

# Reaction mechanisms and kinetics of the elimination processes of 2-chloroethylsilane and derivatives: A DFT study using CTST, RRKM, and BET theories

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## ABSTRACT

The thermal decomposition kinetics of 2-chloroethylsilane and derivatives in the gas phase has been studied computationally using density functional theory, along with various exchange-correlation functionals (UM06-2x and  $\omega$ B97XD) and the aug-cc-pVTZ basis set. The calculated energy profile has been supplemented with calculations of kinetic rate constants under atmospheric pressure and in the fall-off regime, using transition state theory (TST) and statistical Rice–Ramsperger–Kassel–Marcus (RRKM) theory. Activation energies and rate constants obtained using the UM06-2x/aug-cc-pVTZ approach are in good agreement with the experimental data. The decomposition of 2-chloroethyltriethylsilane species into the related products  $[C_2H_4 + Et_3SiCl]$  is characterized by 6 successive structural stability domains associated to the sequence of catastrophes  $C_8H_{19}SiCl: 6-C^{\ddagger}FCC^{\ddagger}[FF]-O: C_6H_{15}SiCl + C_2H_4$ . Breaking of Si–C bonds and formation of Si–Cl bonds occur in the vicinity of the transition state.

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## 1. Introduction

The unusual reactivity of 2-chloroethylsilanes and its derivatives has been studied in the gas phase and in solution of these interesting molecules. It is well known that 2-chloroethylsilanes with three alkyl groups attached to the silicon atom decompose readily to ethylene and chlorosilane by heating, in such a manner that replacement of the alkyl groups by halogen leads to increasing the thermal stability [1–5]. The kinetics of the gas-phase thermal decomposition of 2-chloroethylsilanes and derivatives in a static system indicates that the elimination process predominates to form ethylene, and that the overall decomposition reaction is a first-order, homogenous, and apparently unimolecular  $\beta$ -elimination reaction [6].

Davidson *et al.* [1,7–9] have studied the kinetics of the gas-phase decomposition of 2-chloroethyltrichlorosilane [1], 2-chloroethylethyldichlorosilane [7], 2-chloroethyl-diethylchlorosilane [8], and 2-chloroethyltriethylsilane [9] (see Scheme 1). Sommer *et al.* [5] and Batuev *et al.* [3] have shown that 2-chloroethylsilane with three alkyl groups bonded to silicon

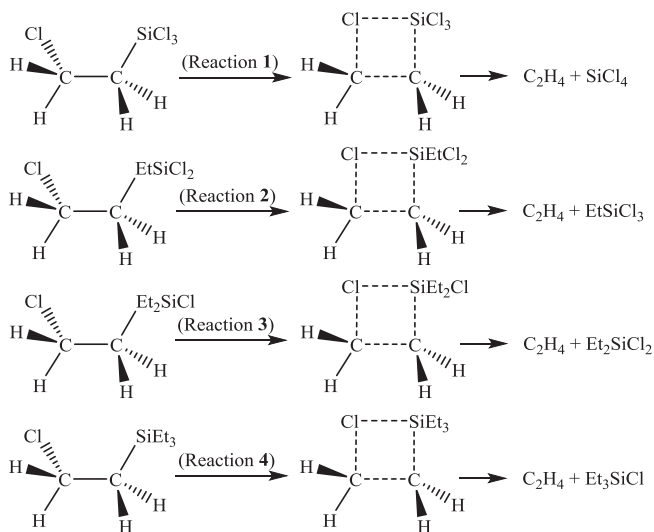
decomposes easily when heated and that replacement of the alkyl groups by halogen increases the thermal stability. When the elimination of 2-chloroethylsilane proceeds by a concerted, four-center, and homolytic process, the replacement of a chlorine attached to silicon by an electron-supplying group would be expected to increase the rate constant, while when a quasi-heterolytic process was involved, the rate constant would also be affected by such a replacement [7]. It follows that the increase in rate constant caused by the replacement of one chlorine atom by an ethyl group is due mainly to a change in steric rather than electronic effects (Scheme 1).

Davidson *et al.* [1,7–9] have showed that the activation energies for the elimination of ethylene from 2-chloroethylsilane decrease when the attached chlorine atom to silicon is replaced by ethyl groups. They also suggested that this is evidence for some charge separation in the transition state, with the expansion of a partial positive charge on silicon. In the elimination reaction electron distribution at silicon atom of reactant and corresponding product are different due to the change from the Si–C bond to a Si–Cl bond, but it is not possible to estimate the extent of this difference reliably [9].

An Arrhenius plot of all the kinetic unimolecular rate constants of the studied compounds that were measured or experimentally

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**Scheme 1.** Elimination processes of 2-chloroethylsilane and derivatives.

inferred over the temperature range 298–417 °C is shown in Fig. 1. Kinetic rate constant of the gas-phase unimolecular  $\beta$ -elimination processes of the compounds **1–4** exhibits positive temperature dependences over the studied temperature range, which is respectively equivalent to Arrhenius activation energies of (45.5  $\pm$  0.5) [1], (46.5  $\pm$  1.5) [7], (41.1  $\pm$  1.0) [8], and (39.2  $\pm$  1.0) kcal mol<sup>-1</sup> [9] (see Fig. 1). A least-square Arrhenius fit of the experimental data yields accordingly the following Arrhenius expression as follows [1,7–9]:

$$k_1(\text{s}^{-1}) = 10^{(11.40 \pm 0.16)} \exp\left(-\frac{45500 \pm 500}{RT}\right); [T = 356 - 417^\circ\text{C}]$$

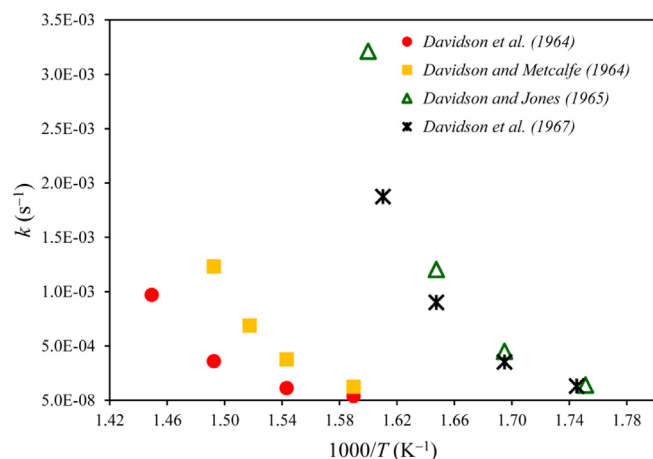
$$k_2(\text{s}^{-1}) = 10^{(12.21 \pm 0.50)} \exp\left(-\frac{46500 \pm 1500}{RT}\right); [T = 356 - 397^\circ\text{C}]$$

$$k_3(\text{s}^{-1}) = 10^{(11.88 \pm 0.26)} \exp\left(-\frac{41100 \pm 1000}{RT}\right); [T = 298 - 352^\circ\text{C}]$$

$$k_4(\text{s}^{-1}) = 10^{(11.07 \pm 0.35)} \exp\left(-\frac{39200 \pm 1000}{RT}\right); [T = 316 - 330^\circ\text{C}]$$

The above Arrhenius rate constants increases with increasing temperature at temperatures ranging from 298 to 417 °C, confirming positive activation energies.

In spite of the fact that the gas-phase reaction kinetics of the elimination processes of 2-chloroethylsilanes and its derivatives has been studied by Davidson *et al.* [1,7–9], the pressure depen-



**Fig. 1.** Arrhenius plot of the rate constant of the unimolecular  $\beta$ -elimination processes of the compounds **1–4**. Legend: (●) Davidson *et al.* [1]; (■) Davidson and Metcalfe [7]; (△) Davidson and Jones [8]; (✱) Davidson *et al.* [9].

dence of the rate constants and molecular mechanisms of these reactions are unknown. The aim of study is to consider activation energies, kinetic rate constants as well as the molecular mechanism of the thermal decomposition processes of 2-chloroethylsilane and derivatives that are sketched in Scheme 1. In this purpose, use shall be made of transition state theory (TST) [10–17], in conjunction with the dispersion-corrected  $\omega$ B97XD [18] and the UM06-2x [19] exchange-correlation functionals and Dunning's augmented triple zeta correlation-consistent polarized valence basis sets (aug-cc-pVTZ) [20]. In addition, kinetic rate constants at the high pressure limit will be supplied using TST, and their falloff behavior at lower pressures will be studied using statistical Rice–Ramsperger–Kassel–Marcus (RRKM) theory [21–23], for the purpose of unraveling the detailed experiments by Davidson *et al.* [1,7–9] at temperature ranging from 298 to 471 °C. Note that the UM06-2x functional has been especially designed both from the kinetic and thermodynamic viewpoint for accurate studies of chemical reactions. The aug-cc-pVTZ basis set is about twice as large as the 6-311+G\*\* basis set that was used in a previous study [24]. The data obtained at the B3LYP/6-311+G\*\* level in this study can be compared with our newly supplied UM06-2x data. The obtained results in the previous study need to be verified using more reliable exchange-correlation functionals and basis set. Due to the self-interaction error of commonly used B3LYP functional [25], transition barriers obtained with the B3LYP energy functional are indeed usually strongly underestimated [26,27].

## 2. Computational details

All the calculations were performed using the Gaussian 09 package of programs [28]. The molecular structures were visualized on the basis of the output data of the density functional theory (DFT) calculations using GaussView [29]. The structures of all stationary points have been optimized by using the DFT along with the  $\omega$ B97XD and UM06-2x exchange-correlation functionals, in conjunction a basis set incorporating diffuse functions (aug-cc-pVTZ basis set).

The corresponding calculations of harmonic vibrational frequencies at the same level of theory were then carried out in order to verify whether the stationary points are local minima or saddle points. For the possible transition state structures, the intrinsic reaction coordinate (IRC) approach have been used [30]. These calculations were carried out in both directions (forward and reverse) with 121 points at the UM06-2x/aug-cc-pVTZ level along the reaction path to check the connections between transition state structures and relevant minima.

Rate constants for the studied unimolecular reactions were investigated in the high pressure limit by means of transition state theory (TST) using the UM06-2x/aug-cc-pVTZ approach to estimate activation energies. The rationale for choosing the UM06-2x exchange-correlation functional is that a recent study by Zhao and Truhlar [19] has shown that it is the best one for studies involving main group thermochemistry, kinetics, noncovalent interactions, and electronic excitation energies to valence and Rydberg states. The M06-2x exchange-correlation functional and its analogs are dedicated for precisely energetic considerations. However, recently it has been established that this approach overestimate activation parameters for elimination processes [31–33].

A computational study on the kinetics of unimolecular reactions can be performed by means of conventional transition state theory. With this approach the rate constants for decomposition reactions can be obtained as follows [34–36]:

$$k_{\text{uni}} = \kappa(T) \frac{\sigma_{\text{B}} k_{\text{B}} T}{h} \frac{Q_{\text{TS}}^{\ddagger}(T)}{Q_{\text{A}}(T)} \exp(-E_{\text{a}}/RT) \quad (1)$$

where  $k_B$ ,  $h$  and  $R$  are the Boltzmann's, Planck's and ideal gas constants, respectively. In this equation,  $\sigma$  denotes a symmetry number, and  $\kappa(T)$  is the Wigner's tunneling correction factor [37].

$$\kappa(T) = 1 + \frac{1}{24} \left( \frac{h \text{Im}(v^\ddagger)}{k_B T} \right)^2 \quad (2)$$

where  $\text{Im}(v^\ddagger)$  is the imaginary vibrational frequency of the relevant transition state. Due to the simplicity of conventional TST, this method gives the upper limit of the kinetic rate constant, and enables to provide reliable rate constants in the limiting high-pressure behavior [23,38], especially important in the case of significant barrier heights.

In line with the temperatures at which the experiments by Davidson *et al.* [1,7–9] were conducted, unimolecular rate constants have been calculated for the temperature range 298–417 °C and pressure of 1.0 bar using conventional TST.

Note that, in practice, standard atmospheric pressures (1 bar) are usually considered to be large enough for reliably calculating rate constants using TST. The falloff behavior of canonical kinetic rate constants from the conventional TST limit ( $P \rightarrow \infty$ ) towards the low-pressure limit ( $P \rightarrow 0$ ) were also calculated with RRKM theory [21–23]. The microcanonical rate constants  $k(E)$  are calculated using the well-known equation of RRKM theory [21]:

$$k(E) = \frac{\sigma N^\ddagger(E)}{h \rho(E)} \quad (3)$$

where  $N^\ddagger(E)$  represents the total number of states at the transition state with energy less than or equal to  $E$ , and  $\rho(E)$  is the density of states at the reactants [33]. The canonical rate constant  $k(T)$  can be written as:

$$k(T) = \frac{1}{Q(T)} \int k(E) N(E) \exp(-\beta E) dE \quad (4)$$

where  $\beta = (kT)^{-1}$  and  $Q(T)$  is the internal state reactant partition function of the reactants.

All supplied TST and RRKM kinetic rate constants were obtained using the Kinetic and Statistical Thermodynamical Package (KiSTheLP) [39]. All these calculations rely upon UM06-2x/aug-cc-pVTZ predictions of ro-vibrational densities of states and activation energy barriers. Collisional stabilization rate constants were calculated using the Lennard–Jones (LJ) collision rate theory [40]. The strong collision approximation is used assuming that every collision deactivates with  $\omega = \beta_c Z_{LJ} [M]$  being the effective collision frequency

where  $Z_{LJ}$  is the Lennard–Jones collision frequency,  $[M]$  represents the total gas concentration, and  $\beta_c = 0.2$  denotes the collisional efficiency. The collision frequencies ( $Z_{LJ}$ ) are calculated from the LJ parameters: The retained LJ potential parameters for argon (as diluent gas) [38] and reactants **1–4** [41] are given in Table 1.

The bonding changes along the reaction pathway of the most favorable reaction [R(**4**); 2-chloroethyltriethylsilane  $\rightarrow$  C<sub>2</sub>H<sub>4</sub> + Et<sub>3</sub>SiCl] has been studied using bond evolution theory (BET) [42] in order to understand the breaking/forming bond process in the elimination reaction of 2-chloroethylsilanes. Bonding changes along reaction paths can be studied from the characterization of the electron-density reorganization using bonding evolution theory, consisting of the joint use of Electron Localization Function (ELF) topology [43,44] and the catastrophe theory (CT) [45] as a new tool for the contemporary understanding of electronic rearrangements in chemical processes. The ELF topological analysis has been done with the TopMod program [46] and the Multiwfn suite of tools [47] using the UM06-2x/aug-cc-pVTZ monodeterminantal wavefunctions along the reaction coordinate to locate the successive structural stability domains (SSDs).

### 3. Results and discussion

#### 3.1. Energetic and thermodynamic parameters

The reader is referred to Tables 2 and 3 for the total internal energies at 0 K, as well as the enthalpies and Gibb's free energies at 298 K of all identified species along the chemical pathways. In line with experimental Arrhenius activation energies [1,7–9], all DFT calculations most clearly show that the gas-phase thermal decomposition of compound **1** is an endothermic process ( $\Delta H \approx 0.7$ – $0.9$  kcal mol<sup>-1</sup>) whereas the decomposition of compounds **2–4** is an exothermic process ( $\Delta H \approx -3.4$ ,  $-4.8$ , and  $-10.0$  kcal mol<sup>-1</sup>, respectively). All unimolecular  $\beta$ -elimination reactions **1–4** are exoergic processes ( $\Delta G < 0$ ) at ambient temperature and pressure. From the energy profiles supplied in Table 2, the production of the Et<sub>3</sub>SiCl species is the most favored process among all products P1–P4, since the reaction pathway **4** is strongly exoergic ( $\Delta G \approx -22$  kcal mol<sup>-1</sup>) and exothermic ( $\Delta H \approx -10$  kcal mol<sup>-1</sup>).

The gas-phase unimolecular  $\beta$ -elimination processes of compounds **1–4** correspond to a concerted, four-center, and homolytic process. The replacement of a chlorine atom attached to silicon atom by an electron-supplying (ethyl) group is expected to decrease the activation energy and therefore the rate constant of the reaction pathway will be increased [48,49].

Note that, the energy barrier height ( $\Delta E_{0K}^\ddagger$ ) for the reaction pathway **4** is lower by about 5.30–6.92 kcal mol<sup>-1</sup> than the barrier for pathway **1**, which is consistent with the replacement of the three ethyl groups attached to silicon by chlorine atoms. Similar observations can be made when Gibb's free activation energies are considered: in spite of slightly unfavorable entropy effects, the Gibb's free energy for the 2-chloroethyltriethylsilane

**Table 1**  
Lennard–Jones (LJ) potential parameters.

Species	LJ potential parameters	
	$\sigma$ (Å)	$\epsilon/k_B$ (K)
Argon (as diluent gas)	3.465	113.5
2-Chloroethyltrichlorosilane ( <b>1</b> )	5.4	484.4
2-Chloroethylethylchlorosilane ( <b>2</b> )	5.8	480.7
2-Chloroethyldiethylchlorosilane ( <b>3</b> )	6.1	477.9
2-Chloroethyltriethylsilane ( <b>4</b> )	6.4	475.9

**Table 2**  
Reaction energies, reaction enthalpies and Gibb's free reaction energies (in kcal mol<sup>-1</sup>) for the thermal decomposition of compounds **1–4** at different DFT levels of theory.

Parameter	$\omega$ B97XD/aug-cc-pVTZ			UM06-2x/aug-cc-pVTZ			
	$\Delta E_{0K}$	$\Delta H_{298K}^\circ$	$\Delta G_{298K}^\circ$	$\Delta E_{0K}$	$\Delta H_{298K}^\circ$	$\Delta G_{298K}^\circ$	$\Delta S_{298K}^\circ$
2-Chloroethyltrichlorosilane $\rightarrow$ C <sub>2</sub> H <sub>4</sub> + SiCl <sub>4</sub>	-0.076	0.688	-10.529	0.197	0.917 (-2.425)	-10.131 (-13.882)	37.057
2-Chloroethylethylchlorosilane $\rightarrow$ C <sub>2</sub> H <sub>4</sub> + EtSiCl <sub>3</sub>	-3.741	-3.414	-13.688	-3.899	-3.031 (-5.447)	-15.129 (-17.137)	40.578
2-Chloroethyldiethylchlorosilane $\rightarrow$ C <sub>2</sub> H <sub>4</sub> + Et <sub>2</sub> SiCl <sub>2</sub>	-5.489	-4.742	-15.891	-5.622	-4.777 (-10.268)	-16.832 (-21.982)	40.437
2-Chloroethyltriethylsilane $\rightarrow$ C <sub>2</sub> H <sub>4</sub> + Et <sub>3</sub> SiCl	-10.273	-9.239	-21.587	-11.013	-9.953 (-14.407)	-22.426 (-25.998)	41.837

- Ref. [24]; The values in parentheses were calculated at the B3LYP/6-311++G\*\* level of theory.

**Table 3**  
Activation energies, enthalpies and Gibbs's free activation energies (in kcal mol<sup>-1</sup>), and activation entropies (in cal mol<sup>-1</sup> K<sup>-1</sup>) of transition states relative to the reactants along chemical pathways 1–4 at different DFT levels of theory (P = 1 atm).

Method	ωB97XD/aug-cc-pVTZ			UM06-2x/aug-cc-pVTZ			Experiment (ΔE <sub>0k</sub> )		
	ΔE <sub>0k</sub>	ΔH <sup>‡</sup> <sub>298K</sub>	ΔG <sup>‡</sup> <sub>298K</sub>	ΔE <sub>0k</sub>	ΔH <sup>‡</sup> <sub>298K</sub>	ΔG <sup>‡</sup> <sub>298K</sub>	ΔS <sup>‡</sup> <sub>298K</sub>		
2-Chloroethyltrichlorosilane → TS1	49.781	49.714	50.279	47.908 [48.674]	47.782 [48.586]	48.490 [49.572]	-2.371 (-8.0 ± 0.7) <sup>a</sup> [-3.308]	(45.5 ± 0.5) <sup>a</sup>	
Imaginary frequency TS1 (cm <sup>-1</sup> )	421.86i			463.31i					
2-Chloroethyltrichlorosilane → TS2	50.503	50.336	51.145	48.030 [50.136]	47.852 [50.077]	48.598 [50.712]	-2.504 (-3.6 ± 2.0) <sup>b</sup> [-2.128]	(46.5 ± 1.5) <sup>b</sup>	
Imaginary frequency TS2 (cm <sup>-1</sup> )	469.85i			487.81i					
2-Chloroethyldiethylchlorosilane → TS3	49.652	49.496	50.448	47.202 [47.504]	47.085 [47.409]	47.129 [48.529]	-0.147 [-3.754]	41.1 <sup>c</sup>	
Imaginary frequency TS3 (cm <sup>-1</sup> )	458.40i			476.10i					
2-Chloroethyltriethylsilane → TS4	44.479	44.723	44.413	40.988 [41.729]	41.191 [41.981]	41.063 [41.991]	+0.430 [-0.034]	39.2 <sup>d</sup>	
Imaginary frequency TS4 (cm <sup>-1</sup> )	282.09i			238.15i					

- Experimental values: a) Ref. [11]; b) Ref. [7]; c) Ref. [8]; d) Ref. [9].  
- Ref. [24]: The values in square brackets were calculated at the B3LYP/6-311++G\*\* level of theory.

(compound **4**) species (41.06–44.41 kcal mol<sup>-1</sup>) is lower than the ones for the 2-chloroethyltrichloro-silane (compound **1**) species (48.49–50.28 kcal mol<sup>-1</sup>).

Indeed, the increase in the rate constant of dehydrochlorination is determined not only by the activation barrier, but also by the activation entropy (ΔS<sup>‡</sup>) which is required to reach the transition state structure [7,50]. The positive activation entropy of reactant **4** compared with reactants **1–3** can be due to a greater charge separation and therefore a less restricted transition state for reactant **4**, or may simply reflect greater restriction on rotations in the initial state of this compound [1,7–9]. In the elimination of ethylene from compounds **1–4**, there will be of course a difference in electron distribution at silicon between compound **1** (Cl–C<sub>2</sub>H<sub>4</sub>–SiCl<sub>3</sub>) and compound **4** (Cl–C<sub>2</sub>H<sub>4</sub>–SiEt<sub>3</sub>) resulting from the replacement of a Si–C bond by a Si–Cl bond [9].

Our best (UM06-2x) estimates for the activation energies are most generally in good agreement with the obtained results at the B3LYP/6-311+G\*\* level of theory [24], and the largest discrepancy compared with this work amounts to ~2.1 kcal mol<sup>-1</sup> for the activation energy characterizing the decomposition reaction of compound **2** into the related products P<sub>2</sub> [C<sub>2</sub>H<sub>4</sub> + EtSiCl<sub>3</sub>] (see Table 3). A possible reason for this difference is that the 6-311+G (2df,p) basis set may fail to quantitatively describe strong polarization effects. Indeed, it is worth reminding that the 6-311G basis set from which the 6-311+G(2df,p) basis set derives is known to exhibit a too compact 2p space [51], and to be effectively a basis set of double-zeta quality only in the s-space, due to improperly balanced expansion coefficients [52,53].

### 3.2. Kinetic parameters

TST and RRKM unimolecular rate constants for the decomposition processes of the studied compounds were calculated in the gas-phase along with the UM06-2x/aug-cc-pVTZ approach (Table 4) at a pressure of 1 bar and at the considered temperatures, in line with the original experiments [1,7–9]. Further RRKM data computed at lower and higher pressures are provided for the same temperatures in Tables S1a–i of the Supporting Information. Theoretical values obtained at a pressure of 1 bar do not differ by more than one order of magnitude from the available experimental data [1,7–9].

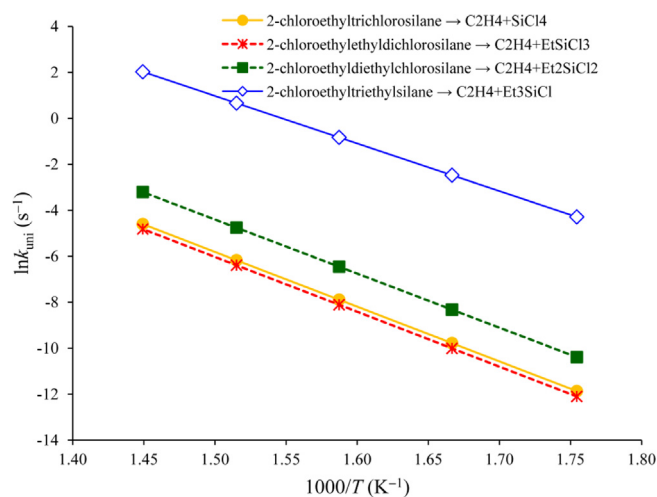
Whatever the considered temperatures, the unimolecular rate constant for the formation of the Et<sub>3</sub>SiCl species is larger than that obtained for the SiCl<sub>4</sub>, EtSiCl<sub>3</sub>, and Et<sub>2</sub>SiCl<sub>2</sub> species, which is in line with a reduction of the activation energy barrier by 6.92, 7.04, and 6.21 kcal mol<sup>-1</sup>, respectively on the corresponding chemical reaction pathways. Indeed, the obtained TST results (Table 4) indicate that kinetic rate constants for the 2-chloroethyltriethylsilane → C<sub>2</sub>H<sub>4</sub> + Et<sub>3</sub>SiCl unimolecular reaction are larger than the rate constants obtained for the other elimination reaction pathways. As is to be expected, because of the involved positive energy barriers, these rate constants increase gradually with increasing temperatures.

An Arrhenius plot of the obtained unimolecular rate constants by means of RRKM theory for pathways 1–4, based on the UM06-2x energy profiles (see Fig. 2) obviously confirms that the production of the C<sub>2</sub>H<sub>4</sub> + Et<sub>3</sub>SiCl (P<sub>4</sub>) species will be the fastest process under atmospheric pressure and at temperatures ranging from 298 to 417 °C. The same observation holds for pressures ranging from 10<sup>-12</sup> to 10<sup>2</sup> bars (Tables S1a–e in the Supporting Information). Since the involved energy barriers are large, the formation of the SiCl<sub>4</sub>, EtSiCl<sub>3</sub>, and Et<sub>2</sub>SiCl<sub>2</sub> species is characterized by lower rate constants at the considered temperatures, compared with the formation of the Et<sub>3</sub>SiCl species: the rate constants for these processes are smaller by at least 2–3 orders of magnitude than

**Table 4**  
Unimolecular rate constants (in  $s^{-1}$ ) for the reported reaction channels obtained by means of TST and RRKM theories ( $P = 1$  bar), according to the computed UM06-2x/aug-cc-pVTZ energy barriers.

Pathway	TST				RRKM			
	R1 → P1 (Reaction 1)	R2 → P2 (Reaction 2)	R3 → P3 (Reaction 3)	R4 → P4 (Reaction 4)	R1 → P1 (Reaction 1)	R2 → P2 (Reaction 2)	R3 → P3 (Reaction 3)	R4 → P4 (Reaction 4)
570	$7.05 \times 10^{-6}$ ( $8.97 \times 10^{-7}$ ) <sup>a</sup>	$5.58 \times 10^{-6}$ ( $2.69 \times 10^{-6}$ ) <sup>b</sup>	$3.09 \times 10^{-5}$ ( $1.09 \times 10^{-4}$ ) <sup>c</sup>	$1.38 \times 10^{-2}$ ( $1.32 \times 10^{-4}$ ) <sup>d</sup>	$6.78 \times 10^{-6}$ ( $8.97 \times 10^{-7}$ ) <sup>a</sup>	$5.33 \times 10^{-6}$ ( $2.69 \times 10^{-6}$ ) <sup>b</sup>	$2.95 \times 10^{-5}$ ( $1.09 \times 10^{-4}$ ) <sup>c</sup>	$6.88 \times 10^{-3}$ ( $1.32 \times 10^{-4}$ ) <sup>d</sup>
600	$5.67 \times 10^{-5}$ ( $6.69 \times 10^{-6}$ ) <sup>a</sup>	$4.52 \times 10^{-5}$ ( $2.09 \times 10^{-5}$ ) <sup>b</sup>	$2.43 \times 10^{-4}$ ( $6.17 \times 10^{-4}$ ) <sup>c</sup>	$8.45 \times 10^{-2}$ ( $8.09 \times 10^{-4}$ ) <sup>d</sup>	$5.48 \times 10^{-5}$ ( $6.69 \times 10^{-6}$ ) <sup>a</sup>	$4.34 \times 10^{-5}$ ( $2.09 \times 10^{-5}$ ) <sup>b</sup>	$2.34 \times 10^{-4}$ ( $6.17 \times 10^{-4}$ ) <sup>c</sup>	$4.23 \times 10^{-2}$ ( $8.09 \times 10^{-4}$ ) <sup>d</sup>
630	$3.75 \times 10^{-4}$ ( $4.12 \times 10^{-5}$ ) <sup>a</sup>	$3.01 \times 10^{-4}$ ( $1.34 \times 10^{-4}$ ) <sup>b</sup>	$1.58 \times 10^{-3}$ ( $2.95 \times 10^{-3}$ ) <sup>c</sup>	$4.37 \times 10^{-1}$ ( $4.18 \times 10^{-3}$ ) <sup>d</sup>	$3.64 \times 10^{-4}$ ( $4.12 \times 10^{-5}$ ) <sup>a</sup>	$2.90 \times 10^{-4}$ ( $1.34 \times 10^{-4}$ ) <sup>b</sup>	$1.52 \times 10^{-3}$ ( $2.95 \times 10^{-3}$ ) <sup>c</sup>	$2.19 \times 10^{-1}$ ( $4.18 \times 10^{-3}$ ) <sup>d</sup>
660	$2.09 \times 10^{-3}$ ( $2.15 \times 10^{-4}$ ) <sup>a</sup>	$1.68 \times 10^{-3}$ ( $7.26 \times 10^{-4}$ ) <sup>b</sup>	$8.62 \times 10^{-3}$ ( $1.23 \times 10^{-2}$ ) <sup>c</sup>	$1.95 \times 10^0$ ( $1.86 \times 10^{-2}$ ) <sup>d</sup>	$2.03 \times 10^{-3}$ ( $2.15 \times 10^{-4}$ ) <sup>a</sup>	$1.63 \times 10^{-3}$ ( $7.26 \times 10^{-4}$ ) <sup>b</sup>	$8.34 \times 10^{-3}$ ( $1.23 \times 10^{-2}$ ) <sup>c</sup>	$9.74 \times 10^{-1}$ ( $1.86 \times 10^{-2}$ ) <sup>d</sup>
690	$1.01 \times 10^{-2}$ ( $9.71 \times 10^{-4}$ ) <sup>a</sup>	$8.13 \times 10^{-3}$ ( $3.39 \times 10^{-3}$ ) <sup>b</sup>	$4.07 \times 10^{-2}$ ( $4.50 \times 10^{-2}$ ) <sup>c</sup>	$7.62 \times 10^0$ ( $7.26 \times 10^{-2}$ ) <sup>d</sup>	$9.80 \times 10^{-3}$ ( $9.71 \times 10^{-4}$ ) <sup>a</sup>	$7.89 \times 10^{-3}$ ( $3.39 \times 10^{-3}$ ) <sup>b</sup>	$3.95 \times 10^{-2}$ ( $4.50 \times 10^{-2}$ ) <sup>c</sup>	$3.82 \times 10^0$ ( $7.26 \times 10^{-2}$ ) <sup>d</sup>

- Experimental values: a) Ref. [1]; b) Ref. [7]; c) Ref. [8]; d) Ref. [9].

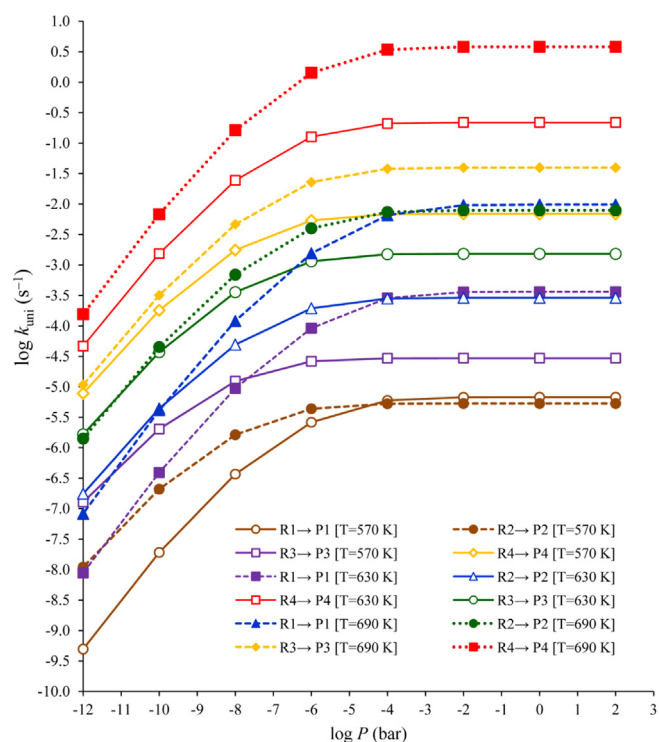


**Fig. 2.** Arrhenius plot of the obtained TST unimolecular rate constants [for compounds 1–4] at the UM06-2x/aug-cc-pVTZ level of theory ( $P = 1$  bar).

the rate constant characterizing the thermal decomposition of the 2-chloroethyltriethylsilane species into  $C_2H_4 + Et_3SiCl$ .

The reader is referred again to Fig. 2 for an Arrhenius plot of the obtained RRKM estimates at a pressure of 1 bar for the decomposition processes of pathways 1–4, according to the UM06-2x/aug-cc-pVTZ estimates of energy barriers. This Figure clearly confirms that the production of the ethylene and 2-chloroethyltriethylsilane species (via pathway 4) is the faster process at all studied temperatures, and this down to extremely low pressures, larger than  $10^{-12}$  bar. The same conclusion holds at much higher and lower pressures ( $10^{-12}$ – $10^2$  bar) (Tables S1a–e of the Supporting Information).

Inspection of Fig. 3 and Table 4 shows that the RRKM unimolecular rate constants obtained for the reported chemical reaction pathways increase with increasing temperatures. Furthermore, upon inspecting the RRKM data displayed in Fig. 3, it appears quite clearly that, in line with rather larger energy barriers, ranging from 41.98 to 47.91 kcal mol<sup>-1</sup>, pressures larger than  $10^{-4}$  bar are in general sufficient for ensuring a saturation within 3% accuracy of the computed unimolecular kinetic rate constants compared with the high-pressure limit (TST) of the RRKM unimolecular rate constants. The comparison with the RRKM data nevertheless indicates that the TST approximation breaks down at pressures lower than  $10^{-6}$  bar for the unimolecular rate constant characterizing pathway 4 [ $k_{uni}(4)$ ] (i.e. 2-chloroethyltriethylsilane →  $C_2H_4 + Et_3SiCl$ ), due to the strong positive activation energy (40.98 kcal mol<sup>-1</sup>)

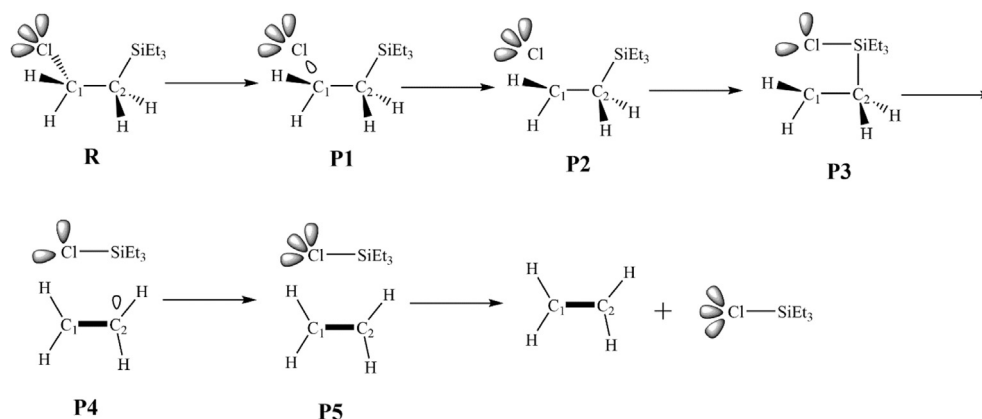


**Fig. 3.** Pressure dependence of the unimolecular rate constants of the  $\beta$ -elimination processes of compounds 1–4 according to the UM06-2x/aug-cc-pVTZ energy profiles.

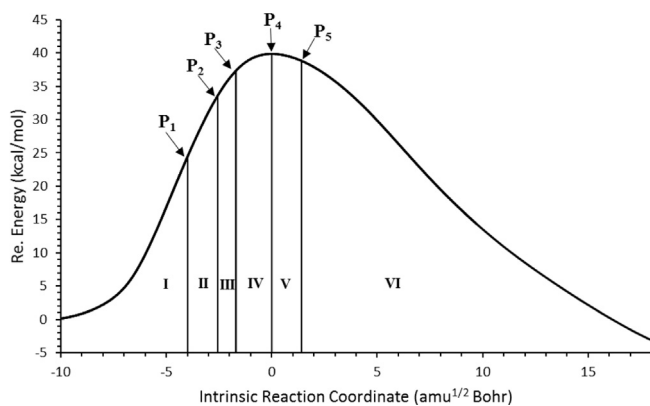
characterizing this reaction pathway. At a pressure of 1 bar, detailed inspection of Table 4 shows that ratios between the TST and RRKM estimates for this rate constant increases from  $\sim 4.7$  to  $\sim 18.2$  as the temperature increases from 298 to 471 °C.

### 3.3. Bonding evolution theory (BET) study of favorable unimolecular reaction

Lewis type representation of bonding changes in different steps on the IRC path, as revealed from BET are graphically represented in Fig. 4. The BET study of the elimination of ethylene from 2-chloroethyltriethylsilane indicates that this reaction along the intrinsic reaction coordinate can be topologically characterized by six differentiated structural stability domains (SSDs). The reaction path of the elimination processes is calculated by means of the IRC method and presented in Fig. 5 with marked domains.



**Fig. 4.** Lewis type representation of bonding changes for the elimination of ethylene from 2-chloroethyltriethylsilane based on the ELF topology. Filled lines indicate disynaptic basins with a population larger than  $3.0 \bar{e}$ . Large ellipses indicate lone electron pairs monosynaptic basins and small ellipses indicate non-bonding electron density with a population smaller than  $1.0 \bar{e}$ .



**Fig. 5.** Energy profile for the elimination of ethylene from 2-chloroethyltriethylsilane calculated by means of the IRC method, with a step size of  $0.1 \text{ amu}^{1/2} \text{ Bohr}$ . The selected points on the IRC are considered for the ELF topological analysis.

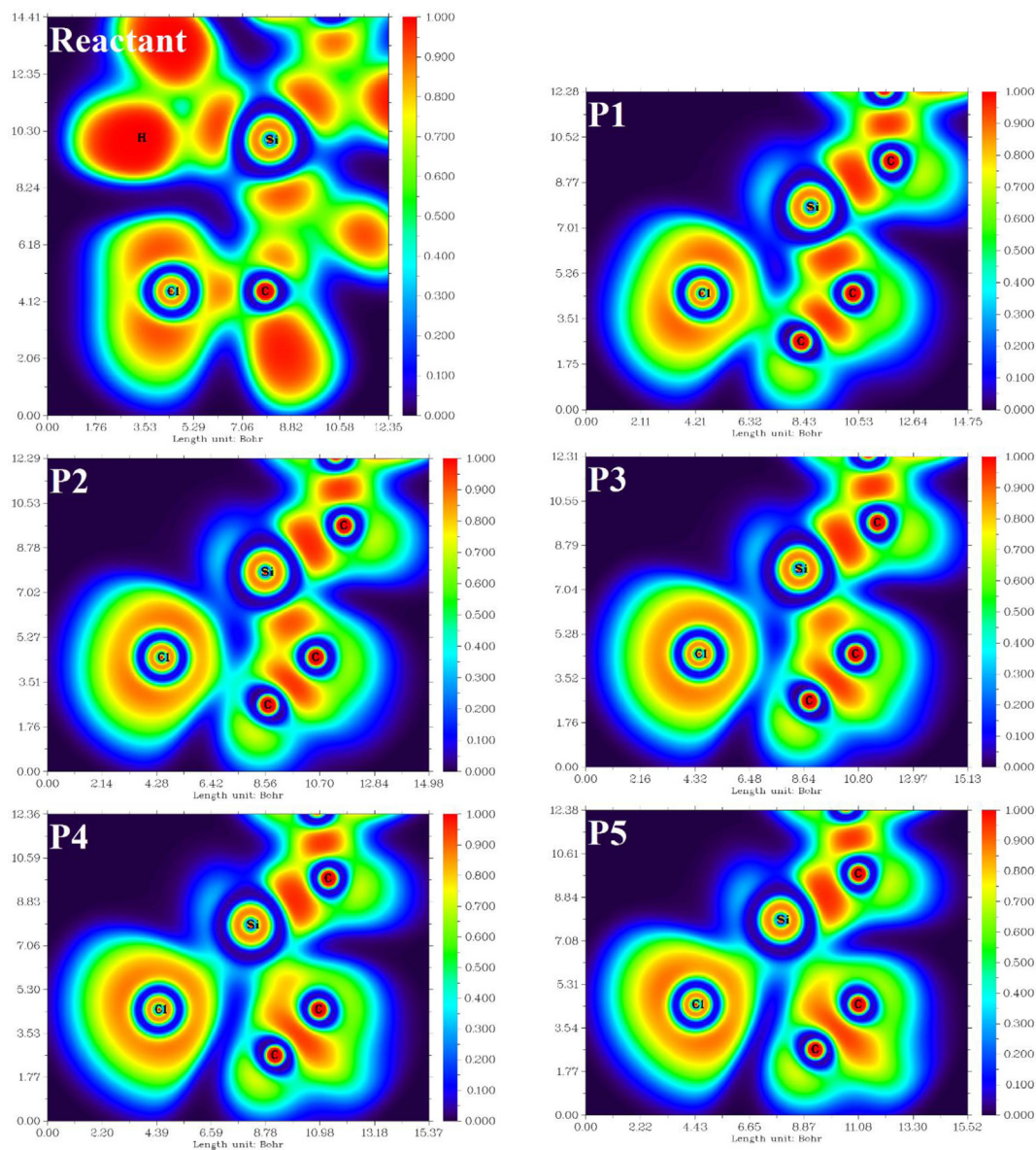
The ELF basin populations for 2-chloroethyltriethylsilane, ethylene,  $\text{Et}_3\text{SiCl}$  and the points  $P_i$  ( $i = 1-5$ ) along the reaction coordinate between the SSDs are tabulated in Table 5. The SSD I, the

most costly phase in term of energy with an increment of  $29 \text{ kcal mol}^{-1}$ , begins at the level of 2-chloroethyltriethylsilane and ends before the  $P_1$  point, i.e.  $R_x = -3.98829 \text{ amu}^{1/2} \text{ Bohr}$ . The ELF analysis of 2-chloroethyltriethylsilane shows three disynaptic basins  $V$  ( $\text{C}_1\text{-Cl}$ ),  $V$  ( $\text{C}_1\text{-C}_2$ ), and  $V$  ( $\text{Si-C}_2$ ) corresponding to the  $\text{C}_1\text{-Cl}$ ,  $\text{C}_1\text{-C}_2$ , and  $\text{Si-C}_2$  single bonds whose electronic population integrates to 1.34, 1.90, and 2.05  $\bar{e}$ , respectively; three disynaptic basins  $V$  ( $\text{C}_3\text{-C}_4$ ),  $V$  ( $\text{C}_5\text{-C}_6$ ), and  $V$  ( $\text{C}_7\text{-C}_8$ ) associated with the  $\text{C}_3\text{-C}_4$ ,  $\text{C}_5\text{-C}_6$ , and  $\text{C}_7\text{-C}_8$  single bonds integrating to 1.81  $\bar{e}$  each; three disynaptic basins  $V$  ( $\text{Si-C}_3$ ),  $V$  ( $\text{Si-C}_5$ ), and  $V$  ( $\text{Si-C}_7$ ) associated with the  $\text{Si-C}_3$ ,  $\text{Si-C}_5$ , and  $\text{Si-C}_7$  single bonds integrating to 2.06  $\bar{e}$  each; and three monosynaptic basins  $V$  ( $\text{Cl}$ ),  $V'$  ( $\text{Cl}$ ), and  $V''$  ( $\text{Cl}$ ) corresponding to the lone pairs located on the Cl atom with total population of 6.42  $\bar{e}$ . The SSD II starts at  $R_x = -3.98829 \text{ amu}^{1/2} \text{ Bohr}$  by the breaking of the covalent bond  $\text{C}_1\text{-Cl}$  and formation of a *pseudoradical* center  $V'''$  ( $\text{Cl}$ ) on the Cl atom with a population of 0.56  $\bar{e}$ . The SSD II ends before  $R_x = -2.55914 \text{ amu}^{1/2} \text{ Bohr}$  and energy in this SSD increases by  $9.13 \text{ kcal mol}^{-1}$ . The SSD III begins at  $R_x = -2.55914 \text{ amu}^{1/2} \text{ Bohr}$  by merging of the *pseudoradical* center generated in the previous SSD II into the lone pairs located at the Cl atom, and displacements of electron charge from the disynaptic basin  $V$  ( $\text{Si-C}_2$ ) to the disynaptic basin  $V$  ( $\text{C}_1\text{-C}_2$ ). The

**Table 5**

Integrated electron populations of valence ELF basins for different points defining the 6 SSDs which characterize the molecular mechanism associated with the elimination of ethylene from 2-chloroethyltriethylsilane.

	R	P <sub>1</sub>	P <sub>2</sub>	P <sub>3</sub>	P <sub>4</sub> (TS)	P <sub>5</sub>	$\text{C}_2\text{H}_4$	$\text{Et}_3\text{SiCl}$
$V$ ( $\text{C}_1\text{-Cl}$ )	1.34	–	–	–	–	–	–	–
$V$ ( $\text{C}_1\text{-C}_2$ )	1.90	2.16	2.37	2.51	3.1	3.46	1.64	–
$V'$ ( $\text{C}_1\text{-C}_2$ )	–	–	–	–	–	–	1.74	–
$V$ ( $\text{C}_3\text{-C}_4$ )	1.81	1.82	1.82	1.82	1.81	1.80	–	1.81
$V$ ( $\text{C}_5\text{-C}_6$ )	1.81	1.80	1.80	1.80	1.79	1.79	–	1.81
$V$ ( $\text{C}_7\text{-C}_8$ )	1.81	1.80	1.80	1.80	1.79	1.79	–	1.81
$V$ ( $\text{Si-C}_2$ )	2.05	1.75	1.47	1.23	–	–	–	–
$V$ ( $\text{Si-C}_3$ )	2.06	2.05	2.04	2.03	2.02	2.03	–	2.06
$V$ ( $\text{Si-C}_5$ )	2.06	2.06	2.07	2.07	2.07	2.06	–	2.07
$V$ ( $\text{Si-C}_7$ )	2.06	2.06	2.07	2.07	2.07	2.07	–	2.07
$V$ ( $\text{Si-Cl}$ )	–	–	–	0.94	1.01	1.06	–	1.55
$V$ ( $\text{C}_2$ )	–	–	–	–	0.37	–	–	–
$V$ ( $\text{Cl}$ )	2.22	2.92	3.18	3.28	3.36	2.78	–	2.46
$V'$ ( $\text{Cl}$ )	2.27	2.93	3.25	3.34	3.50	2.86	–	2.49
$V''$ ( $\text{Cl}$ )	1.93	1.20	1.16	–	–	1.20	–	1.85
$V'''$ ( $\text{Cl}$ )	–	0.56	–	–	–	–	–	–



**Fig. 6.** Color-filled maps of the electron localization function in the molecular plane defined by the Si–Cl–C<sub>1</sub> atoms for reactant, at the selected points along the reaction coordinate for the elimination of ethylene from 2-chloroethyltriethylsilane. The ELF values [0 to 1] are mapped on a blue–green–red color scale indicated on the right of each representation. Blue and red colors indicate respectively no electron and high electron localization.

energy in SSD III increases by 3.70 kcal mol<sup>-1</sup>. The SSD IV starts at  $R_x = -1.70165 \text{ amu}^{1/2} \text{ Bohr}$  by the disappearance of the monosynaptic basin  $V''$  (Cl) and appearance of the disynaptic basin V (Si–Cl) associated to the formation of a new covalent bond Si–Cl with an electron population equal to 0.94  $\bar{e}$ . This SSD ends before the TS point with an increment of 2.54 kcal mol<sup>-1</sup> to reach the breaking of the covalent bond Si–C<sub>2</sub>.

The SSD V begins at the TS point by the disappearance of the disynaptic basin V (Si–C<sub>2</sub>) corresponding to the breaking of the covalent bond Si–C<sub>2</sub> and formation of a new *pseudoradical* center on the C<sub>2</sub> atom with a population of 0.67  $\bar{e}$ . The SSD V ends before  $R_x = -1.40773 \text{ amu}^{1/2} \text{ Bohr}$ . The SSD VI final turning point relates to the depopulation of monosynaptic basins V (Cl),  $V'$  (Cl), and formation of monosynaptic basin  $V''$  (Cl) on Cl atom with an integrated population of 1.20  $\bar{e}$ . Along the SSD VI ethylene detachment and consequently elimination process is completed. Color-filled maps of the ELF in the molecular plane defined by plane of the Si–Cl–C<sub>1</sub> for 2-chloroethyltriethylsilane, and selected

points defining the 6 SSDs characterizing the molecular mechanism of elimination process are presented in Fig. 6. As can be seen the breaking of C<sub>1</sub>–Cl and Si–C<sub>2</sub> bonds starts at P<sub>1</sub> and P<sub>4</sub> points, respectively and their ELF values decreases continuously to approach to zero value at the end of the process. On the other hand, the formation of Si–Cl bond starts at P<sub>3</sub> points with ELF value 0.3. Elimination of ethylene from 2-chloroethyltriethylsilane can be characterized by 6 SSDs associated to the sequence of catastrophes  $C_8H_{19}SiCl: 6-C^1FCC^1[FF]-0: C_6H_{15}SiCl + C_2H_4$ .

#### 4. Conclusion

The thermal decomposition of 2-chloroethyltrichlorosilane (1), 2-chloroethylethyl- dichlorosilane (2), 2-chloroethyldiethylchlorosilane (3), and 2-chloroethyltriethylsilane (4) has been studied in the gas phase on computational grounds using density functional theory along with various exchange-correlation

functionals ( $\omega$ B97XD, and UM06-2x) and the aug-cc-pVTZ basis set.

The calculated energy profiles have been supplemented with calculations of kinetic rate constants under atmospheric pressure and in the fall-off regime, down to pressure of  $10^{-12}$  bar, using conventional TST and RRKM theories. Activation energies and rate constants obtained using the UM06-2x/aug-cc-pVTZ approach are in good agreement with the experimental data. RRKM calculations show that pressures larger than  $10^{-4}$  bar are sufficient for ensuring the validity of the TST approximation for all studied reaction channels. The molecular mechanism of the most favorable reaction has been studied using the BET theory. BET analysis based on ELF topology indicates that the elimination of ethylene from 2-chloroethyltriethylsilane can be characterized by 6 SSDs associated to the sequence of catastrophes  $C_8H_{19}SiCl: 6-C^{\dagger}FCC^{\dagger}[FF]-0: C_6H_{15}-SiCl + C_2H_4$ .

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### Appendix A. Supplementary data

Supplementary data associated with this article can be found, in the online version, at <http://dx.doi.org/10.1016/j.chemphys.2017.01.009>.

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# Reaction mechanisms and kinetics of the elimination processes of 2-chloroethylsilane and derivatives: A DFT study using CTST, RRKM, and BET theories

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## Supporting Information

**Table S1:** Unimolecular rate constants for all reaction steps involved in the reported chemical pathways (results obtained by means of RRKM theory at different pressures and temperatures, according to the computed UM06-2x/aug-cc-pVTZ energy profiles).

**Table S1a:** [ $T = 570$  K]

$P$ (bar)	$k_{\text{uni}}$ (R1→P1)	$k_{\text{uni}}$ (R2→P2)	$k_{\text{uni}}$ (R3→P3)	$k_{\text{uni}}$ (R4→P4)	$\log P$	$\log k_{\text{uni}}$ (R1→P1)	$\log k_{\text{uni}}$ (R2→P2)	$\log k_{\text{uni}}$ (R3→P3)	$\log k_{\text{uni}}$ (R4→P4)
1.00E+02	6.78E-06	5.33E-06	2.95E-05	6.88E-03	2	-5.17	-5.27	-4.53	-2.16
1.00E+00	6.78E-06	5.33E-06	2.95E-05	6.88E-03	0	-5.17	-5.27	-4.53	-2.16
1.00E-02	6.76E-06	5.33E-06	2.95E-05	6.88E-03	-2	-5.17	-5.27	-4.53	-2.16
1.00E-04	6.00E-06	5.29E-06	2.94E-05	6.81E-03	-4	-5.22	-5.28	-4.53	-2.17
1.00E-06	2.62E-06	4.37E-06	2.63E-05	5.39E-03	-6	-5.58	-5.36	-4.58	-2.27
1.00E-08	3.69E-07	1.63E-06	1.24E-05	1.75E-03	-8	-6.43	-5.79	-4.91	-2.76
1.00E-10	1.92E-08	2.09E-07	2.03E-06	1.81E-04	-10	-7.72	-6.68	-5.69	-3.74
1.00E-12	4.96E-10	1.10E-08	1.28E-07	7.82E-06	-12	-9.30	-7.96	-6.89	-5.11

**Table S1b:** [ $T = 600$  K]

$P$ (bar)	$k_{\text{uni}}$ (R1→P1)	$k_{\text{uni}}$ (R2→P2)	$k_{\text{uni}}$ (R3→P3)	$k_{\text{uni}}$ (R4→P4)	$\log P$	$\log k_{\text{uni}}$ (R1→P1)	$\log k_{\text{uni}}$ (R2→P2)	$\log k_{\text{uni}}$ (R3→P3)	$\log k_{\text{uni}}$ (R4→P4)
1.00E+02	5.48E-05	4.34E-05	2.34E-04	4.22E-02	2	-4.26	-4.36	-3.63	-1.37
1.00E+00	5.48E-05	4.34E-05	2.34E-04	4.22E-02	0	-4.26	-4.36	-3.63	-1.37
1.00E-02	5.45E-05	4.34E-05	2.34E-04	4.22E-02	-2	-4.26	-4.36	-3.63	-1.37
1.00E-04	4.60E-05	4.28E-05	2.32E-04	4.14E-02	-4	-4.34	-4.37	-3.63	-1.38
1.00E-06	1.73E-05	3.25E-05	1.94E-04	2.90E-02	-6	-4.76	-4.49	-3.71	-1.54
1.00E-08	2.07E-06	1.00E-05	7.48E-05	7.29E-03	-8	-5.68	-5.00	-4.13	-2.14
1.00E-10	9.54E-08	1.07E-06	9.68E-06	5.80E-04	-10	-7.02	-5.97	-5.01	-3.24
1.00E-12	2.30E-09	4.86E-08	5.09E-07	2.09E-05	-12	-8.64	-7.31	-6.29	-4.68

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**Table S1c:** [ $T = 630$  K]

$P$ (bar)	$k_{\text{uni}}$ (R1→P1)	$k_{\text{uni}}$ (R2→P2)	$k_{\text{uni}}$ (R3→P3)	$k_{\text{uni}}$ (R4→P4)	$\log P$	$\log k_{\text{uni}}$ (R1→P1)	$\log k_{\text{uni}}$ (R2→P2)	$\log k_{\text{uni}}$ (R3→P3)	$\log k_{\text{uni}}$ (R4→P4)
1.00E+02	3.64E-04	2.89E-04	1.52E-03	2.18E-01	2	-3.44	-3.54	-2.82	-0.66
1.00E+00	3.64E-04	2.89E-04	1.52E-03	2.18E-01	0	-3.44	-3.54	-2.82	-0.66
1.00E-02	3.60E-04	2.89E-04	1.52E-03	2.18E-01	-2	-3.44	-3.54	-2.82	-0.66
1.00E-04	2.87E-04	2.82E-04	1.50E-03	2.10E-01	-4	-3.54	-3.55	-2.82	-0.68
1.00E-06	9.20E-05	1.95E-04	1.15E-03	1.27E-01	-6	-4.04	-3.71	-2.94	-0.90
1.00E-08	9.46E-06	4.94E-05	3.59E-04	2.45E-02	-8	-5.02	-4.31	-3.44	-1.61
1.00E-10	3.92E-07	4.41E-06	3.71E-05	1.54E-03	-10	-6.41	-5.36	-4.43	-2.81
1.00E-12	8.83E-09	1.76E-07	1.66E-06	4.68E-05	-12	-8.05	-6.76	-5.78	-4.33

**Table S1d:** [ $T = 660$  K]

$P$ (bar)	$k_{\text{uni}}$ (R1→P1)	$k_{\text{uni}}$ (R2→P2)	$k_{\text{uni}}$ (R3→P3)	$k_{\text{uni}}$ (R4→P4)	$\log P$	$\log k_{\text{uni}}$ (R1→P1)	$\log k_{\text{uni}}$ (R2→P2)	$\log k_{\text{uni}}$ (R3→P3)	$\log k_{\text{uni}}$ (R4→P4)
1.00E+02	2.03E-03	1.63E-03	8.34E-03	9.74E-01	2	-2.69	-2.79	-2.08	-0.01
1.00E+00	2.03E-03	1.63E-03	8.34E-03	9.74E-01	0	-2.69	-2.79	-2.08	-0.01
1.00E-02	2.00E-03	1.63E-03	8.34E-03	9.73E-01	-2	-2.70	-2.79	-2.08	-0.01
1.00E-04	1.49E-03	1.56E-03	8.15E-03	9.12E-01	-4	-2.83	-2.81	-2.09	-0.04
1.00E-06	4.09E-04	9.60E-04	5.59E-03	4.63E-01	-6	-3.39	-3.02	-2.25	-0.33
1.00E-08	3.64E-05	2.01E-04	1.41E-03	6.84E-02	-8	-4.44	-3.70	-2.85	-1.17
1.00E-10	1.37E-06	1.52E-05	1.18E-04	3.46E-03	-10	-5.86	-4.82	-3.93	-2.46
1.00E-12	2.90E-08	5.35E-07	4.55E-06	9.08E-05	-12	-7.54	-6.27	-5.34	-4.04

**Table S1e:** [ $T = 690$  K]

$P$ (bar)	$k_{\text{uni}}$ (R1→P1)	$k_{\text{uni}}$ (R2→P2)	$k_{\text{uni}}$ (R3→P3)	$k_{\text{uni}}$ (R4→P4)	$\log P$	$\log k_{\text{uni}}$ (R1→P1)	$\log k_{\text{uni}}$ (R2→P2)	$\log k_{\text{uni}}$ (R3→P3)	$\log k_{\text{uni}}$ (R4→P4)
1.00E+02	9.80E-03	7.89E-03	3.95E-02	3.82E+00	2.00	-2.01	-2.10	-1.40	0.58
1.00E+00	9.80E-03	7.89E-03	3.95E-02	3.82E+00	0.00	-2.01	-2.10	-1.40	0.58
1.00E-02	9.59E-03	7.88E-03	3.95E-02	3.81E+00	-2.00	-2.02	-2.10	-1.40	0.58
1.00E-04	6.57E-03	7.41E-03	3.80E-02	3.43E+00	-4.00	-2.18	-2.13	-1.42	0.54
1.00E-06	1.55E-03	3.99E-03	2.29E-02	1.43E+00	-6.00	-2.81	-2.40	-1.64	0.16
1.00E-08	1.21E-04	6.94E-04	4.65E-03	1.63E-01	-8.00	-3.92	-3.16	-2.33	-0.79
1.00E-10	4.13E-06	4.51E-05	3.20E-04	6.80E-03	-10.00	-5.38	-4.35	-3.49	-2.17
1.00E-12	8.29E-08	1.41E-06	1.08E-05	1.56E-04	-12.00	-7.08	-5.85	-4.97	-3.81