA-150°C and TiO₂/PPA-150°C-post SPE.

Sample	PPA grafted at the TiO ₂ nanoparticles surface mmol/g	PPA grafted at the TiO ₂ nanoparticles surface mg/g	PPA in titaniumphenylphosphonate mmol/g	PPA in titaniumphenylphosphonate mg/g
TiO ₂ /PPA-150°C	0.083	13.2	1.30	206
TiO ₂ /PPA-150°C-post SPE	0.050	7.9	1.28	202

*: the error is in the order of 0.01 mmol/g

Similar as for TiO₂/PPA-150°C-HCl-treated (**Table 4**), the amount of PPA in the titaniumphenylphosphonate did not change significantly by the SPE. This indicates a high stability of the titaniumphenylphosphonate structure towards the exposure to apolar solvents like toluene and cyclohexane. This in contrast to a loss of 43% of the PPA grafted at the TiO₂ nanoparticles. This points to a loss of PPA modified nanoparticles during the CLWE (conditioning-loading-washing-elution) procedure of the SPE experiment³⁴.

CONCLUSIONS

For the first time the absolute quantification of phosphonic acid (PPA) modified TiO₂ nanoparticles is reported, and this by means of ³¹P MAS NMR and with using octylphosphonic acid as internal standard. By this method, it was further possible to quantify the PPA grafted at the TiO₂ surface separately from the titaniumphenylphosphonate originating from the dissolution of the amorphous part of the matrix at enhanced reaction temperatures. From measurements of hybrid materials obtained by reacting PPA with TiO₂ as a function of increasing reaction temperature, it can be established that the amount of PPA grafted at the TiO₂ nanoparticles surface is quasi independent on the reaction temperature. The amount of PPA involved in titaniumphenylphosphonate formation on the other hand is strongly dependent on the reaction temperature. At reaction temperatures above 100°C, titaniumphenylphosphonate is

predominantly formed. The proposed methodology also allows to evaluate the materials stability under acid, corrosive conditions. The ^{31}P MAS NMR spectra demonstrate that PPA grafted at the TiO_2 nanoparticles is strongly vulnerable to acid induced hydrolysis while the titaniumphenylphosphonate is extraordinary stable. Regarding solid phase extraction (SPE), it is observed that a part of the PPA modified nanoparticles is lost during the extraction of a 1/1 methanol/toluene mixture but that all of the titaniumphenylphosphonate is preserved. Thus, the quantification method allows scientists to fine-tune these TiO_2 hybrid materials and application engineers to evaluate the structure-properties relation and the influence of annealing stress hereon.

ASSOCIATED CONTENT

Supporting information file contains the figures: supporting information 1, supporting information 2, supporting information 3 and the paragraph “Quantitative evaluation of the solid state ^{31}P MAS NMR spectra”

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Author contributions

The manuscript was written through the contributions of all authors. All authors have given approval to the final version of the manuscript.

Notes

The authors declare no competing financial interest.

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