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# Photomediated Controlled Radical Polymerization

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## Nomenclature

$\mathcal{D}$ : Dispersity

AA: Acrylic acid

AET: Associative electron transfer

AFM: Atomic-force microscopy

AIBN: 2,2'-Azobis(2-methylpropionitrile)

ARGET: Activators regenerated by electron transfer

ATRP: Atom transfer radical polymerization

BA: n-Butyl acrylate

BAPO: Phenylbis (2,4,6-trimethylbenzoyl)-phosphine oxide

BDMAT: *S,S'*-bis( $\alpha,\alpha'$ -dimethyl- $\alpha''$ -acetic acid)

BnBiB: Benzyl  $\alpha$ -bromoisobutyrate

BnMA: Benzyl methacrylate

BPN: 2-Bromopropionitrile

BPO: Benzoyl peroxide

Bpy: 2,2'-Bipyridine

Btp: 2-(2'-Benzothienyl)pyridine

BuMA: Butyl methacrylate

CEF: Chain end functionality

CMRP: Cobalt-mediated radical polymerization

CPEC: *S*-2-Cyano-2-propyl-*O*-ethyl xanthate

CPFDB: 2-Cyanoprop-2-yl(4-fluoro) dithiobenzoate

CRP: Controlled radical polymerization

CTA: Chain transfer agent

CuAAC: Copper(I)-catalyzed azide-alkyne cycloaddition

DCF: Dead chain fraction

DCPD: Dicyclopentadiene

DDMAT: *S*-Dodecyl-*S'*-( $\alpha,\alpha'$ -dimethyl- $\alpha''$ -acetic acid)

DET: Dissociative electron transfer

DLS: Dynamic light scattering

DLW: Direct laser writing

DMAEMA: 2-(Dimethylamino)ethyl methacrylate

DMF: Dimethylformamide

DMSO: Dimethyl sulfoxide

DP: Degrees of polymerization

Dppm: (Diphenylphosphino)methane

DT: Degenerative transfer

DVP: Dimethyl vinylphosphonate

EBiB: Ethyl  $\alpha$ -bromoisobutyrate

EBPA: Ethyl  $\alpha$ -bromophenylacetate

EDC: 1-Ethyl-3-(3-dimethylaminopropyl)carbodiimide

EEMA: 2-Ethoxyethyl methacrylate

EGMA: Ethylene glycol methyl ether methacrylate

EMA: Ethyl methacrylate

ESI: Electrospray-ionization

FTIR: Fourier transform infrared spectroscopy

GMA: Glycidyl methacrylate

HEMA: 2-Hydroxyethyl methacrylate

HMTETA: 1,1,4,7,10,10-Hexamethyltriethylenetetramine

HPMA: *N*-(2-Hydroxypropyl) methacrylamide

ICAR: Initiators for continuous activator regeneration

ITP: Iodine transfer polymerization

ISET: Inner-sphere electron transfer

LCST: Lower critical solution temperature

LFP: Laser flash photolysis

LRP: Living radical polymerization

MA: Methyl acrylate

MAA: Methacrylic acid

MBiB: Methyl 2-bromoisobutyrate

Me<sub>6</sub>TREN: Tris[2-(dimethylamino)ethyl]amine

MESA: Methyl(ethoxycarbonothioyl)sulfanyl acetate

MiB: Methyl isobutyrate

MIP: Molecularly imprinted polymer

MMA: Methyl methacrylate

MS: Mass spectrometry

$M_n$ : number-average molecular weight

$M_w$ : weight-average molecular weight

MW: molecular weight

MWD: molecular weight distribution

NAP: *N*-(2-Acryloyloxyethyl)pyrrolidone

NIPAAm: *N*-Isopropyl acrylamide

NMP: Nitroxide-mediated radical polymerization

NVP: *N*-Vinyl carbazole

OEGA: Oligo(ethylene glycol) methyl ether acrylate

OEOMA: Oligo(ethylene oxide) methyl ether methacrylate

OEGMA: Oligo(ethylene glycol) methyl ether methacrylate

OSET: Outer-sphere electron transfer

OMRP: Organometallic mediated radical polymerization

PEGA: Polyethylene glycol acrylate

PET: Photoinduced Electron Transfer

Ph: Phenyl

Phen: 1,10-Phenanthroline

PI: Photoinitiator

PISA: Polymerization-induced self-assembly

PLP: Pulsed-laser polymerization

PMA: Propargyl methacrylate

PMDETA : *N,N,N',N'',N'''*-Pentamethyldiethylenetriamine

PNB: Polynorbornene

Ppy: 2,2'-Phenylpyridine

PS: Photosensitizer

PTZ: 10-Phenylphenothiazine

QCM: Quartz crystal microbalance

RAFT: Reversible addition-fragmentation chain transfer

RDRP: Reversible-deactivation radical polymerization

ROMP: Ring-opening metathesis polymerization

RP: Radical polymerization

SARA: Supplemental activators and reducing agents

SAM: Self-assembled monolayer

SCE: Saturated calomel electrode

SDS: Sodium dodecyl sulfate

SIMS: Secondary ion mass spectrometry

SPMA: 3-Sulfopropyl methacrylate

SR&NI: Simultaneous reverse and normal initiation

St: Styrene

STXM: Scanning transmission X-ray microscopy

TEM: Transmission electron microscopy

TEMPO: 2,2,6,6-Tetramethylpiperidinyloxy-4-yl radical

TERP: Telluride-Mediated Polymerization

TFEMA: 1,1,1-Trifluoroethyl methacrylate

Tmd: 2,2,6,6-Tetramethyl-3,5-heptanedione

TMEDA: Tetramethylethylenediamine

TMPTA: Trimethylolpropane triacrylate

TPO: 2,4,6-Trimethylbenzoyl)diphenyl phosphine

TPMA: Tris((2-pyridine)methyl)amine

TPMA\*: Tris((4-methoxy-3,5-dimethylpyridin-2-yl)methyl)amine

Tren: Tris(2-aminoethyl)amine

TTC: Trithiocarbonate

UV: Ultraviolet

VAc: Vinyl acetate

VBA:  $\alpha$ -Vinylbenzyl alcohol

Vis: Visible

VL: Valerolactone

VP: Vinyl pivalate

XPS: X-ray photoelectron spectroscopy

## **Keywords**

Photopolymerization; radical polymerization; ATRP; RAFT; NMP

## **Abstract**

Photomediated controlled radical polymerization is a versatile method to prepare, under mild conditions, various well-defined polymers with complex architecture, such block and graft copolymers, sequence-controlled polymers, or hybrid materials via surface-initiated polymerization. It also provides opportunity to manipulate the reaction through spatiotemporal control. This review presents a comprehensive account of the fundamentals and applications of various Photomediated CRP techniques, including atom transfer radical polymerization (ATRP),

reversible addition-fragmentation chain transfer (RAFT), nitroxide mediated polymerization (NMP) and other procedures. In addition, mechanistic aspects of other photomediated methods are discussed.

## 1. Introduction

### 1.1 Reversible-deactivation radical polymerization

Conventional radical polymerization is a widely used technique to synthesize polymeric materials, which is typically performed under various reaction conditions such as bulk, in organic solvents or water, under homogeneous and heterogeneous conditions at different temperatures.[1] However, conventional radical polymerization cannot be used to synthesize well-defined polymers with precise average molar masses, narrow molecular weight distributions and pre-designed architectures because of unavoidable radical-radical terminations. Typical living polymerization techniques require a fast initiation and concurrent growth of all chains. To achieve living-like or controlled radical polymerization, several principles should be fulfilled: fast initiation, long lifetime of growing chains, and small fraction of terminated chains. A dynamic equilibrium between dormant species and growing radicals meets these criteria, and a process is termed as reversible deactivation radical polymerization (RDRP), or controlled radical polymerization (CRP).[2, 3] In “living” radical polymerization, there should always be some termination; however, a controlled radical polymerization can be accomplished by rapid exchange of growing radicals and dormant species, by increasing the concentration of dormant species or by reducing polymerization rate,. The dead chain fraction (DCF) is defined as the ratio of terminated chains (T) to all initiated chains (R-X), eq 1.[4] Decreasing DCF, i.e., increasing chain end functionality, could be realized by decreasing degree of polymerization ( $DP_T$ ), conversion ( $p$ ), or rate coefficients of termination to propagation ( $k_t/k_p$ ) and increasing initial monomer concentration ( $[M]_0$ ) and reaction time ( $t$ ).

$$DCF = \frac{[T]}{[R-X]_0} = \frac{2DP_T k_t [\ln(1-p)]^2}{[M]_0 k_p^2 t} \quad (1)$$

In general, there are three types of controlled radical polymerizations. The first type uses dormant species with a weak bond, which provides a propagating radical and a counter-radical ( $T^*$ ) via a homolytic cleavage (Scheme 1a). The counter-radical is a rather stable radical. It can only trap a propagating radical and should not react with monomer. Commercially available nitroxyl radicals

are used in a process, termed nitroxide mediated polymerization (NMP).[5] In a similar approach, organometallic mediated radical polymerization (OMRP) employs a weak bond between carbon and transition metal species, for example Co-porphyrins,[6, 7] and derivatives of Fe or Mo.[8]

### Insert Scheme 1.

It is possible to convert the first type of CRP to a catalytic system by using alkyl halides and transition metal complexes, as in atom transfer radical polymerization (ATRP). [9-11] Transition metal complexes in lower oxidation state ( $M^{I}/L$ ) activate functional alkyl bromide or chloride initiators to give initiating radical and deactivator complex in higher oxidation state ( $X-M^{II}/L$ ). The deactivator species deactivate propagating radicals to alkyl halides (dormant state) (Scheme 1b). Various metal complexes are used to catalyze ATRP including Cu, Fe, Ru, Ni, Os and other metals formed with amines, phosphines and other ligands, and copper catalyst with polydentate amine and pyridine based ligands is most used.[12]

The rate of ATRP is defined by eq 2, where radical concentration  $P_n^*$  is defined by the ATRP equilibrium constant ( $K_{ATRP}$ ), concentration of alkyl halide (dormant species) and the ratio of activator and deactivator concentrations.[10] It is worth to point out that the rate does not depend on the absolute concentration of activator.

$$R_p = k_p[M][P_n^*] = k_p K_{ATRP} \left( \frac{[P_n X][Cu^I/L][M]}{[X-Cu^{II}/L]} \right) \quad (2)$$

According to above equation, the same ratio of activator to deactivator concentrations would result in the same overall rate, even with a very low concentration of catalyst. However, traditional ATRP used high concentration of the activator to maintain the acceptable polymerization rate because each radical termination leads to the irreversible loss of the activator due to the persistent radical effect. [12, 13] Several procedures with low catalyst loading below 100 or even below 10 ppm were developed via regeneration of activator complex, as shown in Scheme 2.[14] The reduction of deactivator in higher oxidation state was accomplished by addition of chemical reducing agents in activator (re)generation by electron transfer (ARGET) ATRP,[14] an external radical initiator in initiators for continuous activator (re)generation (ICAR) ATRP,[15] a current in electrochemically mediated *e*ATRP,[16] zerovalent metals or sulfites in supplemental activators and reducing agents (SARA) ATRP,[17-20] or photochemical reduction in photoATRP.[21-25] Recently, ATRP based on small organic molecule, called metal-free (MF),

organocatalyzed (O) or photoinduced electron transfer (PET) ATRP, was reported using alkyl bromide as ATRP initiator.[26, 27] A similar method using alkyl iodides with various organic catalysts was also reported to provide well-defined polymers.[28]

### **Insert Scheme 2.**

The third RDRP method is based on activation-deactivation process which utilizes a degenerative transfer (Scheme 1c) between transfer agents ( $P_m-X$ ) and growing radicals ( $P_n^*$ ). Similar to conventional radical polymerization, the radical is formed externally, and is not formed internally from the dormant specie. The externally supplied radical generates a small number of growing chains, and also the polymerization stops when the source of radicals is consumed. Various transfer agents were used in degenerative transfer systems, including alkyl iodides, derivatives of tellurium or bismuth,[29] and thiocarbonylthio compounds, such as dithioesters, trithiocarbonates, xanthates and dithiocarbamates.[30] The method with dithio-compound, called reversible addition-fragmentation chain transfer (RAFT) polymerization, undergoes an exchange via addition-fragmentation process (Scheme 3).[31] In RAFT, kinetics follows the steady state approach, similar to the behavior in a conventional radical polymerization, but linear evolution of molecular weight with conversion and narrow molecular weight distributions resemble a controlled process.

### **Insert Scheme 3.**

Some experimentally quantifiable parameters for NMP, ATRP and RAFT were recently quantified and correlated.[32] Several borderline systems were also developed with both spontaneous and catalyzed cleavage, or degenerative transfer and ATRP or OMRP. It is interesting to note that equilibria between active in dormant species, typical for all RDRP systems, are also essential in many biological processes. [33]

## **1.2 Basic chemistry of photomediated processes**

### **1.2.1 Photoinitiators/photosensitizers**

Photoinitiated polymerization is a process that transforms monomers into polymers under ultra violet-visible (UV-Vis) light irradiation.[34] The polymerization is initiated by photogenerated reactive species, either free radicals or ions, which are produced from added photosensitive compounds, namely photoinitiators (PI) and/or photosensitizers (PS) (Figure 1).

### **Insert Figure 1.**

The key feature of photoinitiated polymerization is its ability to be conducted under mild reaction conditions. This provides several advantages, including more environmentally friendly and efficient processes and preparation of novel materials.[35, 36] Since the polymerization may proceed at room temperature, thermally sensitive molecules, such as pharmaceutical compounds, enzymes, peptides or proteins can be safely incorporated into the polymers during polymerization. Solvent-free formulation, or use of less solvent in the polymerization process, reduces the release of volatile compounds giving rise to environmental and health benefits. Unlike thermally initiated polymerization, photoinitiated polymerization provides spatial and temporal controls that make it possible to determine where, when and what type of polymerization process could be applied.[37] Moreover, the rate of initiation and polymerization can be controlled by adjusting light intensity or concentration of photosensitive compounds.

In the free radical mode, which is the focus of this article, photosensitive compounds generate active species either through unimolecular dissociation or bimolecular photoinduced electron transfer reactions upon exposure to UV or visible light.[38-40] A unimolecular photoinitiator undergoes a homolytic bond cleavage, producing active free radicals. In the case of a bimolecular photoinitiator, active free radicals are formed via a hydrogen abstraction or an electron-transfer mechanism between a photosensitizer, which absorbs the light and produces a long-lived excited triplet state, and a co-initiator that can quench the excited molecules (Figure 2). There are many commercially available photoinitiators activated by UV and visible light irradiations and some of them are listed in Table 1.

### **Insert Figure 2.**

### **Insert Table 1.**

The main classes of unimolecular photoinitiators are benzoin ethers, benzil ketals, acetophenones, oxime esters, aminoalkyl phenones, and acylphosphine oxides. Typical bimolecular photoinitiators include benzophenones, thioxanthenes, benzils, and quionones as photosensitizers with alcohols, amines, and thiols acting as hydrogen donors. In addition to conventional free radical polymerization, photosensitive compounds have also been used for photochemical initiation of controlled radical polymerization procedures. In this context, photoinitiators,

photosensitizers, photoredox catalysts play vital role in the process which will be detailed in the following sections.

### 1.2.2 Photoredox catalysts

Photoinduced electron transfer from transition metals (e.g. ruthenium and iridium complexes) and organic dyes (e.g. eosin Y, 9,10-dicyanoanthracene, and triphenylpyrylium salts) have recently received considerable attention, particularly for synthetic applications in a process called photoredox organocatalysis, Figure 3.[41] Due to their excellent light absorption properties (from 400 nm to 800 nm), these catalysts could be applied in various specific organic reactions, namely the formation of carbon-carbon bonds, including atom transfer radical addition to unactivated terminal alkenes, the reduction of organic halides, or the conversion of alcohols to the corresponding halides.[42]

#### Insert Figure 3.

These photoredox catalysts generate active species through single-electron transfer between excited state of a given metal complex or organic dye and an oxidative or reductive quencher under visible light irradiation. Upon light absorption, the photosensitizer (PS) attains a long-lived excited triplet state ( $PS^*$ ) that can be either reductively quenched ( $PS^{\cdot-}$ ) or oxidatively quenched ( $PS^{\cdot+}$ ) by a wide range of organic substrates. The formed radical ions ( $substrate^{\cdot+}$ ) are subsequently converted to the final products in a number of possible reactions (Figure 4).[43]

#### Insert Figure 4.

Photoredox catalysts are activated under very mild conditions, with most reactions proceeding at room temperature without the need for reactive radical initiators. They exhibit wide functional group tolerance and minimization of waste products.[41] These catalysts were initially based on ruthenium or iridium compounds because of their intense visible light absorption, long lived excited states and suitable redox potentials. In general, they were utilized in (i) photodynamic therapy, (ii) dye-sensitized solar cells and (iii) organic light-emitting diodes, and (iv) polymerization reactions. Heterogeneous semiconductors such as mesoporous carbon nitride and various metal oxides and sulfides were also used as visible light photoredox catalysts.[44] However, their activation mechanisms were more complex than transition metal-polypyridyl complexes.

Very recently, several reviews on photo-initiated/mediated/induced polymerization have been published,[45-48] but this review highlights the recent progress in the use of light to regulate controlled/living radical polymerization from mechanism, reaction, material preparation to advanced applications.

## **2. Photomediated ATRP**

### **2.1 Cu-based systems in the presence of photoinitiators**

The significant progress made in controlled radical polymerization techniques during the past few years has enabled the synthesis of well-defined polymers by photochemical means. Preliminary examples of successful photomediated CRP were based on integration of photosensitive compounds (photoinitiators, photosensitizers or photoredox catalysts) in well-known CRP systems. The role of photosensitive compounds was to convert the physical energy of incident light into chemical energy in the form of reactive radicals with high electron-donating properties. These radicals were capable of regenerating a lower oxidation state activator from a higher oxidation state deactivator (Figure 5).[49] Many UV and visible light sensitive compounds including unimolecular[50-53] and bimolecular free radical photoinitiators, [50, 54] dyes, [52] semiconducting nanoparticles[55-60] and metal carbonyls[61] were used to photochemically initiate and control the ATRP process (Table 2).

**Insert Figure 5.**

**Insert Table 2.**

Upon UV irradiation, unimolecular photoinitiators such as 2,2-dimethoxy-2-phenyl acetophenone or (2,4,6-trimethylbenzoyl) diphenylphosphine oxide undergo a unimolecular bond cleavage to generate free radicals, which could be used for the activation of an ATRP at room temperature. On the other hand, bimolecular photoinitiators such as benzophenone or camphorquinone, require a second compound called a co-initiator, including tertiary amines, thiols or alcohols, to produce such intermediates. Both reverse ATRP[62] and simultaneous reverse & normal initiation (SR&NI)[63] were investigated for photoinduced ATRP.[54] In the reverse ATRP system, the polymerization of methyl methacrylate (MMA) proceeded in the absence of an added alkyl halide, but with a poor control over molecular structure. The results also indicated that unimolecular photoinitiators were more suitable due to their mechanistic advantages including higher quantum yield and short lifetimes of transient steps, which prevented any quenching by the metal ions.

Addition of alkyl halide to the mixture to provide a SR&NI ATRP system, not only improved initiation efficiency but also produced well-defined polymers with more narrow molecular weight distribution (1.11~1.26).

Subsequently, a photoinduced ATRP was conducted in an inverse microemulsion media using hexane as the continuous phase and poly(oxyethylene) (3) oleyl ether and poly(oxyethylene) (6) oleyl ether surfactants along with aqueous oligo(ethylene glycol) monomethyl ether methacrylate solutions containing (poly(ethylene glycol) ( $M_n = 550$  g/mol) as a costabilizer), tris((2-pyridine)methyl)amine (TPMA) or *N,N,N',N'',N''*-pentamethyldiethylenetriamine (PMDETA) as ligand, with a water-soluble photoinitiator (Irgacure 2959) and an alkyl halide (poly(ethylene oxide) based macroinitiator) under UV irradiation.[53] Photochemically generated radicals from both photoinitiator and direct photolysis of the Cu(II)/ligand complex not only initiated the polymerization but also regenerated the Cu(I) activators, which could be lost due to termination reactions. The method was particularly advantageous for tuning the final particle size by switching on/off the UV light and adjusting, the fraction of the aqueous phase in the recipe. The UV light enabled to control growth of polymer chains, while the polymer molecular weights were determined by the size of the micelles.

Many attempts have been directed towards extending spectral sensitivity to longer wavelengths. The use of visible-light sensitive compounds such as photoinitiators, dyes, mesoporous carbon nitride and various metal oxides and sulfides were employed for the key step of regeneration of Cu(I) activator in an ATRP.[49] In one case, the excited dye molecule and hydrogen donor molecule such as amine underwent electron transfer reaction to form radical-cation/radical-anion pairs, which essentially yielded free radicals in the system.[52] The radicals produced from the hydrogen donors were able to activate the ATRP, but the other radicals generated from the dye molecules were not capable of reducing the Cu(II). They acted mostly as radical chain terminators or underwent other reactions leading to bleaching the dye. A unimolecular photoinitiator, (2,4,6-trimethylbenzoyl)diphenylphosphine oxide, could be also utilized for the visible-light induced ATRP and provided better control of the molecular characteristics under the same experimental conditions.[51] Considering their role in photoinduced click reactions and photodecomposition pathway, acylgermane photoinitiators were also expected to activate ATRP within the visible range.[64]

In another approach, the use of camphorquinone and benzhydrol as photosensitive compounds was investigated for visible light-induced ATRP.[54] In this system, the camphorquinone

absorbed visible light and abstracted a hydrogen from benzhydrol to form two ketyl radicals, which were inactive toward the monomer, but were able to reduce Cu(II) to Cu(I) species. In addition, the PMDETA ligand also could act as hydrogen donor and produced free radicals that were active in the ATRP equilibrium. The reverse ATRP process was poorly controlled and only polymers with broad molecular weight distributions were obtained in the absence of benzhydrol. The addition of benzhydrol to the process provided better control over molecular characteristics, forming polymers with narrow molecular weight distributions and good agreement between theoretical and experimental molecular weights. The SR&NI ATRP system was also investigated under similar conditions and was more successful than photoinduced reverse ATRP.

Zinc oxide[57], titanium dioxide[55, 60] and mesoporous carbon nitride[56] nanoparticles, are known light-sensitive compounds which absorb visible light and consequently released their electrons in the conductive band (Figure 6). These photochemically generated electrons could be efficiently used to modulate an ATRP process by regenerating Cu(I) activators from Cu(II) deactivators. Recently, the photocatalytic activity of TiO<sub>2</sub> or ZnO nanoparticles was improved by the addition of photosensitive compounds such as dyes, reduced graphene oxides and iron nanoparticles. The addition of these compounds not only extended the optical absorption characteristics of the system up to the visible range but also enhanced the rate of polymerization.[58] The process was also implemented for surface-initiated ATRP of various monomers, including sulfopropyl methacrylate potassium salt, 2-hydroxyethyl methacrylate, *N*-isopropylacrylamide and oligo(ethylene glycol) methacrylate in an aqueous environment. The visible light-initiated polymerization using TiO<sub>2</sub> nanoparticles in the presence of an initiator-modified substrate resulted in polymer brush growth in a controlled manner.[55, 60] The thickness of the polymer brushes increased linearly as a function of irradiation time. An especially attractive feature of this system was the heterogeneous nature of the nanoparticles, which made them easily separable and removable from the reaction media without causing any contamination in the ATRP process. Furthermore, this system did not require any expensive compounds, metals, and conventional radical initiators that might lead to side reactions. In a recent study, the visible light-induced ATRP was successfully mediated by benzyl alcohol grafted niobium(V) chloride (NbCl<sub>5</sub>) nanoparticles as a heterogeneous photocatalyst with ethyl 2-bromoisobutyrate as initiator led to the well-defined polymers with predicted molecular weights and narrow molecular weight distributions ( $M_w/M_n = 1.16\sim 1.30$ ). [59] The niobium complex nanoparticles could be reused as photocatalysts in five further cycles without significant loss of control or activity.

### **Insert Figure 6.**

Dimanganese decacarbonyl ( $\text{Mn}_2(\text{CO})_{10}$ ) is a photosensitive compound that can homolytically decompose to manganese pentacarbonyl ( $^{\bullet}\text{Mn}(\text{CO})_5$ ) when exposed visible light illumination or sunlight.[61] These radicals can not only abstract halogen atoms from alkyl halides to generate carbon centered radicals, but also regenerate the Cu(I) activators thereby allowing photoinduced ATRP of (meth)acrylates and styrene (Figure 7). The polymerization was instantly stopped in the absence of light. Subsequent re-exposure to visible light restored the growth of polymer chains with the same kinetic characteristics that were observed in the former light-on process. This versatile process was also applied to the synthesis of graft copolymers from commercially available poly(vinyl chloride) via the grafting from approach. The direct abstraction of halide atoms from poly(vinyl chloride) by ( $^{\bullet}\text{Mn}(\text{CO})_5$ ) radicals enabled the grafting of MMA from the polymer backbone. Successful formation of a graft copolymer was confirmed by gel permeation chromatography with monomodal molecular weight distribution and a significant shift of the peak value toward higher molecular weight.

### **Insert Figure 7.**

## **2.2 Cu-mediated system in the absence of photoinitiators**

### **2.2.1 Cu-mediated photoATRP**

Among the various polymerizations that can be grouped under the term photoATRP, copper-mediated polymerizations play one of the most important roles due to their close resemblance to the classical copper/amine ATRP reactions. In fact, the effect of light on classical ATRP was first mentioned in 2000.[65] A strongly enhanced rate of polymerization was then observed under the influence of light; however it took more than 10 years before this observation was further investigated. PhotoATRP via reduction of initially added copper(II) salt was reported in the presence of an excess of PMDETA.[21, 22] This methodology was further developed using PMDETA and TPMA derivatives, leading to a successful photoATRP with less than 100 ppm copper catalyst to monomer.[23, 24] A photoinduced copper-mediated ATRP was later reported for polymerization of acrylates in the presence of an excess of tris[2-(dimethylamino)ethyl]amine ( $\text{Me}_6\text{TREN}$ ) (Figure 8).[25, 66]

### **Insert Figure 8.**

Copper-mediated photoATRP is a versatile technique that has been used for the polymerization of many substituted acrylates; including methyl, [23, 25, 56, 66-75] ethyl,[23, 25, 71] *n*-butyl, [25, 67, 69, 70] *t*-butyl,[25, 70] ethylene glycol methyl ether,[25, 67, 71, 72] oligo(ethylene glycol) methyl ether,[25, 67, 68] di(ethylene glycol) ethyl ether, hydroxyethyl,[69, 70] hydroxypropyl,[69] lauryl,[69] stearyl,[69] isooctyl,[69] solketal,[67, 69, 71] and glycidyl acrylate.[69] Other monomers, such as *N*-isopropylacrylamide[60] [2-(acryloyloxy)ethyl]trimethylammonium chloride,[76] 3-(dimethylamino)propyl[76] and 2-ethylhexyl[76] acrylate were successfully polymerized in surface initiated reactions (see section 8).

Polymerization of methacrylates is more challenging in copper-mediated photoATRP, since fewer monomers were reported. They include methyl methacrylate[21-25, 56, 58, 74, 75, 77-80], 2-methacryloyloxyethyl phosphorylcholine[81], glycidyl methacrylate[81], and oligo(ethyl glycol) methyl ether methacrylate.[23] The polymerization of 2-hydroxyethyl,[55] 3-sulfopropyl,[55] 2-(dimethylamino)ethyl,[60] and *tert*-butyl methacrylate,[80] as well as [2-(methacryloyloxy)ethyl]trimethylammonium chloride,[55] and *N*-(2-hydroxypropyl)methacrylamide[82] were described but only for surface initiated reactions. Especially, a family of iminophenylboronate methacrylates was polymerized under photoATRP conditions at low temperature.[83] Styrene[25, 56] has also been polymerized via copper-photoATRP. Regarding the rate of polymerization, a correlation between light intensity, irradiation wavelength spectrum and copper/ligand concentrations exists for the respective monomers. The ligand and its specific molar ratio to CuBr<sub>2</sub> plays an important role.

### 2.2.2 Comparison of copper complexes

The structure of the copper complex Cu<sup>II</sup>(L<sub>n</sub>)X has a significant impact on the polymerization due to the solubility of the formed complex[21, 67, 68] and the potential participation of the ligand in the reduction process.[66, 84] An excess of amine ligand was often added to the reaction mixture.[22, 23, 25, 80] Nevertheless, copper catalyst concentrations in the range of and below 100 ppm were successfully used.[23-25] The lowest concentrations, below 1 ppm, were reported only for surface initiated ATRP (refer to section 8).[76, 82]

Most acrylate monomers were subjected to photoATRP using copper(II) bromide as the copper source and Me<sub>6</sub>TREN as ligand in dimethyl sulfoxide (DMSO).[25, 67, 69, 71-74] Generally, very good control over molecular weight distribution was achieved for various monomers, in the presence of different solvents and light sources. In almost all reports, an excess of 6 to 1 was used

for Me<sub>6</sub>TREN to copper(II). Ratios of copper to Me<sub>6</sub>TREN of 1:2 and 1:3 using ethyl  $\alpha$ -bromoisobutyrate (EBiB) for the polymerization of methyl acrylate (MA) in DMSO provided good control over the reaction.[25] Equimolar ratios usually resulted in no polymer formation within 90 min reaction time. This observation was also confirmed by a similar reaction that no polymerization occurred during the first 4 hours of irradiation, but that subsequent irradiation led to a slow well-controlled polymerization,[74] with lower excess ligand.[25]

Copper complex [Cu(Me<sub>6</sub>TREN)](O<sub>2</sub>CH)(ClO<sub>4</sub>) was reported as an efficient catalyst for photoATRP.[73] With this copper(II)-formate complex, near-quantitative conversion of methyl acrylate within 2 hours was reached, while maintaining excellent end-group fidelity and low dispersities even with only 100 ppm copper catalyst. Thereby, the long-term stability of the reported copper(II)-formate complex was underpinned by the possibility to chain extend the obtained polymers even after 6 months' storage of the complex solutions.

A [Cu(phen)<sub>2</sub>]Br (phen: 1,10-phenanthroline) photocatalyst was employed for visible light mediated photoATRP.[78] Excellent control was achieved with only 80 ppm catalyst, even if polymerization was slow due to the low light intensity of a 0.9 W LED ( $\lambda$ =465 nm) light source. [Cu(phen)<sub>2</sub>]<sup>+</sup> was not reactive enough to undergo thermal copper-mediated radical polymerization at room temperature, in contrast to a photoexcited form [Cu(phen)<sub>2</sub>]<sup>+\*</sup>. The reaction was probably mediated by an oxidative quenching of the photoexcited copper catalyst and not by an amine-based reductive quenching.[66, 84] In the presence of a large excess of triethylamine, the reaction was significantly accelerated but was still much slower than a CuBr<sub>2</sub>/Me<sub>6</sub>TREN based ATRP without excess ligand.[74]

In addition to copper(II) bromide based catalysts, copper(0), copper(I) bromide, and copper(II) chloride were used as the copper source.[25] Recently, the food supplement, copper(II) gluconate was used for photoATRP in the presence of additional Me<sub>6</sub>TREN ligand.[72] This finding underlined the importance of the copper(II)-Me<sub>6</sub>TREN complex as the deactivating mediator; incomplete substitution of gluconate with Me<sub>6</sub>TREN resulted in a significant loss of control, or even an uncontrolled radical polymerization. A slow ligand exchange required to premix copper(II) gluconate and Me<sub>6</sub>TREN prior to the reaction. This process could be accelerated when the mixture was heated or exposed to UV-light. A faster ligand exchange was observed when two equivalents of sodium bromide per equivalent of copper(II) were added, enabling direct photo polymerization without the requirement of premixing. Moreover, the presence of impurities in the

food supplement did not impair the reaction speed nor the end-group fidelity or polymer dispersity.

Next to Me<sub>6</sub>TREN, other amine ligands have also been investigated. Replacing Me<sub>6</sub>TREN with tris(2-aminoethyl)amine (TREN) in a polymerization of methyl acrylate with EBiB as initiator and only 400 ppm copper(II) bromide in DMSO led to only a minor difference in control. However, the use of PMDETA resulted in a significant reduction in both polymerization rate and control.[25] Other aliphatic ligands such as TPMA (Scheme 4) and tris((4-methoxy-3,5-dimethylpyridin-2-yl)methyl)amine (TPMA\*) were used in dimethylformamide (DMF).[23] While a significant reduction in polymerization rate was observed for both TPMA and TPMA\* ligands, the reactions maintained low dispersity with only 100 ppm copper. A comparable reaction with Me<sub>6</sub>TREN in DMSO was 10 times faster without compromising polymerization control.[25] In contrast to the success with aliphatic amines, the use of 2,2'-bipyridine (bpy) as a ligand did not result in any polymerization.

#### **Insert Scheme 4.**

PhotoATRP of methacrylates was investigated with various copper complexes. The most common ligands employed were TPMA [23, 24, 74] and PMEDTA (Scheme 4). [21-24, 56, 75, 77, 79] Me<sub>6</sub>TREN resulted in a loss of control and reduced polymerization rates.[25, 77] TPMA generally gave lower dispersities and faster polymerization rates compared to PMDETA,[23, 74] and the opposite effect was observed in anisole with copper concentrations above 100 ppm.[24] A reduction in polymerization rates and a loss of control was found for both 50 and 200 ppm CuBr<sub>2</sub>/PMDETA in anisole. Depending on the solvent and copper to ligand ratio a further reduction in copper concentration resulted in a loss of control and polymerization rate (Figure 9).[24, 74] Interestingly, equimolar ratio of TPMA to CuBr<sub>2</sub> gave very similar polymerization rates and molecular weight distributions.[74] Use of PMDETA as ligand works as well under equimolar conditions at 100 ppm CuBr<sub>2</sub>. [24, 74] Additionally, CuCl<sub>2</sub> was used with alkyl bromide initiators without compromising polymerization rate or control.[77] However, initiation with chlorinated initiators was generally too slow for photoATRP of methyl methacrylate.

#### **Insert Figure 9.**

In contrast to an ATRP with CuBr<sub>2</sub> as the copper(II) source, the polymerization of methyl methacrylate in DMSO with CuSO<sub>4</sub>·5H<sub>2</sub>O as the copper(II) source proceeded faster and provided

better control when TPMA was used in excess.[77] The use of an organic copper(II) salt for the polymerization of methyl methacrylate with 4 equivalents TPMA per copper(II) was further investigated, since these catalysts feature higher solubility in organic solvents.[77] While there were no significant differences between  $\text{CuBr}_2$  and organic copper(II) salts detectable for conversions above 60%, there were smaller differences for lower conversions, since the deactivator  $\text{CuBr}_2$  was generated in situ. The same behavior was also found for methyl acrylate.[66] The use of  $\text{CuBr}$  instead of  $\text{CuBr}_2$  resulted in almost identical results.[80] A very similar behavior was found when  $\text{CuO}$  was used instead of copper(II) salts.[74] When oxygen was present, the reaction started after an induction period which was strongly dependent on the amount of amine ligand. Subsequent to the induction period, the reaction followed first-order kinetics forming polymers with low dispersities.

### 2.2.3 Influence of the light source in photoATRP

Not only ligand and ligand concentrations, but also selection of light source has a very significant influence on photoATRP reactions. Since the absorption spectra of different copper complexes and reaction mixtures are generally different, the light source has to be selected accordingly.[23] In one of the earliest studies on photoATRP, various light sources were compared for the polymerization of methyl methacrylate using 100 ppm  $\text{CuBr}_2$  to monomer and TPMA\* as ligand in DMF (Figure 10).

#### Insert Figure 10.

Sunlight mediated and controlled the reaction most efficiently.[23] After 12 hours, the sunlight-mediated reaction yielded 81% conversion with  $M_n = 21,000$  and  $\mathcal{D} = 1.09$ . The same reaction mediated by LED ( $\lambda = 392$  nm;  $0.90 \pm 0.05$  mW/cm<sup>2</sup>) light required 27 hours to reach 71% conversion, and maintained slightly better control with  $\mathcal{D} = 1.05$ . In contrast, use of blue ( $\lambda = 450$  nm;  $10.0 \pm 0.5$  mW/cm<sup>2</sup>) and red ( $\lambda = 631$  nm;  $8.9 \pm 0.5$  mW/cm<sup>2</sup>) light resulted in only 26% and 0% conversion, respectively, after more than 20 hours, as these wavelengths lay mostly outside the absorption spectrum of the copper complex.

The use of light at various wavelengths was also investigated for the polymerization of methyl acrylate using  $\text{Me}_6\text{TREN}$  as ligand,  $\text{EBiB}$  as initiator, and copper(II) bromide as catalyst precursor at various concentrations (down to 100 ppm) in DMSO.[25] The photoATRP with UV light (360 nm;  $4 \times 9$  W nail curing device) gave a significantly faster polymerization and access to a broader range of copper to initiator ratios than sunlight. In agreement with previous

studies,[23] blue light resulted in a good polymerization control with slower polymerization rates.[25]

Generally, higher light intensities provide faster chain activation, and faster polymerization. However, high intensities also resulted in side reactions, providing poorer control. Important to note is that following Lambert-Beer law strong light intensity gradients exist in reaction vessels, leading to almost 100% depletion of UV light within a distance of few millimeters. Only by intensive stirring, all material can be brought into the illumination area of the vessel, and hence polymerization rate and the level of control also depends on the geometry of the reactor and the positioning of the lamps.[74]

Temporal control is one of the biggest advantages of photoreactions over thermal procedures. In principle, no reaction should occur in the absence of light and reactions should immediately stop when light irradiation is switched off. Temporal control is usually assessed by sequential light-on light-off experiments. Indeed, in most studies, a strongly reduced polymerization rate but not a full cessation was observed during the light-off periods (Figure 11).[21, 24, 25, 56, 61, 74, 79] Recently, a limited temporal control was observed during light-off periods.[73] Incomplete cessation of polymerization in dark time periods hints at formation of active Cu(I) species that persist and which then further activate dormant species.

### **Insert Figure 11.**

#### **2.2.4 Initiators in photoATRP**

In absence of an ATRP initiator, only low conversions and/or uncontrolled polymerizations were observed for photoATRP reactions.[22-25, 56, 66, 74, 79, 80] Higher initiator concentration lead to faster polymerizations.[66] EBiB was used in basically all photoATRP of acrylates, but also 2,3-dihydroxypropyl 2-bromo-2-methylpropionate and 2-(2-hydroxyethoxy)ethyl 2-bromo-2-methylpropionate initiated polymerizations successfully. Bifunctional initiators, such as ethylene bis(2-bromoisobutyrate and bis[2-(2'-bromoisobutyryloxy)ethyl]disulfide, were successfully used for the synthesis  $\alpha,\omega$ -telechelic block copolymers. However, increasing the number of initiation sites per molecule led to a higher degree of termination compared to monofunctional initiators.[67]

In contrast to the polymerization of acrylates, a variety of initiators were successfully employed for polymerization of methacrylates. Ethyl  $\alpha$ -bromophenylacetate (EBPA) and 2-bromopropanitrile (BPN) initiated the polymerization efficiently and yielded polymers with lower

dispersities compared to EBiB for the polymerization of methyl methacrylate.[23, 24] In accordance to the above described influence of CuCl<sub>2</sub> on photoATRP, chlorinated initiators led to slower polymerization and poorer control due to slow initiation.[77, 85]

### 2.2.5 Solvents in photoATRP

A sufficient solubility of the copper complex species in the polymerization medium is very important for an effective reaction to proceed, and also solvent polarity plays an important role in the success of polymerizations. For example, bulk methyl methacrylate polymerization did not show good control due to insufficient solubility of the copper complex.[69, 77] When methanol was added to the polymerization mixture, much better results were obtained.[21] Most photoATRP reactions of acrylates were carried out in DMSO (25-50 vol%),[25, 56, 66, 67, 69-75, 86] as well as in ionic liquids,[68] acetonitrile, [25, 56] DMF,[23, 25, 69] and other solvents such as toluene and alcohols have been reported.[25, 69] A variety of solvents were screened under similar conditions, and methanol was found to yield best results with near-quantitative conversion of monomer with  $\bar{D} = 1.05$  within 90 min,[25] whereas the use of other alcohols, i.e. ethanol and 2,2,2-trifluoroethanol only resulted in 38% and 54% conversion, respectively.[69] In contrast, the use of toluene as solvent resulted in a relatively uncontrolled polymerization with  $\bar{D} = 1.54$  at 62% conversion after 90 min.[25]

When water was used as solvent for the polymerization of oligo(ethylene glycol) methyl ether acrylate (OEGA<sub>480</sub>) poor control over the polymerization was observed,[69] while the same reaction proceeded well in DMSO. Nevertheless, a 1 to 1 mixture of water and DMSO provided polymerization rates similar to pure DMSO and a further increase towards pure DMSO improved the control. Various ionic liquids were also identified as excellent solvents for polymerization of methyl acrylate, enhancing the polymerization rate by factor 3 and concomitantly giving better control for lowest copper catalyst loadings.[68] Other monomers, or resulting polymers, showed poor solubility in at least some of these ionic liquids, limiting their utility for block copolymerizations. The use of 1-octyl-3-methylimidazolium hexafluorophosphate [C<sub>8</sub>mim][PF<sub>6</sub>] was beneficial for many monomers and allowed block copolymerization without compromising fast polymerization rates or dispersity.

Methyl methacrylate was predominantly polymerized in DMF,[23, 58, 75, 77, 79] and rarely in DMSO,[74, 77] anisole,[24, 74, 77] acetonitrile,[56] methanol,[21, 81], 1-butyl-3-methylimidazolium hexafluorophosphate [bmim][PF<sub>6</sub>],[79] or in the absence of solvent. [22, 56] PhotoATRP of MMA performed with BPN as initiator and 100 ppm Cu[TPMA]Br<sub>2</sub> in DMSO

was significantly faster than when the polymerization was performed in anisole under same conditions.[24, 74] Recently, photoATRP of oligo(ethylene oxide) methyl ether methacrylate (OEOMA) using  $\text{CuBr}_2/\text{TPMA}$  as catalyst in aqueous media was reported.[87] In this study, TPMA was selected as ligand due to formation of a stable Cu(I) complex that does not undergo significant disproportionation.[88] The polymerization was originally investigated using poly(ethylene oxide) bromophenylacetate ( $\text{PEO}_{2000}\text{BPA}$ ) as macroinitiator with a high concentration of Cu (20,000 ppm vs. monomer, Cu to initiator: 9:1) in very highly diluted solution (90% water). The conversion of OEOMA<sub>500</sub> reached 60% after 4 h of irradiation with visible light, providing polyOEOMA<sub>500</sub> with  $M_n = 97,000$ ,  $M_w/M_n = 1.07$ . Good control was maintained with Cu concentration as low as 220 ppm; however, the reaction with 110 or 22 ppm resulted in an uncontrolled radical polymerization. This was attributed to the dissociation of the halide from  $\text{CuBr}_2/\text{TPMA}$  complex, resulting in a low concentration of active deactivator, a major challenge for ATRP in aqueous media.[89, 90] This problem was resolved by adding 5 mM or 30 mM NaBr into the reaction mixture. The photoinduced ATRP of OEOMA<sub>500</sub> with 22 ppm Cu catalyst and 30 mM salt after 4 h gave 55% conversion, and formed a polymer with  $M_n = 103,000$  and  $M_w/M_n = 1.17$ .

### Insert Figure 12.

This system featured low concentrations of catalyst, as low as *ca.* 20 ppm molar ratio to monomer, or 0.3 mg/L copper amount in aqueous media with irradiation by visible light at room temperature, which could be an attractive procedure for preparation of water soluble copolymers for a broad range of applications. A similar system using  $\text{Me}_6\text{TREN}$  as a ligand in aqueous media was recently reported to polymerize acrylates. [91]

#### 2.2.6 Mechanism of Cu-mediated photoATRP

Several investigations were reported to elucidate the mechanism of photoATRP.[23, 25, 66, 84] In the early work, the activation rate of alkyl halides by  $\text{Cu}^I/\text{HMTETA}$  or TPMA catalyst was measured, and negligible difference was found in activation rate with or without irradiation. [23, 92] In contrast to the  $[\text{Cu}(\text{phen})_2]\text{Br}$  system (see section 2.2.1),[78]  $\text{CuBr}_2$  with TPMA or  $\text{Me}_6\text{TREN}$  was proposed to undergo a reductive quenching cycle, Scheme 5. No polymerization of MA was observed in the absence of free ligand or an appropriate electron donor such as trialkylamine, i.e., at  $\text{CuBr}_2:\text{Me}_6\text{TREN} = 1:1$ , whereas increasing the ratio of ligand to copper resulted in an increased rate of polymerization, Figure 13.[66]

### Insert Scheme 5.

### Insert Figure 13.

A series of experiments and kinetic simulations were conducted in which one or more components were removed from the standard model conditions. [66] It was determined that the major pathway to radical (re)generation was the photochemical reduction of Cu<sup>II</sup> complexes by free amine moieties, similar to an ARGET ATRP process. Small contributions from other processes such as the interaction between alkyl halides and ligands, ligands with monomers, and the photochemical cleavage of alkyl halides were still present. These observations were further supported by a more detailed investigation at a molecular level.[84] Pulsed-laser polymerization (PLP) was used to carry out the polymerization, and electrospray-ionization mass spectrometry (ESI-MS) was utilized to analyze the exact structure of the polymer and its chain end groups. In analogy to a previous report,[66] different combinations of photoATRP reaction components EBiB, CuBr<sub>2</sub>, Me<sub>6</sub>TREN, MMA and DMSO were investigated and led to a similar conclusion.[84]

PhotoATRP was also conducted with two fluorescent lamps ( $\lambda = 400\text{-}750\text{ nm}$ ) at a distance 10 cm from the reaction mixture.[80] This study raised one important question, whether photochemical processes influence various systems such as normal, ICAR and ARGET ATRP. ICAR ATRP polymerizations were conducted under ambient lighting or in the dark, and negligible differences were observed.[93] This suggested that under typical ICAR conditions, only chemical decomposition of AIBN contributes to the activator (re)generation. PhICAR (photochemical plus ICAR) ATRP was then investigated to clarify the effect of light on the activator (re)generation procedure. Previous photo-system used a light source at a distance of 10 cm, while the following reactions conducted the reaction at 1 m distance from the light source. At low concentration of 2,2'-azobis(2-methylpropionitrile), AIBN, 0.035 equiv. with respect to ATRP initiator, photochemical process in 392 nm photoreactor (0.9 mW/cm<sup>2</sup>) contributed *ca.* 50% to activator regeneration while at higher concentration of AIBN, 0.2 equiv with respect to ATRP initiator, the contribution dropped to *ca.* 15%.

The effects of ambient light on normal and ARGET ATRP were also investigated.[94] An enhancement in polymerization rate was observed with ambient light under normal ATRP conditions, but not under ARGET ATRP conditions. Under both conditions, excess amine ligand PMDETA was used. The rate enhancement in normal ATRP could be due to (re)generation of activator Cu(I) via photoreduction of Cu(II) by excess ligands. In ARGET ATRP, the reduction

of Cu(II) by chemical reducing agents such as ascorbic acid might be faster than photoreduction with excess amines. Therefore, the contribution of photochemical process is negligible, and no rate enhancement was observed in ARGET ATRP.

### 2.3 Fe-mediated photoATRP

PhotoATRP systems based on iron catalysts were reported in 2014 [95, 96] and 2015.[97-100] Iron could be considered as a preferred transition-metal for ATRP due to its abundance, lower toxicity, and lower cost than copper.[101, 102] Iron-mediated ATRP successfully polymerized various methacrylates and styrenes; however, polymerizations of acrylates were more challenging.[103-108] The Fe-mediated photoATRP polymerization of various methacrylates including methyl[97-100], ethyl[98, 100], n-butyl[98, 100], t-butyl[98], isobutyl[98], glycidyl[98], benzyl[100], 2-(trimethylsilyloxy)ethyl[100], and furfuryl methacrylate[100] were reported.

#### Insert Table 3.

PhICAR ATRP of MMA was conducted using  $\text{CCl}_4$  as ATRP initiator, tetramethylethylenediamine (TMEDA) as ligand and  $\text{FeCl}_3$  and AIBN as photoinitiators.[95] Benzoyl peroxide (BPO) was also successfully used as a conventional radical initiator under irradiation conditions. Typical features of controlled radical polymerization were observed in this system, such as linear semilogarithmic plots, linear relationships of  $M_n$  with conversion and low dispersity. Additionally, it was observed that  $\text{FeCl}_3/\text{bpy}$  complex was reduced by ethanol under irradiation at room temperature.[95] A series of Fe-mediated photoATRPs were conducted by using this photoreduction mechanism to provide well-defined PMMA. An induction period was usually observed at the early stage of the polymerization period, and polymerization completed stopped in the absence of irradiation.

#### Insert Figure 14.

Nanosized  $\alpha\text{-Fe}_2\text{O}_3$  was used as photoinitiator for the photoATRP of MMA with  $\text{FeCl}_3/\text{PPh}_3$  as catalyst in PEG as solvent.[97] The polymerization proceeded in a controlled manner; however the dispersities of the synthesized PMMAs ranged from 1.39 to 1.68. It was proposed that under irradiation, an electron ( $e_{\text{CB}}^-$ ) and a positive hole ( $h_{\text{VB}}^+$ ) were generated in the  $\alpha\text{-Fe}_2\text{O}_3$  nanoparticle, and Fe(III) was reduced to Fe(II) by the electron ( $e_{\text{CB}}^-$ ). Once Fe(II) was formed, an ATRP equilibrium could be established.

Photoredox catalysts were extensively used in organic and material chemistry.[41, 109-113] An iron based catalyst,  $\text{Fe}(\text{bpy})_3\text{Br}_2$ , was used to activate alkyl bromide to generate the radical in enantioselective organocatalytic alkylation of aldehydes.[114] A similar iron-based photoredox catalyst  $\text{Fe}(\text{bpy})_3(\text{PF}_6)_2$  was investigated for photoATRP of various methacrylates,[98] providing both linear semilogarithmic plot and a linear relationship of  $M_n$  with conversion. However, the molecular weights of the polymers obtained were ca. 10 times higher than theoretic values, indicating limited initiation efficiency.

Three iron-based photoredox catalysts were designed, synthesized with reduction potentials from -2.1 to -2.7 V for the excited states of the catalysts.[99] With these reduction potentials, alkyl bromides such as EBPA could be reduced to radicals, initiating an ATRP. The “living” nature of this system was confirmed by linear kinetic plot, linear correlation between  $M_n$  and conversion, and clean chain extension. The broad molecular weight distribution ( $M_w/M_n = 1.5 - 2.2$ ) for this system could be due to photo-degradation of iron photoredox complexes during polymerizations.

Polar solvents such as DMF and MeCN are efficient ligands for  $\text{FeBr}_2$ -catalyzed ATRP of MMA.[108] Based on this observation, a photoinduced Fe-mediated ATRP of methacrylates using only air-stable  $\text{FeBr}_3$  without additional ligands, reducing agents or radical initiators was reported.[100] Simple irradiation of a reaction mixture of methacrylate,  $\text{FeBr}_3$  and EBPA in MeCN provided well-defined polymers with high bromine chain-end functionality. A mechanistic investigation showed that irradiation resulted not only in the photoreduction of Fe(III) to Fe(II) in presence of excess monomer (Figure 15) but also in photoactivation of Fe(II) activators. Since the activation required irradiation, full temporal control with no detectable monomer conversion was observed when the irradiation was turned off.

**Insert Figure 15.**

#### **2.4 Ir-mediated photoATRP**

Inspired by powerful organic transformations reported by MacMillan,[41, 115] Yoon,[42, 116] and Stephenson,[117, 118] ATRP was successfully photocatalyzed using the most commonly used photoredox catalyst, *fac*- $[\text{Ir}(\text{ppy})_3]$  (ppy: 2,2'-phenylpyridine).[119] Initially, the polymerization of MMA was investigated using EBPA as conventional ATRP initiator and 0.2 mol% *fac*- $[\text{Ir}(\text{ppy})_3]$  as catalyst under visible light irradiation which resulted in an uncontrolled polymerization. This indicated that the photoredox catalyst was only involved in activating dormant species rather than deactivating the propagating radicals. This was also previously

observed in photoinduced uncontrolled radical and ring-opening polymerizations using Ir photoredox complexes.[120] Interestingly, decreasing the loading of the Ir complex catalyst to 0.005 mol% resulted in excellent control over the polymerization, providing PMMA with  $M_n = 22,900$  and  $M_w/M_n = 1.25$ . Another interesting property of this system was polymerization of methacrylic acid (MAA), which is one of the most challenging monomers to be polymerized under ATRP conditions due to a carboxylic acid functional group. Other methacrylates such as ethylene glycol methyl ether (PEGMA),[121] 1,1,1-trifluoroethyl (TFEMA),[122] and propargyl methacrylate (PMA)[123] were also polymerized in a controlled fashion under Ir-catalyzed photoATRP conditions.

**Insert Figure 16.**

Polymerization proceeded only under irradiation and in the presence of *fac*-[Ir(ppy)<sub>3</sub>], and the reaction completely stopped by turning off the light, indicating that the excited state of Ir catalyst activated the alkyl bromide.[119] The simplified proposed mechanism for photoredox catalyst Ir-catalyzed ATRP is shown in Scheme 6. A very strong reductant *fac*-[Ir(ppy)<sub>3</sub>]\* ( $E^\ominus = -1.73$  V vs SCE) is formed under irradiation by visible light. This excited Ir<sup>III\*</sup> species would reduce an alkyl bromide or polymer with bromine as chain end functionality to generate the propagating radical. The radicals then added monomers before being deactivated by the oxidizing Ir<sup>IV</sup> complex. The deactivation reaction between Ir<sup>IV</sup> complex and radicals closed the catalytic cycle to regenerate the ground-state Ir<sup>III</sup> catalyst and dormant polymer chain with bromine chain end, which was important for a controlled radical polymerization.

**Insert Scheme 6.**

**Insert Table 4.**

Ir-catalyzed photoATRP was then extended to the polymerization of various acrylates including methyl, ethyl, n-butyl, t-butyl acrylates as well as acrylic acid in the presence of benzyl  $\alpha$ -bromoisobutyrate (BnBiB).[124] Random copolymerization of acrylic acid (AA) and ethyl acrylate (EA) was prepared at different molar ratios, up to 50 mol% AA. In the polymerization with low ratios of AA (5 – 20 mol%), copolymers with low dispersities, between 1.25 and 1.40 were achieved but higher dispersities (1.51 – 1.90) were observed with higher ratios of AA (30 – 50 mol%). Compared to the polymerization of MAA under Ir-catalyzed photoATRP conditions,

control over the copolymerization of AA was more difficult to achieve. This could be related to the displacement of the terminal halide by a pendant nucleophilic carboxylate.[125]

A new iridium complex Ir(btp)<sub>2</sub>(tmd) (btp: 2-(2'-benzothienyl)pyridine; tmd: 2,2,6,6-tetramethyl-3,5-heptanedione) was designed and investigated for radical or cationic controlled polymerizations under mild irradiation ( $\lambda$ : 457 – 532 nm).[126] Compared to *fac*-[Ir(ppy)<sub>3</sub>], Ir(btp)<sub>2</sub>(tmd) catalyst complex exhibited a stronger absorption in visible light range, and the luminescence lifetimes of newly designed complex Ir(btp)<sub>2</sub>(tmd) was longer than *fac*-[Ir(ppy)<sub>3</sub>] (4400 ns vs. 1300 ns). Consistent with previous observations, the photoATRP with higher loading of Ir catalyst, 1.5% mol ratio to initiator, resulted in an uncontrolled radical polymerization of MMA, but simply reducing the amount of catalyst to 0.14 mol% ratio to initiator resulted in much better controlled polymerization, providing linear relationship between  $M_n$  and conversion as well as polymers with low dispersities (1.2 - 1.3).

## 2.5 Ru-mediated photoATRP

Another photoredox catalyst tris(bipyridine)ruthenium(II) chloride (Ru(bpy)<sub>3</sub>Cl<sub>2</sub>, Figure 17) was recognized as a photosensitizer capable of initiating a free radical polymerization in the presence of a sacrificial electron donor trimethylamine in 1985.[127] A combination of Ru(bpy)<sub>3</sub>Cl<sub>2</sub> and *N,N*-diisopropylethylamine was used to initiate radical polymerization in the presence of EBiB under visible light irradiation in 2011.[128] This system efficiently polymerized various methacrylates, including MMA, EMA, nBMA, tBMA, iBMA, GMA, and 2-ethoxyethyl (EEMA) methacrylate. However, polymerization of styrene (St) and acrylates using this system was less successful.

### Insert Figure 17.

At the same time, a cyclometalated ruthenium (II) complex was reported to catalyze photoATRP utilizing a different photo-activation mechanism.[129] A series of cyclometalated ruthenium (II) compounds were synthesized, and both *cis*-[Ru(*o*-C<sub>6</sub>H<sub>4</sub>-2-py)(phen)(MeCN)<sub>2</sub>]PF<sub>6</sub> (**RuI**, Figure 17) and *cis*-[Ru(*o*-C<sub>6</sub>H<sub>4</sub>-2-py)(bpy)-(MeCN)<sub>2</sub>]PF<sub>6</sub> (**RuII**, Figure 17) were photosensitive.[130] Both compounds were inactive under thermal conditions. However, visible light irradiation of complex **I** in methanol resulted in the dissociation of one acetonitrile ligand generating a reactive 16-electron Ru intermediate **I'**. The Ru complex **I'** with one vacant site activated an alkyl halide to produce a carbon-centered radical and a 17 electron complex **II**, which acted as an efficient deactivator in a classic ATRP equilibrium (Scheme 7), allowing polymerization of typical

monomers, such as MMA, BA and St, with good control. Almost linear semilogarithmic kinetic plots for each monomer indicated a constant concentration of radicals during the polymerization under irradiation. Although the initiation efficiency was not quantitative, the  $M_n$  increased linearly with the monomer conversion. PMMA with a low value for  $M_w/M_n$  (1.22) was synthesized, but polymerization of *n*-butyl acrylate (BA) and St resulted in formation of polymers with higher dispersity,  $M_w/M_n$  1.5 – 1.6.

### Insert Scheme 7.

## 2.6 Au-mediated photoATRP

Recently, a photoATRP system based on a dinuclear gold (I) photoredox catalyst  $[\text{Au}_2(\mu\text{-dppm})_2]\text{Cl}_2$  (dppm: (diphenylphosphino)methane) was reported.[131] Controlled radical polymerizations of methacrylates and acrylates were achieved in the presence of EBPA as the conventional ATRP initiator and low loadings, 1.25 mol% ratio of Au(I) catalyst to initiator, under various irradiation conditions. Similar to the Ir-mediated photoATRP (Scheme 6), an excited state  $\text{Au}_2^{\text{II}*}$  complex ( $E^\ominus = -1.6$  V vs SCE) was photogenerated, which reduced a C-Br bond at the polymer chain end or in a low molecular weight initiator to provide a carbon-centered radical. This activation process was further supported by a rate constant of  $9.2 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$  for the reaction between  $\text{Au}_2^{\text{II}*}$  and EBPA by laser flash photolysis (LFP) measurements. Additionally, the oxidized gold complex  $[\text{Au}_2(\mu\text{-dppm})_2\text{Cl}_2]^+\text{Br}^-$ , detected by UV-vis spectra, quenched the propagating radical to regenerate the bromine chain end and ground-stated  $\text{Au}_2^{\text{II}}$  complex. The “living” nature of this process was confirmed by the predictable  $M_n$  and retention of chain end functionality (CEF). However, the molecular weight distributions were broad (1.62 - 1.85). This indicated that the deactivation was not very efficient and some side reactions competed with the main reaction pathways.

## 2.7 Metal-free ATRP

### 2.7.1 Polymerization of methacrylates

Inspired by the success of transition metal-based photoredox catalysts, a series of organic based small molecules were examined as catalysts for photoATRP, due to their intrinsic photoredox properties. 10-Phenylphenothiazine (PTZ, Figure 18) was reported to successfully polymerize methacrylates in a controlled manner with traditional ATRP initiator using 380 nm irradiation.[26] When the reaction was conducted in the absence of irradiation, initiator or catalyst, uncontrolled radical polymerization was observed. The PTZ-mediated photoATRP could be temporally

controlled, with no monomer conversion in the absence of irradiation, Figure 19a. The molecular weight increased linearly with conversion, also with several “on-off” irradiation cycles, Figure 19b. The linear semilogarithmic kinetic plot indicated a controlled radical polymerization process, Figure 19c. Interestingly, 2-(dimethylamino)ethyl methacrylate (DMAEMA) could be polymerized with precise control over molecular weight generating polymers with low dispersity (~1.1) under PTZ-mediated photoATRP conditions. In contrast, Ir-based photoATRP of DMAEMA only provided an uncontrolled polymerization due to formation of a radical cation from the amine moiety in the monomer by the oxidized Ir<sup>IV</sup> complex.

**Insert Figure 18.**

**Insert Figure 19.**

Perylene was also reported as an organic photoredox catalyst for a metal-free photoATRP of methyl methacrylate and other functionalized vinyl monomers.[132] Perylene behaved as an activator to reduce a conventional ATRP initiator, such as EBPA, to induce radical polymerization under a visible light irradiation. However, limited initiation efficiency was determined from 2% to 40% in various solvents when catalytic amounts, 0.11 equiv to initiator, of perylene were used. Although a linear semilogarithmic kinetic plot and good temporal control by light was observed, the molecular weight decreased with conversion. This indicated that the system catalyzed by perylene partially resembled an uncontrolled radical polymerization. This could be attributed to the fact that oxidized perylene radical cation did not efficiently quench the radicals. Retention of the bromine chain end functionality was limited, which was confirmed by several chain-extension polymerizations and MALDI-TOF. This might arise because of bromine abstraction from EBPA by the propagating radical as well as other undesired radical terminations, rather than a clean deactivation pathway.

Fluorescein was recently used as metal-free catalyst for photoATRP of methyl methacrylate.[133] In the presence of trimethylamine as a sacrificial electron donor, fluorescein was involved in a reductive quenching cycle to activate the alkyl bromide and generate a propagating radical. In contrast, both PTZ and perylene-catalyzed photoATRP systems were proposed to undergo an oxidative quenching cycle, and they did not require a sacrificial electron donor. A more detailed discussion is presented in section 2.7.3. Model experiments clearly showed that control over polymerization could only be obtained in the presence of fluorescein, trimethylamine, EBPA and irradiation. A polymerization with fluorescein but without trimethylamine resulted in formation

of a trace amount of uncontrolled polymer (3% monomer conversion,  $M_n = 62100$ ,  $M_w/M_n = 1.92$ ). The fluorescein-catalyzed photoATRP of MMA was investigated with various amounts of catalyst (10%, 20%, and 30% mol ratio to initiator) and trimethylamine, Figure 20. All three experiments gave almost linear first-order kinetic plots, and the polymerization with higher loading of fluorescein gave slower polymerization rate. The increasing molecular weight with conversion supported a controlled polymerization process. However, the molecular weights for obtained PMMA were much higher than the theoretical values, indicating limited initiation efficiency. Interestingly, the initiation efficiency was significantly improved by targeting higher degrees of polymerization (DP). For example, when targeting a DP of 800, 86% initiation efficiency was obtained at 39% monomer conversion. This system fulfills some criteria for a controlled polymerization process such as linear kinetic plot, increasing molecular weight and high chain end functionality. However, this system was inferior to PTZ-catalyzed photoATRP due to inherent limitations in initiation and deactivation. Very recently, a visible light mediated metal-free ATRP was developed using 5,10-diaryl-5,10-dihydrophenazine derivatives (Figure 18). [134]

**Insert Figure 20.**

### **2.7.2 Polymerization of acrylonitrile**

The PTZ-catalyzed photoATRP system was also applied to the synthesis of polyacrylonitrile and produced polymers with predictable molecular weights and relatively low dispersities.[27] In addition to 10-phenylphenothiazine, 10-(4-methoxyphenyl)- and 10-(1-naphthalenyl)-phenothiazine were synthesized and successfully applied to the polymerization of acrylonitrile. PhotoATRP of acrylonitrile was conducted with EBPA as ATRP initiator and catalytic amounts of the PTZ derivatives, 10 mol% to initiator. Linear semilogarithmic kinetic plots were observed for acrylonitrile polymerizations in both DMSO and DMF, while the polymerization in DMSO gave a faster polymerization than in DMF (Figure 21a). The molecular weights increased with conversion, but at the beginning of the polymerization the molecular weight was higher than the theoretical value, indicating slow initiation (Figure 21b). 2-Bromopropionitrile was also used as an ATRP initiator instead of EBPA for the polymerization of acrylonitrile under PTZ-catalyzed photoATRP conditions. However, the initiation efficiency was not improved. High bromine chain end functionality was confirmed by  $^1\text{H}$  NMR spectra and successful chain extension of polyacrylonitrile macroinitiator. This is an especially viable method to synthesize polyacrylonitrile as a precursor for carbons for electronic applications since it contains no

residual transition metal.[135-138] On the other hand, Cu-mediated ATRP of acrylonitrile usually provided polymers with very narrow MW distribution ( $M_w/M_n < 1.10$ ).[139, 140]

**Insert Figure 21.**

### 2.7.3 Mechanism of metal-free ATRP

An extensive investigation recently elucidated the mechanism of metal free PTZ-catalyzed photoATRP, especially as a comparison with Cu-mediated ATRP.[141] A series of phenothiazine derivatives, including phenothiazine, 10-methylphenothiazine 10-(4-chlorophenyl)-phenothiazine and phenyl benzo[*b*]phenothiazine and 9-phenylcarbazole and thianthrene were tested in metal-free photoATRP in order to investigate and identify structure-reactivity relationships. Successful polymerization of MMA was conducted with all selected catalysts in the presence of EBiB as the ATRP initiator. All compounds efficiently photoactivated the alkyl bromide to generate radicals which initiated the polymerizations. However, only N-aryl phenothiazine mediated polymerizations gave polymers with controlled molecular weights and relatively narrow molecular weight distributions, indicating that the other compounds were not efficiently involved in the deactivation process.

Upon irradiation, excited *fac*-[Ir(ppy)<sub>3</sub>]\* and Ph-PTZ\* show strong reduction ability, -1.73 V and -2.10 V vs SCE, respectively (Figure 22). The most active catalyst for Cu-mediated ATRP is [Cu<sup>I</sup>(TPMA\*3)]<sup>+</sup> ( $E^\ominus = -0.18$  V vs SCE, TPMA\*3 = tris((4-methoxy-3,5-dimethylpyridin-2-yl)methyl)-amine; note that redox potential of [CuBr<sub>2</sub>(TPMA\*3)] is more negative -0.42 mV due to a strong halidophilicity of Cu(II) species). [142] These values suggested that both *fac*-[Ir(ppy)<sub>3</sub>] and Ph-PTZ have much greater reactivity for the activation of alkyl halides via an outer-sphere electron transfer (OSET) processes rather than typical inner-sphere electron transfer (ISET) in Cu-based ATRP[143, 144] The activation step was studied by laser flash photolysis (LFP) experiments. Different concentration of EBPA solutions was reacted with excited Ph-PTZ\*, resulting in different lifetimes. By plotting the concentration and lifetime to Stern-Volmer equation, a rate constant  $5.7 \times 10^9 \text{ M}^{-1}\text{s}^{-1}$  was obtained for the reaction between EBPA and Ph-PTZ\*.

**Insert Figure 22.**

The OSET activation step of a metal-free ATRP was subsequently carefully evaluated by modified Marcus theory.[145-148] Based on calculations, the ground state Ph-PTZ could not

activate methyl 2-bromoisobutyrate (MBiB,  $k_{\text{act,OSET}} = 1.0 \times 10^{-14} \text{ M}^{-1}\text{s}^{-1}$ ), while Ph-PTZ\* reacted with MBiB with a very large rate constant ( $k_{\text{act,OSET}} = 5.8 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$ ). This value could not be directly compared to the activation rate constant of a traditional Cu-catalyzed ATRP due to the low quantum yield ( $\phi = 1.3 \times 10^{-3}$ ) for Ph-PTZ\*. Therefore the effective rate constant was significantly decreased by the low lifetime of the excited state and the quantum yield.

The triplet excited state of *fac*-[Ir(ppy)<sub>3</sub>] has a much longer lifetime (1900 ns) compared to Ph-PTZ (4.5 ns), which resulted in much higher quantum yield for *fac*-[Ir(ppy)<sub>3</sub>] ( $\phi = 0.4$ ). This could also explain why photoATRP of MMA required only 50-100 ppm of *fac*-[Ir(ppy)<sub>3</sub>], but 1000 ppm of Ph-PTZ. A much higher proportion of the Ir complex would be involved in both activation and deactivation, while most of formed Ph-PTZ\* would quickly decay back to the ground state, being unable to efficiently participate in the activation and deactivation processes.

Another interesting observation in this study was that a controlled polymerization was achieved only for compounds that exhibited a reversible cyclic voltammetry. Irreversible redox behavior was detected for 9-phenylcarbazole and thianthrene (Figure 23), indicating that the radical cation derivatives were not stable enough to deactivate the propagating radical during a polymerization. The deactivation pathways were then investigated by both Marcus and DFT theory calculations. Five possible deactivation mechanism of MiB (methyl isobutyrate) radicals were proposed, (Scheme 8): (a) the ISET mechanism through a concerted Br atom transfer; b) the outer-sphere dissociative electron transfer (DET) between MiB• to Ph-PTZ<sup>•+</sup>Br<sup>-</sup>, providing dissociation form of catalyst and Br<sup>-</sup>; (c) the stepwise outer-sphere electron transfer (OSET-I) from MiB• to Ph-PTZ<sup>•+</sup>Br<sup>-</sup> to form an MiB<sup>+</sup> carbocation; (d) the outer-sphere electron transfer between MiB• and the dissociated Ph-PTZ<sup>•+</sup> (OSET-II); and (e) the termolecular outer-sphere associative ET (AET-ter) from Ph-PTZ<sup>•+</sup> to MiB• and Br<sup>-</sup> to form the ground-state catalyst **2** and MBiB.

### Insert Figure 23.

Among the five possible pathways, the ISET pathway required an activation free energy of 10.5 kcal/mol. Both DET and OSET-II involved the formation of carbocation had activation free energies of 13.4 and 9.2 kcal/mol, respectively, while OSET-I had the highest barrier, 13.6 kcal/mol. It was concluded that AET-ter had the lowest activation free energy, 3.9 kcal/mol, indicating that it was the most favorable pathway, consistent with experimental observations. Metal-free photoATRP with an alkyl chloride initiator was not controlled, ruling out the possibility of OSET-II. Additionally, the fact that there was no formation of oligomers implied

no carbocation intermediates were formed, which excluded DET, OSET-I and OSET-II pathways. Consequently, the most favorable AET-ter pathway was further supported by the calculated reaction rates. The rate of AET-ter deactivation was calculated to be ~200 times faster than the rate of radical-radical termination, while rates of other deactivation pathways (ISET, DET, and OSET-I) were much slower, or in the same order of magnitude (OSET-II) as termination.

#### **Insert Scheme 8.**

By combining all the information summarized above, an overall mechanism was proposed as an oxidative quenching cycle for PTZ-catalyzed photoATRP (Scheme 9). Excitation of ground state of Ph-PTZ by irradiation provided the excited Ph-PTZ\*, which had sufficient reduction potential to reduce alkyl bromide in a dissociative electron transfer process to form a carbon-centered radical. At the same time, neutral catalyst Ph-PTZ was oxidized to the radical cation form, Ph-PTZ<sup>•+</sup>, which was in equilibrium with the ion pair Ph-PTZ<sup>•+</sup>Br<sup>-</sup>. The radicals initiated polymerizations before being deactivated by Ph-PTZ<sup>•+</sup> in a termolecular associative electron transfer, which closed the catalytic cycle to reform the catalyst and polymer chain with bromine as chain end.

#### **Insert Scheme 9.**

Several side reactions were also possible in this metal free PTZ-catalyzed photoATRP. A background reaction photoinduced homolytic cleavage of the C-Br bond would provide carbon centered and bromine radicals which could initiate a free radical polymerization. The carbon centered radical could abstract bromine from ATRP initiator or be deactivated by Ph-PTZ<sup>•+</sup> to form an active polymer chain. Although Ph-PTZ\* had the ability to reduce the radical to a carbanion, this process should be very slow due to the low concentrations of these two active intermediates in the reaction medium.

Unlike PTZ-catalyzed photoATRP, metal-free photoATRP based on fluorescein was proposed to undergo a reductive quenching cycle (Scheme 10). [133] Upon irradiation by visible light, an excited species of fluorescein was formed and quenched by excess electron donor trimethylamine to provide the radical anion form of fluorescein and an amine radical cation. Then, the radical anion form of fluorescein reduced the C-Br bond in the ATRP initiator, or polymer chain end, to generate the propagating radical and regenerate the catalyst in ground state. It was proposed that trialkylamine radical cation oxidized the bromide anion to bromine radical which quenched the

propagating radical to form a polymer with a bromine chain end. The proposed deactivation pathway was a pure radical-radical coupling reaction, which should have very slow rate. On the other hand, the deactivation reaction between  $\text{Br}^\bullet$  and  $\text{P}_n^\bullet$  competed with bromine radical-initiated polymerization. This led to an inefficient deactivation procedure which caused the molecular weight of the formed polymers to be much higher than theoretical values as well as displaying broad molecular weight distributions.

**Insert Scheme 10.**

## 2.8 Applications

High level of control over block copolymer formation provided access to responsive materials that can be utilized in complex environments. Full sequential control, thus close to 100% end group fidelity, allows for the formation of multiblock copolymers, permitting to mimic the primary structure and potential function of simple proteins. The photoATRP procedures featured herein provide excellent “living” characteristics and have already been employed for formation of multiblock copolymers. An amphiphilic diblock copolymer based on a block of poly(methyl acrylate) and one of poly(solketal acrylate) that was subsequently hydrolyzed formed nanoparticles in a self-assembling process and were characterized via transmission electron microscopy (TEM) and dynamic light scattering (DLS) in water (Figure 24). [69]

**Insert Figure 24.**

The sequential synthesis of a decablock copolymer with a dispersity of only 1.16 and a molecular weight of 8,500 was reported.[70] Methyl, di(ethylene glycol) ethyl ether and *t*-butyl acrylate were polymerized in a sequential procedure adding neither copper(II) nor ligand in between chain extension steps. Since only a  $\text{DP}_n = 13$  was targeted, reaction times were kept between 2 and 7 hours. Increasing reaction times was required for an increasing number of blocks. Since multiblock copolymer synthesis via photoATRP relies on the presence of a bromine, or a (pseudo)halide functionality, monomer must always be present to keep radical-radical coupling reactions at a minimum.[149] Therefore, monomer conversion was monitored via Fourier transform infrared spectroscopy (FTIR) to keep radical coupling reactions to a minimum, nevertheless a clearly visible shoulder appeared in the SEC graph with increasing number of blocks (Figure 25). Hydrolysis of the incorporated *t*-butyl acrylate units led to formation of a pH and temperature responsive polymer with self-assembly properties. The hydrodynamic radii of

the formed decablock copolymer particles at pH = 6 changed with temperature due to thermoresponsive poly(di(ethylene glycol) ethyl ether acrylate) blocks.

**Insert Figure 25.**

High molecular weight multiblock polyacrylates were also investigated.[71] Blocks of 3, 10, 25 and 100 repeating units were targeted. With increasing target  $DP_n$  per block, a decrease in both polymerization rate and end group fidelity was observed, limiting the number of blocks for the multiblock copolymer targeting segments with  $DP_n = 100$  to 6, whereas  $DP_n = 3$  allowed 11 blocks with comparable dispersity of around 1.2, in agreement with results obtained for other controlled polymerization techniques. In contrast to the previous approach, fresh  $CuBr_2/Me_6TREN$  in DMSO was added along with every 2<sup>nd</sup> ( $DP_n=25$  and 100) or 3<sup>rd</sup> monomer addition, respectively, improving the degree of control as well as polymerization rates. Reported reaction times were longer in these reactions, due to a different light source and lower polymerization temperature.[67]

Recently, bifunctional initiators were used for the synthesis of  $\alpha,\omega$ -telechelic multiblock copolymers.[67] The use of bifunctional initiators allowed for the synthesis of copolymers with up to 19 blocks, with a bimodal molecular weight distribution ( $D = 1.45$ ), indicating more termination than with the monofunctional analogues since the number of initiating sides was doubled. Although no temperature dependence was found for the synthesis of the homopolymers, the degree of control was dependent on the temperature, with lower temperatures giving better control. Consequently, an optimized protocol allowed for the formation of a 23-block copolymer (11 block extensions) with target  $DP_n = 2$  and a dispersity of only 1.18. The optimized protocols also allowed the preparation of undecablock copolymers (target  $DP_n = 100$ ) with 150,000,  $D = 1.20$ . Despite the good agreement between theoretical and actual molar mass, a clearly visible shoulder (~30% peak height) was observed in the SEC trace. The symmetric nature of the telechelic polymers was demonstrated via SEC after cleavage of a polymer formed with a bifunctional initiator with a disulfide linkage, yielding a polymer with half of the initial molecular weight. A multiblock star polymer up to 21 arms was recently synthesized via photoinduced Cu-based ATRP. [150]

Sequence-defined polymerizations were one step beyond sequence-control, providing a precise number of monomer units per block, i.e. the precise equivalent to proteins or DNA. Sequence-defined polymers could potentially be used to encode information.[151] Recently, the synthesis of

sequence-defined acrylate oligomers was reported using photoinduced copper-mediated single unit monomer insertions.[152, 153] Starting from EBiB, methyl acrylate and up to 4 other acrylate monomer units were sequentially inserted into the copolymer. The conversion was monitored in situ via FTIR spectroscopy. After each monomer insertion step, the desired products were isolated via preparative recycling SEC. Interestingly, chlorine end groups were identified as unreactive side products after purification, which could be linked to presence of traces of chlorinated solvents during the photoreaction.

Synthesis of molecularly imprinted polymer (MIP) nanoparticles was reported, which recognized testosterone and  $\beta$ -propranolol, based on photoATRP processes using an iridium catalyst.[154] The MIPs were crosslinked polymer matrices formed in the presence of a template molecule. The template molecules were removed after the polymerization and left a cavity that preferentially binds the template molecule. PhotoATRP was a beneficial procedure for MIP formation as it could be performed at low temperatures. The controlled nature of the network formation enhanced molecular recognition by providing more homogenous recognition sites. Iridium catalysts were chosen to ensure compatibility with acidic monomers, which were required to reach good imprinting.

One of the biggest constraints of photoreactions is the low illumination efficiency arising from the optical density of the reaction mixture. Efficient exposure to light is only given for the outermost layers of a reaction container, as strong photo-gradients were imposed by Beer-Lambert's law. To overcome this limitation, photoreactions can be carried out in continuous flow reactors that provided shorter path lengths compared to batch reactors, [155] and which provide scalability of polymerizations by increasing tube lengths and parallelization of reactors. Homogeneous reaction conditions as well as scalability and more efficient heat release are some additional advantages of flow reactors.[156] The significant impact of efficient illumination in flow reactors is best visualized by the speed-up of otherwise very slow [2+2] cycloadditions, which can be accelerated from several hours to a minute reaction time when switching from batch to flow operation.[157, 158] PhotoATRP in a continuous flow reactor for methyl and *n*-butyl acrylate was reported with CuBr<sub>2</sub>/6Me<sub>6</sub>TREN in DMSO.[86] The reaction was carried out at 15 °C in microreactors. Good control over the molecular weight distribution was obtained along with a speed-up by a factor of 4 compared to (small scale) batch systems. Very high end group fidelity was demonstrated by ESI-MS and underpinned by block copolymer formation. More recently, MMA was polymerized in a flow reactor. PMDETA was selected as ligand in order to

attain good block structures.[75] In this case the reaction was around 7 times faster than in batch. The strongest increase in polymerization rate of a photopolymerization reaction so far was observed for Co-mediated photoCMRP.[159] Remarkably, control was also maintained for UV irradiation in flow above room temperature, whereas crosslinked polymers were found for the same reaction in batch systems. Recently, Ir-mediated flow photoATRP was reported for the synthesis of PMMA, demonstrating similar advantages for flow reactions over batch processing.[160] As mentioned above, rate increase and avoidance of side reactions caused by too high light intensities directly at the reactor wall, is one advantage of using flow reactors. However, highest potential is the simple scalability of the reactions. While batch photoreactions become increasingly inefficient when increasing reactor dimensions, flow reactors are easily scaled up while maintaining short optical pathlengths. Flow reactors are hence ideal to produce larger amounts of material compared to batch, where often only a few grams are synthesized.

### **3. Photomediated RAFT polymerization**

The inherent difficulty in performing photoRAFT resided in the fact that RAFT agents are typically good UV-absorbers and are therefore prone to UV-degradation. Conditions must be chosen to allow for initiation of the polymerization while retaining the integrity of the RAFT end group functionalities. Several strategies are discussed in the following sections, starting with the so-called iniferter polymerization mechanism. Photoinduced iniferter reactions can be seen as precursor technique of RAFT as some features between both techniques are similar. Iniferter initiation processes played an important role in the development of modern photo RAFT procedures and are thus described first.

#### **3.1 Iniferter polymerization**

Iniferter (*initiator-transfer agent-terminator*) polymerization is among the oldest living/controlled radical polymerization techniques.[161, 162] An iniferter agent is defined as a molecule that can decompose into radical fragments to initiate chain growth. Iniferter-derived radicals were thereby expected to end chain growth by end-capping the growing chains. This end-capping was either reversible, or to be prone to transfer reactions. The iniferter technique was followed by a photoinitiated iniferter method. In a photoiniferter polymerization a UV-labile bond is homolytically cleaved, followed by polymerization and end-capping of polymer chains. The polymers, containing specific thioether or thioester end groups, can be reactivated by UV-light, effectively establishing a reversible termination equilibrium. Dithiocarbamates form one class of photoiniferter agents, but could also act as RAFT agents. Dithiocarbamates are good control

agents for vinyl acetate RAFT polymerizations.[163] Vice versa, classical RAFT agents, such as dithiobenzoates and trithiocarbonates can in principle also act as RAFT mediators and concomitantly as photoiniferters (see below).

Photoiniferter reactions can be considered as the first example of a successful photoRDRP process. Polymerization of a variety of monomers was possible, including styrene, acrylates, methacrylates, vinyl esters and derivatives thereof, however with varying levels of success.[161, 164-167] As mentioned above, dithiocarbamates were most commonly used but dithiols, selenides or phosphines could also act as photoiniferters.[162, 164, 167-171] The recently developed nitroxide-mediated photopolymerization can also be classified as a photoiniferter process.[172] Complex block copolymer architectures have been reported for classical photoiniferter polymerization.[166, 173-175]

### **3.2 RAFT initiated by conventional photoinitiators**

RAFT polymerizations can be initiated by conventional PIs. Most dithioesters are colored and absorb light in the range of ca. 500-600 nm. Hence, even under visible light irradiation, decomposition of the RAFT agent may occur, rendering photoRAFT polymerization reactions with these transfer agents difficult. Trithiocarbonates are typically blue-shifted, and hence absorb in the region below 400-450 nm. Still, trithiocarbonates decompose under UV light, but usually to lesser extent and are thus a better choice for any photoRAFT reaction.

The first photoRAFT polymerization was demonstrated in 2002.[176] Two different dithiobenzoates were employed to control styrene polymerization. Initially, well controlled polymers were obtained, ( $D < 1.2$ ), however only at <30% monomer conversion, thus short irradiation times. When extending the reaction time to higher conversions, discoloration of the RAFT polymer was observed, accompanied by broadening of the molecular weight distributions. The effect was attributed to the loss of RAFT agent in secondary decomposition reactions after homolytic cleavage of the C-S bond. Interestingly, no external photoinitiator was employed, and hence the polymerization relied on initiation of the reaction by the iniferter mechanism (see section 3.3).

Later, trithiocarbonate, specifically DDMAT, *S*-dodecyl-*S'*-( $\alpha,\alpha'$ -dimethyl- $\alpha''$ -acetic acid, was used for photoRAFT reactions.[177] When exposed to longer wavelength UV light, > 365 nm, only a few percent of the RAFT agent decomposed even after irradiation over a few hours. With this RAFT agent, in combination with a conventional PI (2,4,6-trimethylbenzoyl)diphenyl

phosphine (TPO), poly(methyl acrylate) was obtained up to 90% conversion with dispersity below 1.1 and with an excellent linear correlation between  $M_n$  and conversion. Block extension and reactivation of the residual macro RAFT agents were presented, underpinning that if the wavelength regions of the RAFT agent and the PI were chosen appropriately, polymerizations could be efficiently performed.[178] This approach was used in kinetic studies,[179-182] via so-called pulsed laser polymerization experiments, and propagation and termination kinetics of macroradicals could be assessed. The RAFT process was employed to probe chain-length dependencies, especially in the termination reaction. In these experiments, laser generated populations of initiating radicals rapidly enter the main equilibrium of the RAFT process, hence converting all radicals into macroradicals with a specific chain length.[181, 183] Later this concept of pulse laser-induced RAFT polymerization was extended to study RAFT kinetics.[184-190]

PI RAFT reactions have successfully been applied to numerous monomers. The outcomes of polymerizations under UV and/or visible light irradiation were comparable to results obtained under thermal polymerization conditions, but also under mild reaction conditions. This was especially interesting for preparation of polymers with distinct lower critical solution temperature (LCST) behavior. Table 5 gives an overview of the selection of photoRAFT reactions reported in literature. In almost all cases, trithiocarbonates were chosen over dithiobenzoates for the reasons discussed above. Light sources varied largely, with a trend towards visible light irradiation with a full irradiation spectrum or blue and green light. In general, when UV light in the range of 365 nm was employed, a superimposition of the typical RAFT mechanism with the iniferter reversible termination mechanism could be postulated (see section 3.3 for details). A distinction between both mechanisms was difficult since product were identical for both mechanisms. Little information could be gathered from the typical tests of controlled polymerization. The presence of reversible termination iniferter reactions was indicated by a dependence of the overall polymerization rate on the concentration of the RAFT agent. Increasing  $c_{\text{RAFT}}$  resulted in increasing rate, which could only be explained on the basis of the iniferter mechanism. Obviously, the higher the wavelength employed for the light source, the less influence was from iniferter reactions.

Homo- and block copolymers were synthesized such as polystyrene, PS, polystyrene)-*b*-poly( $\alpha$ -vinylbenzyl alcohol) PS-*b*-PVBA[191] and poly(butyl acrylate), PBA[192] using the RAFT agent S,S'-bis( $\alpha,\alpha'$ -dimethyl- $\alpha''$ -acetic acid), BDMAT. Poly(glycidyl methacrylate), PGMA, and

poly(butyl methacrylate), PBMA and block copolymers were interestingly polymerized with a dithiobenzoate, namely 2-cyanoprop-2-yl(4-fluoro) dithiobenzoates, CPFDB.[193] Even when a broadband irradiation source was used, no breakdown of the RAFT agent was observed on the timescale of the polymerization. Block copolymerizations were targeted with visible light. The ethyl substituted derivative of DDMAT was used to obtain water-soluble polymers based on poly(*N*-(2-acryloyloxyethyl) pyrrolidone) (PNAP) and poly(poly(ethylene glycol) acrylate) (PPEGA) directly by reaction in water as reaction medium. [194] Another water-soluble polymer, poly(*N*-isopropyl acrylamide) (PNIPAAm), was obtained using BDMAT and Irgacure 2959 as PI in DMF solution.[195] Also, vinyl acetate was polymerized using xanthates.[196, 197] With the xanthate methyl(ethoxycarbonothioyl)sulfanyl acetate (MESA), poly(vinyl acetate) (PVAc) with well controlled molecular weight was obtained. Interestingly, some degree of stereocontrol could also be achieved as the solvent and polymerization temperature had an influence over tacticity. In this way, stereoblock copolymers were synthesized.

#### **Insert Table 5.**

### **3.3 Initiator-free photoRAFT**

As described above, contributions from iniferter initiation and reversible termination to the PI-induced photoRAFT reactions were commonly observed, even if the mechanistic situation was somewhat ambiguous regarding quantification of the effects.

When RAFT agent is continuously decomposed under UV-incidence but only to a low extent, the radicals produced from the iniferter mechanism can maintain the reaction, essentially leaving the RAFT equilibrium unaffected. Scheme 11 describes this case. The RAFT mechanism with its pre- and main equilibrium remained fully active and was unaltered compared to a conventional RAFT process. The only difference is that the initial RAFT agent could serve as a radical source by homolytic C-S bond cleavage. The thiyl radical that is at that instance created can trap propagating macroradicals, yielding a macroRAFT agent that takes part in the main RAFT equilibrium. They can also dimerize forming thiuram. The R group radicals initiate chain growth and act as a conventional initiator. Once all initial RAFT agent is consumed by the pre-equilibrium (composition plays only a minor role with regards to RAFT agent conversion), radicals are generated by the reversible termination action of the macroRAFT. A beneficial side-effect of this process is that all residual chains (with the exception of chains stemming from bimolecular termination) carry not only the desired trithiocarbonate (TTC) end group in  $\omega$ -

position, but also the R group in  $\alpha$ -position. Hence, polymers are more pure than thermally initiated RAFT polymers, or polymers obtained from PI-initiated photoRAFT, where a fraction of initiator-derived radicals is always found in the polymer as a side product. A similar approach was previously reported for concurrent ATRP/RAFT process.[198-200]

### Insert Scheme 11.

As with the PI-induced RAFT reactions, trithiocarbonates are preferred for initiator-free photoRAFT reactions, as decomposition should be kept low to reduce consumption of the RAFT agent in secondary decomposition reactions. The average radical concentrations must also be kept low in order to prevent irreversible bimolecular termination. Therefore a UV light source is preferred over visible light, as trithiocarbonates absorb typically between 350 and 400 nm.[178] For example, poly(acrylic acid) (PAA) was obtained with different molecular weight directly from polymerization in water at room temperature by reaction with BDMAT, achieving polymers in the molecular weight range of 3,000 to 40,000 with dispersities below 1.1.[201] The same principle was applied to polymerization of styrene. However, instead of a conventional TTC RAFT agent, the complexity of agent was increased by polymerizing a bifunctional TTC with dichloro-*p*-xylene, which yielded a polymer with a molar mass of few thousand Da with several TTC moieties built into the main chain. Polymerization of such polyTTC directly yielded multi-segmented polymer materials, which could be later degraded into their single segments by aminolysis. These results show that no significant difference exists between a thermal RAFT process (employing then a thermal initiator) and an initiator-free photoRAFT reaction. Segment growth of the PS units was uniform, well controlled and could be reinitiated after further monomer addition, yielding polymers of up to 30,000 with each segment of ca. 4,500 ( $D = 1.28$ ).[202] Using the same RAFT agent, the polymerization of PBA and PS was achieved.[203] Both monomers were successfully polymerized with very good match between theoretical and experimental molecular weight. Since BDMAT is a symmetrical TTC, chains grew to either side of the RAFT agent. Consequently, chain extension led to formation of triblock copolymers. In this manner, high molecular weight PS-*b*-PBA-*b*-PS polymer with  $M_n > 200,000$  was obtained (see Table 6 for a collection of examples of initiator-free photoRAFT polymerizations). DFT calculations were performed on the bond dissociation energy of the C-S bond, showing that the tertiary nature of the leaving group of BDMAT was essential for the process. If the R group was replaced by a less stabilized radical fragment, less control was achieved over the RAFT polymerizations, which was also confirmed by experiment.[204] PVAc was obtained using

xanthates as mediators/initiators.[205] Similar to the phenylbis(2,4,6-trimethylbenzoyl)-phosphine oxide (BAPO)-initiated polymerizations (Table 5), well controlled polymerizations were observed with a polymerization rate increasing with increasing xanthate concentration. To facilitate decomposition and re-initiation of the R group, a specific derivative of MESA was employed, *S*-2-cyano-2-propyl-*O*-ethyl xanthate (CPEC), which decomposed to yield a cyanoisopropyl fragment, as would normally be obtained from AIBN. Polymerizations without BAPO were slower, but still faster than their thermal counterparts.

In a detailed study on the kinetics of initiator-free photoRAFT, the effect of RAFT agent concentration (DDMAT) on methyl acrylate polymerization was examined, alongside a detailed study on the effect of increasing light intensities.[206] In-situ NMR measurements further confirmed the suggested mechanism, shown in Scheme 11. Higher light intensities led to increased radical concentrations and hence acceleration of polymerization. Figure 26 displays the fourfold rate increase in rate of polymerization when going from 3 mW·cm<sup>-2</sup> to 48 mW·cm<sup>-2</sup>. It is noteworthy that the intensity increase showed no negative effect on the polymerization. In all cases polymer of almost identical molecular weight and dispersity was obtained.

#### **Insert Figure 26.**

Interestingly, the iniferter and RAFT mechanism could also be decoupled.[207] Starting from a xanthate-terminated poly(*N*-vinyl carbazole) (PNVP) or xanthate-terminated PVAc, the iniferter reaction was used to initiate chain growth with acrylic crosslinkers. As xanthates are not efficient mediators for activated monomers, no RAFT equilibrium was established. Rapid crosslinking occurred, showing that RAFT end groups could also generally serve as efficient macroinitiators for free radical polymerizations.

Whether TTCs could be activated under visible light conditions was then investigated.[208] In fact the excited  $\pi^*$  state can also be reached by absorption in the blue light region when TTCs were employed. Consequently, polymers could be obtained by irradiation with blue LED. Polymers with pristine end group fidelity were obtained, which was further exemplified by multiple chain extensions of the RAFT polymer. Hexablock amphiphilic block copolymers were obtained using this process, thus confirming a high end group fidelity. Results from polymerizations from this study are summarized in Table 6. The potential of initiator-free photoRAFT was further explored for precision polymer design (Figure 27).[209] Very pure star polymers could be obtained based on PMA polymers made via DDMAT, with variation in the

number of arms depending on the length of the PMA building block and crosslinker concentration. Polymeric star nanoparticles were obtained with molecular weights ranging from roughly 55,000 (7 arms,  $D = 1.59$ ) to 38,450,000 (~720 arms,  $D = 1.94$ ). The reshuffling of trithiocarbonates and also thiocarbamates was successfully applied to prepare self-healing materials under UV and also visible light conditions under oxygen-free and ambient conditions.[210, 211]

**Insert Table 6.**

**Insert Figure 27.**

Recently, visible light mediated photolysis RAFT agent was applied under a catalyst-free photoRAFT conditions using blue and green light.[212] It was further extended to a polymerization-induced self-assembly (PISA) system. [213]

As described above, good polymerization results were achieved for PI-free as well as PI-initiated photoRAFT polymerizations. Polymers with similar structures were prepared with a slight advantage of PI-free photoRAFT, since no PI fragments would incorporate as the end group in the  $\alpha$ -position of a fraction of the chains. An important difference that did exist, however, was reaction rate. Photoinitiator-free photoRAFT scaled in overall polymerization rate with the initial RAFT agent concentration, while PI-initiated photoRAFT scaled with PI concentration. Consequently, photoRAFT reactions in absence of PIs typically proceeded on the timescale of several hours to tens of hours, while PI-initiated photoRAFT often proceeded within tens of minutes to only few hours reaction time. The only exception was vinyl acetate polymerization mediated by xanthates. VAc polymerized even in absence of a PI within one hour, with PI reactions running to completion in only a few minutes. The long reaction times in the PI-free scenario therefore often limit the achievable conversion and many examples discussed above were stopped at conversions below 50 %.

### **3.4 Photoinduced electron transfer (PET)-RAFT**

An extension to the PI-induced RAFT concept was recently introduced, which was termed photoinduced electron transfer (PET)-RAFT polymerization.[214] Inspired by the work of Ir-mediated photoATRP,[119, 215] photoredox catalysts were used to initiate a polymerization, which then proceeded via a classical RAFT polymerization pathway.

PET-RAFT can be classified as an efficient PI-induced photoRAFT process, as it featured similarities with PI-free photoRAFT. Kinetically, the reaction is driven by the metal complex

concentration (ppm range). From a product point of view, no PET-initiator is incorporated into the polymer product. Thus, PET-RAFT combines the advantages of the above described photoRAFT systems. Moreover, the use of metal complexes for initiation allows the procedure to work with materials that absorb in the visible light region, hence giving the possibility to completely decouple the RAFT process from any background-iniferter or RAFT agent decomposition reactions. Consequently, reactions proceeded comparatively fast (the system given in Scheme 12 was not optimized for rate) but later PET-RAFT systems allowed for reaction times in the order of just few hours. End group fidelity was high. This system resembles concurrent ATRP/RAFT systems (without irradiation). [198-200]

#### **Insert Scheme 12.**

In the initial paper,[214] PET-RAFT was shown to be highly versatile, demonstrating the successful photo-controlled polymerization of MMA, MA, tBuA, St, *N,N*-dimethylacrylamide, *N*-(2-hydroxypropyl) methacrylamide (HPMA), NIPAAm, oligo(ethylene glycol) methyl ether methacrylate (OEGMA), OEGA, isoprene, vinyl acetate, vinyl pivalate (VP), NVP and dimethyl vinylphosphonate (DVP). Trithiocarbonates were used for activated monomers; while xanthates were used for non-activated ones, Scheme 12. Polymerizations proceeded in a well-controlled manner, and up to decablock copolymers were achieved.

Follow-up research significantly broadened the application of PET-RAFT. Polymerizations were largely independent of oxygen content, and a similar degree of livingness of polymerizations was observed in presence or absence of air (~80 %).[216] This feature also allowed for the synthesis of a variety of quasi-pentablock copolymers by sequential addition of monomers without purification steps in between chain extensions.[217] Using a ruthenium catalyst ( $\text{Ru}(\text{bpy})_3$ ), similar oxygen tolerance could be observed, leading to well-defined polymer products and synthesis of decablock copolymers.[218] In a further study, PET-RAFT conditions were combined with Lewis acid mediators that influence tacticity of polymers, by which stereoblock copolymers could be obtained.[219]

#### **Insert Table 7.**

PET-RAFT, similar to photoATRP, is not necessarily dependent on metal-based catalysts and organocatalysts are able to induce a PET process. Thus, efficient PET-RAFT reactions were demonstrated using eosin Y,[220] donor-acceptor porphyrins[221] or even chlorophyll.[222] In

all three cases, good control over polymerization was achieved, leading to different polymers, see Table 7 for examples. Porphyrins were combined with metals in order to tune the absorption wavelength over the whole visible light spectrum.[223] PET-RAFT was already adopted for use under the name of conventional PI-induced photoRAFT. Curcumin was demonstrated as an efficient PET-RAFT catalyst.[224] Nb(OH)<sub>5</sub> nanoparticles were also used as initiators, as well as dimanganese decacarbonyl Mn<sub>2</sub>(CO)<sub>10</sub>, [225, 226] and 10-phenylphenothiazine (PTZ) as organocatalyst.[227]

PET-RAFT system was extended to conduct the polymerization by near-infrared/far-red irradiation.[228, 229] It was especially useful to polymerize some photosensitive monomers such as *o*-nitrobenzyl methacrylate.[230] A selective activation of RAFT agents under different wavelength was reported, achieving single monomer addition. [231]

### **3.5 Applications**

Despite being a relatively recent field of research, photoRAFT in its different variations was applied in a number of different applications, with a strong focus on crosslinked materials and (nano)particle synthesis. Examples of various applications are given in the following section.

#### **3.5.1 Advanced photoRAFT**

As discussed in section 2.8, continuous flow photoreactors could avoid several obstacles in photoinduced reactions.[155] The RAFT/iniferter reaction of a symmetrical trithiocarbonate[232] was significantly faster in a Booker-Milburn designed flow reactor compared to batch reactions.[233] Within one hour, high monomer conversions were reached for a series of acrylates and acrylamides, with dispersities below 1.2, allowing the synthesis of PDMA with molecule weights above 100,000 while batch reactions require much more extended reaction times. Photoinduced RAFT of (meth)acrylates and acrylamides was conducted in a tubular continuous flow reactor with Irgacure 369 and Irgacure 819 at ambient temperature. [234]

#### **3.5.2 Particle synthesis**

In the initial photoRAFT dispersion polymerization, styrene was polymerized in the presence or absence of AIBN with a dithiobenzoates CTA.[235] The dithiobenzoates carried a carboxy moiety on the R-group to make the agent surface active. The suspension was stabilized by poly(vinylpyrrolidone). The typical behavior of an initiator-free photoRAFT polymerization was observed; increasing rate with increasing RAFT agent concentration, but interestingly no large variation in rate when AIBN was added. This could be attributed to the low efficiency of AIBN as

photoinitiator. While control over molecular weight was lower than in a solution polymerization and dispersity was in the range of 2, overall satisfying results were obtained, demonstrating that photoRAFT was applicable to dispersed systems. An improvement of the dispersion system was reported with sodium dodecyl sulfate (SDS) as surfactant. The RAFT agent used was a surfmer, coined surfiferter.[236] A well-controlled PMMA was obtained via emulsion approach, with molecular weights in the range of 400,000 and dispersities around 1.3.

A dispersion polymerization was employed to produce close to monodisperse microspheres from PI-initiate photoRAFT polymerizations of MMA.[237-239] The use of a trithiocarbonate (BDMAT) allowed preparation of monodisperse particles in a one-stage procedure by delaying the initial polymerization rate during the nucleation phase. Polymers followed a controlled polymerization pathway and could be reactivated after addition of fresh monomer, Figure 28. In this way, a variety of functional monomers could be introduced to form functional surfaces on the microspheres, including poly(vinylpyrrolidone) or HPMA. This work also presented a pathway towards synthesis of functional microspheres by using hydrophilic macroRAFT agents as surfmer to stabilize the particles.[240] Various macroRAFT agents based on the TTC DDMAT and PEGA, AA and 4-vinyl pyridine were prepared, for example the microspheres containing AA on the surface showed pH-responsive behavior. Subsequently, variations of core-shell microsphere particles were synthesized, e.g. particles with thermo-responsive properties based on PEGA and methoxyethyl acrylate.[241] The concept was also suitable for synthesis of metal-composite particles,[238, 240] or particles suitable for bioconjugation via 1-ethyl-3-(3-dimethylaminopropyl)carbodiimide (EDC) coupling.[242] In a similar procedure monodisperse microparticles were obtained by direct photoRAFT of crosslinking monomers.[243]

**Insert Figure 28.**

Miniemulsion polymerization has also been addressed in addition to suspension and conventional emulsion polymerization, by employing a PET-RAFT process. In the controlled polymerizations of styrene using the Ir(ppy)<sub>3</sub> catalyst system, good temporal control was achieved and characteristics of controlled polymerization was retained.[244]

Nanoparticle synthesis was reported using polymerization-induced self-assembly (PISA) procedure.[245, 246] MacroRAFT agents based on oligo(ethylene glycol) acrylate were chain extended in a PET-RAFT process with a Ru(bpy)<sub>3</sub> catalyst under blue light irradiation. Self-assembly was induced by chain extension with benzyl acrylate. Depending on catalyst

concentrations (Figure 29), different particle morphologies were obtained. While the catalyst had no direct influence on individual block lengths, it altered the rate of polymerization and hence longer benzyl acrylate blocks were obtained with higher catalyst loadings. Overall, spherical micelles, vesicles and worm-like micelles could be obtained by this method simply by variation of solid content and catalyst concentration.

**Insert Figure 29.**

### **3.5.3 Synthesis of gels and networks**

PhotoRAFT reactions were applied to the synthesis of networks in the bulk phase. Dental resins became more stress-resistant when RAFT agents were added to the photopolymerization formulation. The RAFT agent regulated the segment length between two crosslinks, altering the strength of the network.[247] Post-polymerization modifications became possible when the networks were formed in which the typical RAFT groups were incorporated into the main chain rather than being attached at the end of chain branches. A polymer gel synthesized from monomer and a trithiocarbonate-containing crosslinker monomer did not only lead to a controlled network polymerization, but also allowed control over the mesh size in the network after polymerization, simply by irradiating the gels with UV light.[248] The C-S bonds were reversibly cleaved by the action of the UV light, giving the network the chance to rearrange depending on the external circumstances, in a quasi-self-healing process. In a similar approach, a symmetrical RAFT agent with dienophile end groups could form crosslinks in tetrazine Diels-Alder reactions. The polymerization could be carried out following the iniferter initiation mechanism. Also, a reshuffling of the network structure and chain extension of the network segments became accessible by light irradiation in a post-polymerization process when fresh monomer was present.[249, 250]

## **4. Photomediated NMP**

Nitroxide-mediated radical polymerization (NMP) is one of the most common CRP techniques that employs reversible thermal homolysis of C–ON bond of an alkoxyamine to produce an initiating radical and a persistent radical, a nitroxide. The conventional NMP requires high temperatures and long polymerization times for the activation of strong oxygen-carbon bonds present in the dormant species. The controlled polymerization of meth(acrylate)s is rather difficult due to their high equilibrium constants. The slow recombination of the nitroxide and the sterically-hindered chain-end radicals lead to high concentration of active radicals that promote

the formation of “dead” polymers via self-termination at low conversions. This limitation has been partly compensated by recently developed alkoxyamines having higher equilibrium constants at lower temperatures thereby providing acceptable rate of polymerization for meth(acrylate)s. However, the use of the second generation alkoxyamines in NMP still require a high temperature, such as heating at 90 °C. Photosensitive alkoxyamines that generate reactive species by light irradiation have been recently introduced to mediate controlled radical polymerization at ambient temperatures. The initial work on photoinduced NMP was based on the homolysis of TEMPO (2,2,6,6-tetramethylpiperidinyloxy-4-yl radical)-based alkoxyamines.[251] In this case, photochemical activation of TEMPO derivatives containing either xanthone or pyrene moieties generated reactive carbon-centered radicals and stable TEMPO radicals. Although the photosensitization step was successful, these molecules could not be used in polymerization reactions due to the low efficiency in generating the carbon-centered radicals.[252, 253] After these initial results, new photosensitive alkoxyamines bearing different chromophores covalently linked to the TEMPO-OH moiety were designed. Some successful polymerizations were reported using these compounds but either uncontrolled polymers with much higher molecular weights and broader molecular weight distribution or very low monomer conversion were also obtained (Table 8).

Quite recently, new second generation alkoxyamines bearing photosensitive benzophenone, quinolone or xanthone groups were developed to accomplish NMP under more ambient conditions.[172] In addition to experimental studies on the influence of the types and positions of chromophore as well as nitroxide groups on the NMP process, theoretical studies were extensively investigated.[254-257] In the first step, the photosensitive alkoxyamine generated an active alkyl and a stable nitroxide radical under UV light (Figure 30). The alkyl radicals could initiate polymerization, whereas the nitroxide radicals were responsible for reversibly deactivating the propagating radicals, establishing a rapid equilibrium between active and dormant species. In order to enhance the energy transfer as well as generation of alkyl radicals, the chromophore groups were deliberately located close to the aminoxyl function, 1 or 2 carbons, which was responsible for the cleavage of the C-O bond. Faster polymerization rates, 80% conversion in 500 s. for NMP of n-butyl acrylate were observed in addition to partial living characteristics, as suggested by the close agreement between experimental and theoretical molecular weights and the relatively low molecular weight distributions.

**Insert Figure 30.**

Notably, these photosensitive alkoxyamines could potentially act as a bimolecular photoinitiator in combination with a co-initiator, e.g. alcohols, amines or thiols, and produce two different radicals during irradiation. Some of these radicals might contribute to an independent initiation resulting in a polymerization that proceeds in an uncontrolled manner. The photosensitive alkoxyamines could be also efficiently direct the thermally induced NMP process by heating at 120 °C. Using these two different NMP modes, the thermal polymerization of styrene at 120 °C and sequential photopolymerization of isobornyl acrylate, block copolymers with moderate molecular characteristics were obtained.[258] In a recent study, photosensitive nitroxides bearing different chromophore groups (benzophenone, naphthalene and quinoline) were developed for the photoinduced NMP of methyl methacrylate in the presence of 2,2-dimethoxy-2-phenyl acetophenone as photoinitiator.[58] In contrast to linear polymer growth with narrow molecular weight distributions (1.3–1.4), a loss of end-group functionality was confirmed by a chain extension study.

### **Insert Table 8.**

## **5. Other Controlled Radical Polymerizations**

### **5.1 Photomediated CMRP**

The use of organocobalt(III) compounds in cobalt-mediated radical polymerization (CMRP) is receiving growing attention because of the reversible formation of a carbon–cobalt covalent bond under mild conditions.[259, 260] Since the Co-C bond can be easily cleaved by heat or light, CMRP was successfully applied for various monomers such as vinyl acetates, acrylates and acrylamides.[261] In earlier studies, UV light not only facilitated the generation of initiating radicals from photosensitive alkyl cobaloxime compounds but also manipulated the reversible cleavage of the Co–C bond throughout the polymerization.[262] However, the scope of the cobaloxime compounds was limited to acrylates. The use of bis-acetylacetonatocobalt complex ( $\text{Co}(\text{acac})_2$ ) in CMRP not only broadened the range of monomers but also allowed the polymerization to be conducted at lower temperatures, even below room temperature.[263] The photoinitiated version of CMRP using  $\text{Co}^{\text{II}}(\text{acac})_2$  and UV irradiation of a typical azo-initiator provided well-defined polymers with very high molar masses. Subsequently, this method was further extended by using commercially available photoinitiators, such as 1,2-diphenyl-2,2-dimethoxyethanone and bis(2,4,6-trimethylbenzoyl)phenylphosphine oxide. The photodissociation of organocobalt(III) compounds into an organic radical and a cobalt(II) metal-centered radical upon exposure to UV or visible light was confirmed by both spin-trapping and

kinetics of polymerization experiments. The Co-C bond was highly photosensitive so that the polymerization only occurred under light irradiation when the light source turns off, polymerization was immediately stopped. Furthermore, the polymerization kinetics could be tuned by photoinitiator type, its concentration and irradiation time, which directly affected the initiating efficiency of the system.

Recently, organocobalt porphyrin derivatives have been developed for thermally and photochemically initiated CMRP to eliminate the need to use conventional radical initiators.[264] By applying a visible light source or sun-light to the organocobalt porphyrin, a number of structurally different monomers, both hydrophilic and hydrophobic, could be successfully polymerized in a controlled manner. Some key features of this system included mild reaction conditions and no need for additional compounds such as dyes, photosensitizers, or reducing agents. The organocobalt porphyrins were highly photosensitive so that the visible light irradiation could facilitate mediation of chain growth by reversibly activating the Co-C bond (Figure 31). A series of well-defined multi-block copolymers containing acrylate and acrylamide segments were prepared via photoinduced CMRP using a simple household light source or sunlight.[265] Furthermore, this system provided excellent temporal control over the chain-growth process, enabling the synthesis of complex macromolecular structures.

**Insert Figure 31.**

Structurally similar organocobalt salen complexes (salen is a common abbreviation for N<sub>2</sub>O<sub>2</sub> bis-Schiff base bis-phenolates) were also used as control agents for the polymerization of acrylates, acrylamides and vinyl acetate, under ambient conditions, i.e. visible light irradiation at room temperature.[266] Since the organocobalt salen complex has a relatively weak Co-C bond, the direct photolysis of this complex generates an organic radical that can initiate polymerization and a persistent metal-centered radical that is responsible for the reversible termination of the radical on the growing polymer chain end. Because of its convenient synthesis, just mixing commercial available cobalt(II) salen complexes with a traditional photoinitiator, this system was more suitable for industrial applications, compared to organocobalt porphyrin systems.[267] This system enabled the preparation of well-defined homopolymers and copolymers with predetermined molecular weight and narrow molecular weight distributions under mild conditions.

The photomediated CMRP process was recently implemented in continuous flow reactors using (Co(acac)<sub>2</sub>) complex with UV light irradiation.[159] The flow reactor process not only led to a

significant improvement in polymerization rates, but also allowed scalable synthesis of polymers by simplifying purification steps. This process was successfully applied to the polymerization of vinyl acetate and its copolymerization with a less reactive olefin 1-octene to obtain random copolymers containing up to 50 mol% of 1-octene. This continuous photoflow system enabled reaching high conversion in several polymerization processes by increasing the light intensity and efficiency that reduced reaction times from days to minutes.

## 5.2 Photomediated ITP

Iodine transfer polymerization (ITP) was one of the earliest CRP methods and is based on the cleavage of C-I bond via a degenerative chain transfer mechanism.[268] The method was particularly suitable for the polymerization of fluorinated and phosphonic monomers and was used industrially to produce high value thermoplastic elastomers based on hard-soft-hard triblock copolymers. The basic ITP system consisted of a monomer, an iodinated chain transfer agent and a conventional initiator, but did not comprise any metal. Due to the weak C-I bond, the iodinated chain transfer agents were very sensitive to both heat and light, and tended to decompose during synthesis, isolation, or storage. To overcome these limitations, the required iodinated compounds could be formed *in situ* from a free radical initiator in conjunction with iodine (I<sub>2</sub>). Conventional thermal initiators such as BPO, AIBN and potassium persulfate were mainly used, but recently, photoinitiators were used as alternatives that allowed conducting the polymerization under ITP conditions.[269]

The first example of photoinitiated ITP was a degenerative iodine transfer process under visible light irradiation using dimanganese decacarbonyl (Mn<sub>2</sub>(CO)<sub>10</sub>) in conjunction with alkyl iodide compounds.[270-272] The photochemical homolysis of the metal-metal bond of Mn<sub>2</sub>(CO)<sub>10</sub> generated a highly reactive metal-centered radical •Mn(CO)<sub>5</sub>, which directly abstracted iodine atoms from alkyl iodides. The newly formed carbon-centered radicals not only initiated ITP of unconjugated monomers such as vinyl acetate and vinylidene fluorides, but also conjugated monomers such as styrenes and acrylates in the presence of appropriate alkyl iodide-based initiators. The generated I-Mn bond in I-Mn(CO)<sub>5</sub> was too stable to reversibly deactivate propagating radicals and the system mainly operated by a degenerative chain transfer mechanism (Figure 32).

**Insert Figure 32.**

The system provided well-defined polymers with narrow molecular weight distribution and the experimental molecular weights relatively close to theoretical values and temporal control for the growth of polymer chains by switching on/off the light source. This system was also applied to the preparation of well-defined poly(vinylidene fluoride) and its copolymers using a variety of alkyl and perfluoroalkyl halides (Cl, Br, I).[273] The visible light-induced ITP using bifunctional alkyl and perfluoroalkyl iodides afforded telechelic poly(vinylidene fluoride)s with very high (> 95 %) iodide chain-end fidelity, which were further utilized for preparation of a series of ABA triblock copolymers containing vinyl acetate, *tert*-butyl acrylate, methyl methacrylate, isoprene, styrene, and acrylonitrile segments.[274-276]

The influence of UV light on photomediated ITP was also investigated under various experimental conditions.[277] ITPs of butyl methacrylate with cyanopropyl iodide as chain transfer agent and a conventional initiator, AIBN, were conducted at 60 °C under dark conditions and in the presence of UV light. The UV irradiation not only provided a significant increase in the molar-mass of the polymers but also lowered molecular weight distribution from above 1.5 down to 1.1 compared to non-irradiated system. The C–I bond was light sensitive and homolytically cleaved to give R• and I• radicals upon UV irradiation. Since the light produces radicals needed for the ITP, its use not only accelerated the polymerization rate but also allowed formation of polymers with narrow molecular weight distributions. The effect of UV light on the polymerization rates was also investigated by continuous and intermittent UV irradiation experiments. The system was perfectly dormant with no monomer conversion occurring during the dark time, whereas the chain growth could be immediately restarted upon further irradiation.

Recently, a new concept for ITP was developed and consists of an alkyl iodide (dormant species) and an amine (catalyst) and uses visible light (350–600 nm).[278] It is among the most simple, inexpensive, and robust photoinduced LRPs. Mechanistically, the key reaction is the photolysis of free dormant species and dormant species/catalyst complex, which frequently occurs as the main activation process. The subsequent photochemical dissociation of this complex prevented “irreversible” reactions in the process and enabled the polymerization of a variety of acrylates to yield relatively well-defined polymers ( $M_w/M_n < 1.4$ ). The temporal control of the system was also demonstrated by switching light “on” and “off”. No polymerization occurred in the dark, meaning that the formed amine complex efficiently quenched the propagating radicals. By turning on the light, the restoration of the polymer chain growth of MMA was achieved in all cycles. The wavelength flexibility of this process further was extended to the visible range by

employing various organic dyes (CY-1 ( $\lambda_{\text{max}} = 530$  nm) and CY-3 ( $\lambda_{\text{max}} = 602$  nm) cyanine dyes, and SP-1 ( $\lambda_{\text{max}} = 718$  nm) spiropyran dye) as photocatalysts.[279] Although, the similar mechanism was described, the photocatalyst acted as antenna that transferred light energy to cleave the C–I bond of the complex (Figure 33). This process was responsive to the irradiation wavelength; the polymerization could be conducted at the desired wavelengths by choosing suitable organic dyes. All three organic dyes produced well-defined polymers ( $M_w/M_n < 1.3$ ) in good yields. The process was compatible with various functional groups and could be used to synthesize a wide range mechanistically different block copolymers in a one-pot mode. A block copolymer of MMA and Valero lactone (VL) was prepared by selectively regulating ITP and ring-opening polymerization using orthogonal irradiation techniques. The visible light exposure of the reaction mixture led to PMMA with 67% yield, while no conversion of VL could be observed, consistent with selective light irradiation. The following UV light irradiation ( $\lambda = 350\text{--}380$  nm) resulted to 80% VL monomer conversion with only 1% further MMA conversion, and resulted in formation of a well-defined block copolymer ( $M_w/M_n = 1.2$ ). By taking advantage of the wavelength-responsive nature of this technique it might lead to one-pot synthesis of block copolymers using different polymerization mechanisms.

**Insert Figure 33.**

### **5.3 Photomediated TERP**

Photochemical stimuli was efficiently implemented for the activation of organotellurium-mediated radical polymerization (TERP) of meth(acrylate)s.[280] Two distinct activation mechanisms, namely the degenerative chain transfer and reversible termination, were involved and the TERP process mainly occurred by the degenerative chain transfer mechanism (Figure 34).[281] The direct photolysis of the organotellurium compound led to homolytic cleavage of C–Te bond with high quantum yield, 0.84. The directly generated tertiary carbon-centered radicals were responsible for the quantitative initiation of TERP even under visible light. The reversible deactivation process not only prevented irreversible termination reactions but also provided the synthesis of well-defined polymers with narrow molecular weight distribution ( $M_w/M_n = < 1.2$ ), with the exception of poly(acrylic acid) which was slightly broader,  $M_w/M_n = 1.34$ .

**Insert Figure 34.**

In order to extend its wavelength sensitivity into the visible light range, ditelluride compounds, having a stronger absorption coefficient, were used the TERP process.[282] Since the quantum efficiency of organotellurium compounds was so high, the light intensity of the TERP process could be reduced by either using filters or altering UV light source with visible light sources like LED lamps. Indeed by tuning of the light intensity, well-defined polymers with narrow molecular weight distributions were synthesized under LED irradiation with or without a neutral density filter. In addition, no polymerization occurred in the absence of light, indicating temporal control of the polymer chain growth between active and inactive states by turning the light source on and off. However, the organotellurium compounds were very sensitive to air and light, foul-smelling, and difficult to handle. Because of these drawbacks, the development of simple preparation methods for organotellurium compounds still remains a challenging topic. In a recent study, a highly reactive benzoyl phenyltelluride, easily prepared from diphenyl ditelluride and benzoyl chloride in one-step, was developed for the photomediated TERP of various monomers under mild condition, visible light irradiation (400-500 nm) at room temperature.[283] Although the process was not feasible for styrenes and methacrylates, it produced well-defined polymers with narrow molecular weight distributions as low as 1.2 to 1.3 from acrylamides and acrylates.

## **6. Radical intermediates in other photomediated controlled polymerizations**

As discussed in previous sections, photomediated reactions generate free radicals which promote controlled radical polymerization. However in addition to free radicals, other reactive intermediates such as radical anions and radical cations, derived from photoredox catalyzed reactions, could participate in several other controlled polymerizations.

A photoinduced ring-opening metathesis polymerization (ROMP) was developed and mediated by pyrylium salts acting as metal-free photoredox catalysts.[284] Pyrylium salts have been identified as efficient catalysts for photoinduced redox reaction in organic transformations.[285, 286] The polymerization of norbornene (Figure 35), with a relatively high ring strain among ROMP monomers, was carried out with 1-ethoxy-1-propene as initiator and a pyrylium salt in a molar ratio of 1000:1:0.03 in dichloromethane under a blue LED irradiation (450 – 480 nm). After 120 min of irradiation, conversion of norbornene reached 61% based on  $^1\text{H}$  NMR, and polynorbornene (PNB) was formed with  $M_n = 60,200$  and  $M_w/M_n = 1.6$ , which was close to the theoretical value of 57,400. The synthesized PNB was confirmed by  $^1\text{H}$  NMR analysis.  $M_n$  increased with conversion, and the reaction could be stopped or restarted by turning the irradiation source off or on.

### Insert Figure 35.

The “living” property and “on/off” behavior of this photoinduced metal-free ROMP was clarified by the proposed mechanism (Scheme 13). Vinyl ethers with oxidation potentials in the range of 1.43 to 1.30 V vs SCE, could be feasibly oxidized by the excited state of pyrylium cations shown in Figure 35. The oxidation potential was calculated to be 1.74 V vs SCE.[287] After the electron transfer from vinyl ether **A** to the excited pyrylium cation, a vinyl ether radical cation **B** was formed as a key intermediate, which subsequently reacted with a cycloalkene monomer to provide the [2+2] cycloaddition intermediate **C**. Due to high ring strain, ring-opening of cyclobutane radical cation **C** formed **D** with radical cation chain end, which continually propagated with additional monomers to provide polymer chain **E**. It was also proposed that the propagating radical cation had a dynamic redox couple with the reduced pyrylium species, which allowed the reduction of radical cation chain end to terminate polymerization (**E** → **F**). This oxidation and reduction cycles represented the activation and deactivation cycles in a controlled polymerization.

### Insert Scheme 13.

After this original methodology for the synthesis of PNB, photoinduced metal-free ROMP of dicyclopentadiene (DCPD), another common ROMP monomer, was carried out to form linear polyDCPD.[288] The cyclopentane moiety in polyDCPD also allowed the preparation of a cross-linked polyDCPD by thiol-ene chemistry.

## 7. Combination of photomediated radical polymerizations and click processes

The term “click chemistry” defines a collection of reactions allowing for the efficient building of chemical libraries under mild reaction conditions using readily available reagents.[289] The criteria for being classified as click chemistry reaction are wide in scope, modular, proceeds in high yields, with none, or inoffensive by-products, tolerant to many functional groups and final products that can be easily isolable by non-chromatographic methods. The copper(I)-catalyzed azide-alkyne cycloaddition (CuAAC) was the first example of “click” chemistry.[290, 291] Since ATRP and CuAAC processes are both based on a similar Cu(I) catalytic system, their sequential or simultaneous combinations provide versatile routes for the construction of a variety of complex and functional macromolecular structures from heterofunctional molecules.[292-295] The mechanistic combination of ATRP and CuAAC techniques were usually achieved in a

sequential manner.[296-302] The general approach was either the use of an alkyne-functionalized ATRP initiator to synthesize a polymer bearing an  $\alpha$ -alkynyl end-group or a nucleophilic displacement of the terminal halogen of polymers prepared by ATRP with sodium azide. Application of a subsequent CuAAC click reaction with these reactive polymers afforded the desired product in a quantitative yield under ambient conditions. On the other hand, there were a few examples of the simultaneous combination of ATRP and CuAAC techniques.[299, 300, 303-305] The combination of photoinduced CuAAC and ATRP was also accomplished and the results comparable with the thermally combined version. Quite recently, synthesis of block[306] and graft[307] polymers were successfully achieved by combined one-pot photoinduced ATRP and CuAAC protocols (Figure 36). For preparation of the block copolymer, clickable azido-terminated polystyrene and prop-2-ynyl 2-bromo-2-methylpropanoate as ATRP initiator were mixed with methyl methacrylate and Cu(II)/PMDETA in acetonitrile. Photochemically regenerated Cu(I) activators could activate both CuAAC and ATRP processes under UV light irradiation. After 24 h, a polystyrene-*b*-poly(methyl methacrylate) block copolymer was obtained with 68% monomer conversion. The monomodal GPC peak with a clear shift to a higher molecular weight region confirmed the successful block copolymer formation without any detectable contaminant homopolymer.

**Insert Figure 36.**

The utility of this one-pot protocol for the synthesis of well-defined polystyrene-*g*-poly(methyl methacrylate) was also demonstrated.[307] In this case, a side-chain azide functional copolymer, namely poly(styrene-co-4-azidomethylstyrene), was used instead of azido-terminated polystyrene. After 4 h of UV light irradiation, the graft copolymer was obtained with good conversion of added monomer (72%) under mild reaction conditions. The reaction was carried out at room temperature without the necessity of an added photosensitive molecule. By varying the recipe and reaction conditions, choosing the proper ligand, initiator, and solvent, the combined one-pot photoinduced ATRP and CuAAC enabled the construction of the desired macromolecular structures in a well-controlled manner. ATRP was successfully combined with ring-opening, ionic and even polycondensation processes and this approach can be extended to photomediated processes. [308-311] In a similar approach, PET-RAFT was combined with orthogonal ROMP to synthesize block copolymer at ambient temperature.[312] Very recently, a photoacid-mediated ring opening polymerization irradiated by visible light was combined with photoRAFT to result in a one-pot sequential polymerization under two different wavelength. [313]

## 8. Surface-initiated photomediated radical polymerization

The temporospatial control of photoinduced reactions is ideal for manipulation of surfaces. Photopolymerizations in specific areas can be used to directly grow polymer brushes in specific patterns following CRP protocols (hence retention of end group functionality and the ability to form block copolymers). In the following, the performance of photomediated RDRP reactions for surface modification and patterning is highlighted.

### 8.1 PhotoATRP

The first reported copper-mediated photoATRP surface-initiated polymerization used 0.5 mol%  $\text{CuCl}_2/\text{bpy}$  as catalyst and  $\text{TiO}_2$  as photosensitizer.[60] Various hydrophilic, or even ionic, methacrylates such as MeOGEMA, NIPAAm and 3-sulfopropyl methacrylate (SPMA) were grafted from a flat gold substrate or titanium nanowires with water-methanol mixtures as solvent at low light intensities. SPMA films grew linearly with time and reached more than 200 nm in thickness (Figure 37). Increased thicknesses were found with increased light intensities, whereas the  $\text{TiO}_2$  concentration have no influence once a certain level was exceeded.

**Insert Figure 37.**

Polymer brush patterns were generated via combination of  $\mu\text{CP}$  and photoATRP.[60] More recently,  $\text{TiO}_2$  particles were imprinted with a dye, yielding faster reactions and making the reaction accessible to the use of visible light ( $66 \text{ mW/cm}^2$ ).[55] Remarkable nanoscale-resolved patterns were generated with anodized aluminum oxide as a photomask. Recently, the grafting from a nanocomposite of  $\text{TiO}_2$  and reduced graphene oxide as photosensitizer were reported in combination with tetrasulfonated copper phthalocyanine. The resulting pMMA brushes were cleaved in a saponification reaction over 7 days and showed a dispersity of around 1.3 and  $M_n = 25,000$ . [58]

The first copper-photoATRP grafting without photosensitizer was described for MMA from silicon substrates.[80] PMMA films of 15 nm thickness were formed after 5 hours irradiation. Successful block copolymer formation with tBMA and micropatterning indicated the controlled nature of the initial grafting from polymerization and exemplified the full potential of photoATRP grafting. Photosensitizer-free copper-photoATRP of acrylates was first reported to graft methyl acrylate from cellulose substrates.[85] Subsequent SEC characterization of the cleaved brushes confirmed the livingness of the procedure, however the polymers displayed

higher dispersities and higher molecular weights than polymers formed in the presence of a sacrificial initiator. Interestingly, block copolymer formation from secondary chlorinated initiators was feasible, although never reported for polymerizations in solution.

PhotoATRP grafting of various acrylates from silicon substrates was investigated.[76] *N*-2-hydroxypropylmethacrylate and MeOEGMA were grafted from substrates under photoATRP conditions.[82] In both cases, (meth)acrylate monomers, CuBr<sub>2</sub>/Me<sub>6</sub>TREN (1/6) in DMSO were exposed to UV light (360 nm, 36 W) with air cooling.[76, 82] The grafting rate was proportional to the irradiation time, even in the presence of ppb-range concentrations of catalyst. The catalyst concentration had some influence on the polymerization rate, and both very low (100 ppb) and relative high (100 ppm) copper catalyst concentrations led to a decrease in final film thickness compared to medium concentrations (1-10 ppm), Figure 38. The grafting rate was proportional to the monomer concentration. As reported for photoATRP in solution, the choice of solvent had significant influence on the polymerization rate. Interestingly, monomers with high excluded volume interactions such as *t*-butyl, 2-ethylhexyl acrylate or the ionic [2-acryloyloxyethyl]trimethylammonium chloride led to film thicknesses of up to 950 nm after 1 hour, whereas film thickness with other monomers stayed below 250 nm. Reinitiation and block copolymer formation of *t*-butyl and 2-hydroxyethyl acrylate were performed on the polymer brushes. No phase separation was found for the *Pt*BA and the poly(2-hydroxyethyl acrylate) blocks, rather a gradual composition shift from a 2-hydroxyethyl acrylate enriched upper layer towards almost pure *Pt*BA at the polymer-silicon interface was revealed via X-ray photoelectron spectroscopy (XPS) depth profiling (Figure 38).[76] A similar result was reported for *N*-(2-hydroxypropyl)methacrylamide and MeOEGMA.[82] ToF-secondary ion mass spectrometry (SIMS) depth profile showed maximum of the bromine concentration at the polymer/air interface and the polymer/silicon interface depending on the choice of sputter ion.[76] Basic spatial control was achieved via a shadow mask, but temporal control remained challenging.[76, 82] Antifouling properties of poly(HMPA) brushes were also demonstrated. [82] Photochemical modification of polypropylene surface with benzophenone containing bromoisobutyrate as ATRP initiator was followed by ATRP of dimethylaminoethyl methacrylate and subsequent quaternization with ethyl bromide. The quaternary ammonium moieties showed antibacterial activity against *Escherichia coli* (*E. coli*). The biocidal activity of the resultant surfaces depends on the amount of the grafted polymers (the number of available quaternary ammonium units). With the same grafting density, the surface grafted with relatively high MW polymers ( $M_n > 10,000$ ) showed almost 100% killing efficiency (killing all of the input *E. coli* ( $2.9 \times 10^5$ ) in the shaking test), whereas a lower biocidal

activity (85%) was observed for the surface grafted with shorter PQA chains ( $M_n = 1,500$ ).[314] In a related approach, thermoresponsive star polymers were photo-crosslinked in order to control of cell-surface interactions to enhance cell detachment.[315]

**Insert Figure 38.**

Besides Cu- and Ir-based catalysts, Fe[99] and Mn-mediated[316] photoATRP grafting protocols were also reported. In the Mn-mediated photoATRP processes  $Mn_2(CO)_{10}$  was used to graft NIPAAm, styrene, *N*-vinylpyrrolidone and *t*-butyl methacrylate from various substrates.[316] A linear increase up to 160 nm film thickness within only 6 min was observed. Despite these thick films, XPS measurements showed intense signals originating from the substrate, indicating a less than homogenous grafting. However, block copolymer formation, reinitiation and other control experiments were not assessed. Iron-catalyzed photoATRP grafting of pentafluoropropyl acrylate was reported.[99] Photomasks were employed to generate spatially resolved patterns. Reinitiation and block copolymer formation were only demonstrated for the polymer film.

Ir-mediated photoATRP of MMA from silicon substrates carrying a self-assembled monolayer (SAM) of BiB-based was successful, Figure 39.[317]. The substrates were covered with a thin glass slide to ensure short path lengths of the employed visible light. MMA, OEGMA and various fluorinated methacrylates were polymerized and yielded film thicknesses between 60 and 120 nm within a one hour brush polymerization time.[317, 318] Ionic monomers such as 2-methacryloyloxyethyl phosphorylcholine and sulfobetaine methacrylate formed only very thin layers, even with trifluoroethanol as co-solvent.[318] Catalyst concentrations were kept at 50 ppm for MMA,[317] but were increased to 500 ppm for OEGMA and fluorinated monomers.[318] The linear relationship between irradiation time, light intensity and film thickness and performed chain extension reactions confirmed the controlled nature of the reported grafting from procedures.[317, 318] This gave a direct access to 3D architectures on the surface via use of surface of photomasks with areas of different optical densities.[317] Due to short diffusion path lengths of the excited Ir-catalyst species, patterns with  $\mu m$  resolution were readily accessible. On-off experiments confirmed good temporal control. In a recent report, grafting on commercial PVC with Ir based catalyst was successful in solution and on surfaces. [319]

**Insert Figure 39.**

Block copolymers of OEGMA and TFEMA were grafted from surfaces, while the high hydrophobicity of TFEMA and its pronounced surface activity reduced its chain extension capability especially for thick layers of pTFEMA.[318] CF<sub>3</sub> moieties were present at the polymer/air interface even when pOEGMA was grafted as the second, outer block, underlining the low surface energy of CF<sub>3</sub>. The sequential use of Ir-mediated photoATRP with thermal Cu-mediated ATRP gave access also to ionic monomers. The generated micron-scale patterns were investigated with optical microscopy, scanning transmission X-ray microscopy (STXM) and SIMS. Ir-mediated photoATRP was also reported for the modification of membranes with antifouling polymers such as hydroxyethyl methacrylate (HEMA), GMA and OEGMA.[320] Reaction times were extended to 24 hours under fluorescent light (50W) and a linear relationship of grafting yield (mass) and time was up to 12 h exposure time, irrespective of the catalyst concentration which was generally kept below 300 ppm. Furthermore, a linear dependency on the monomer concentration was reported by the same group.

Recently, a visible-light mediated phenothiazine-catalyzed metal-free ATRP was extended to fabricate of well-defined polymer brushes from flat, curved surfaces and particle surfaces.[321] Single-layer patterns, gradient structures as well as block copolymer architectures could be readily obtained under irradiation from compact fluorescent lamps and natural sunlight. A linear relationship between brush height and irradiation time was observed, confirming the “living” nature of whole process. Moreover, chain-extension of a uniform PMMA brush (~30 nm) with TFEMA increased the brush thickness of 26 nm, indicating a good retention of active chain ends. In a following work, 2-bromo-2-phenylacetate based tetherable initiators were prepared and gave better initiation efficiency and performance in PTZ-catalyzed metal-free SI-ATRP. [322]

## **8.2 Photoiniferter and photoRAFT**

The first photoiniferter grafting methodology was reported in 1996.[323] Various monomers such as MMA, styrene and dimethylaminopropyl acrylamide were grafted from polymer substrates armed with 30 mol% diethyldithiocarbamate functionalities. These polymer substrates became more readily available after a dithiocarbamate-containing methacrylate monomer was synthesized.[324] Triblock copolymers were grafted in a controlled manner, visualized via fluorescence microscopy after staining, and film thicknesses up to 420 nm were reached.[323] Diblock copolymer brushes were also grafted from a silicon substrate, with a significantly reduced grafting rate, potentially due to the lower amount of initiating sites.[325] Photomask and neutral-density filters were employed to generate complex 3D structures, Figure 40.[323, 326]

More elaborated patterns with polymer grafting including formation of a gradient in film thicknesses were generated with x-y-adjustable photomasks.[327] A methodology for the generation of 2D and 3D patterned polymer substrates was developed.[328]

Patterns of second[329] or even third[330] generation hyperbranched brushes on glass slides were prepared from poly(p-chloromethylstyrene) brushes functionalized with additional iniferter initiators (Figure 40). Hydrophilic monomers such as MAA, OEGMA or DMAAm were grafted from the new iniferter initiator sites in a subsequent polymerization step.[329] The growth of PMMA brushes was monitored in situ via quartz crystal microbalance (QCM) and provided a film thickness of 290 nm PMMA which corresponded with a molecular mass of 137,500.[331] In addition to the functionalization of flat surfaces, the surfaces of nanoparticles of alumina and silica were also successfully modified via photoiniferter, underlining the high versatility of photoiniferter procedures for any surface modification.[332-334]

#### **Insert Figure 40.**

The investigation of the kinetics of photoiniferter surface grafting showed a linear evolution of film thickness with time only for short reaction times, followed by a decrease in polymerization rate indicating the presence of termination events.[335] This behavior is expected for any surface-initiated RDRP.[336] The addition of a diethyldithiocarbamate functionality was suggested to suppress chain transfer events and to improve the polymerization control.[334, 337, 338] The effect of diethyldithiocarbamates on the grafting kinetics was investigated[339] and the reinitiation ability dependence of the concentration of the diethyldithiocarbamate tetraethylthiuram disulfide was modeled in a subsequent study.[340] Furthermore, it was suggested that an increase the concentration of deactivating species could improve the livingness.[335] Also the influence of the light intensity was investigated. Similar to photoATRP, a linear dependence of the grafting rate on the monomer concentration was identified. A methodology was developed to generate patterns with gradient film thicknesses from silicon substrates.[341]

Temperature- and pH-responsive polystyrene particles were prepared using a photoiniferter to graft pNIPAAm and p(NIPAAm-co-acrylic acid) and pNIPAAm-*b*-p(NIPAAm-co-acrylic acid) as well as p(NIPAAm-co-acrylic acid)-*b*-pNIPAAm diblock copolymers, respectively, from polystyrene particles featuring DC-initiators sites.[342, 343] Temperature, pH and ionic strength-responsive diblock copolymer brushes of PMAA and PNIPAAm were grafted from silicon

substrates.[344] The ellipsometric film thickness of the diblock copolymer brushes on silicon substrate was expanded by factor up to 4 via changes in pH and temperature or addition of ions with respect to the dry film thicknesses. Surface-bound PNIPAAm blocks show LCST behavior over a temperature range of around 10 °C, in contrast to a sharp transition in solution. When a disulfide-containing photoiniferter initiator was employed to form a SAM and subsequently graft pNIPAAm from gold substrates, a significantly wider temperature range was not observed for LCST behavior.[345] In a further study, the topographic changes upon pH changes were characterized via atomic-force microscopy (AFM).[346] Very recently, conductive polymer brushes based on polythiophene were reported.[347] 3-Methylthienyl methacrylate was grafted via a photoiniferter mechanism and was then subsequently oxidized and doped with iodine to give a conductive polythiophene polymer parallel to the polymethacrylate backbone (Figure 41). Film thicknesses up to 290 nm were reported, and the conductivity decreased by factor 7 within one hour after doping, indicating oxidation.

**Insert Figure 41.**

Hydrophilic,[348, 349] ionic,[350] and carboxybetaine[351-355] monomers were employed for the preparation of antifouling coatings on gold, polymers and silicon substrates via photoiniferter grafting from reactions. Furthermore, monomers based on amino acids were also reported in photoiniferter grafting polymerization.[356, 357] Protein immobilization via post-polymerization EDC coupling and controlled drug release systems based on drugs encapsulated in block copolymer brushes were reported.[348] In one example bovine albumin protein was attached to a dithiocarbamate-initiator on glass slides.[358] Subsequently poly(dimethylacrylamide) and polyNIPAAm brushes were grown from the surface. Brushes featuring bovine albumin reached heights of 400 to 500 nm. The combination of poly(carboxybetaine) brushes with long brushes for specific protein antibody binding enabled access to antifouling films with high loading binding characteristics.[353-355] A similar approach based on polyPEGMA was reported.[349]

Another example of selective particle functionalization was synthesis of SiO<sub>2</sub> particles functionalized with multiple layers of molecular imprinted polymers, which were subsequently successfully employed for the separation of D- and L-phenylalanine anilide in affinity column chromatography.[359] Recently, MIPs for the anti-inflammatory drug fenbufen were generated using a photoiniferter, leading to potential applications in analytical devices.[360] Likewise, hybrid-MIPs for the identification and subsequent oxidation of catechol were produced using

photoiniferter.[361] Unrelated to that, photoiniferters were also employed to produce microfluidic devices, underpinning again the versatility of the photoiniferter methodology.[362]

Recently, initiator-free photoRAFT was employed to graft PMA brushes from crosslinked pEGDA particles.[208] An initiation efficiency of 95% was determined for the embedded trithiocarbonate RAFT agent. No additional photoinitiator was used since radicals were directly generated from the transfer agent 2-(((butylthio)carbonothiolyl)thio)propanoic acid upon exposure to blue LED-light. PhotoRAFT was employed to graft PAA[363, 364] and PHEMA[365] from polypropylene microporous membranes, although only a few repeating units were attached in the presence of chain transfer agent in solution. Subsequently, a PNIPAAm block was grafted via thermal RAFT to yield block copolymers.[363, 365] Further control experiments were not performed. These modifications of the surface of membrane were conducted by initially exposing the surface to high power UV light in the presence of benzophenone which subsequently formed surface-bound semipinacol radicals to initiate the polymerization in the presence of a chain transfer agent.

The potential of photoRAFT for the production of MIP-SiO<sub>2</sub> composite beads was demonstrated.[366] Photoiniferter and photoRAFT, activated with a surface-bound azoinitiator in the presence of a RAFT agent, were compared. PhotoRAFT and photoiniferter surface functionalization were carried out at 15 °C. L-Phenylalanine anilide was imprinted on a poly(MAA-co-EGMA) matrix. A loss of sulfur concentration was observed for both systems, indicating a loss of control potentially due to recombination events. However, the same result was found in an earlier study where photoiniferter was reported to allow successful formation of various layers.[359] In contrast to a photoiniferter procedure, the photoRAFT system maintained the pore size distribution of the original silica particles.[366] Furthermore, particles modified with the photoRAFT featured a more homogenous polymer film. With respect to enantioselectivity and retention in column chromatography, photoRAFT surface functionalization was superior to photoiniferter grafting. Silica particles functionalized via photoiniferter grafting were not able to separate racemic mixtures of phenylalanine anilide, whereas the respective particles, after photoRAFT grafting, featured good separation for film thicknesses beyond 1 nm, which was attributed to formation of more homogenous films. Retention factor and enantioselectivity were linearly dependent on the graft density for photoRAFT, in contrast to the results of photoiniferter grafting from reactions where a bell shaped curve was observed.

### **8.3 Photomediated NMP**

The development of novel photosensitive alkoxyamines extended the use of NMP process in many applications such as photolithography, microfluidic, data storage devices, 3D-printing, biochips, etc. For example, a homogeneous acrylate-based multilayered film with a thickness close to 50 nm and a good spatial resolution could be prepared by using photosensitive alkoxyamines.[367] The living character of the photoinduced NMP enabled reinitiation of the polymerization producing multi-layered objects with nanometric scale thickness. The photoinduced NMP process was also proposed and developed for fabrication functional micropatterns with complex structures such as hydrophilic/hydrophobic or luminescent surfaces simply by using a mask (Figure 42). [368] The self-healing process was reported in covalent polymer networks via the photodissociation of alkoxyamine junctions.[369]

**Insert Figure 42.**

### **8.4 PhotoCRP in direct laser writing**

Direct laser writing (DLW) with photoinduced controlled radical polymerizations was first reported for photomediated NMP.[367] Line widths of 1.5  $\mu\text{m}$  were achieved with a 375 nm laser (16 mW) and an inverted microscope objective with piezoelectronically positioned sample. Polymeric microstructures of up to 2  $\mu\text{m}$  height were grafted from the surface of a crosslinked polymer film featuring suitable alkoxyamine groups with only 30 ms irradiation time. The structure height was proportional to the irradiation time for the first 10 ms. Recently, microstructures grafting via direct laser writing on Ir-mediated photoATRP was reported (Figure 43).[126] An almost linear dependence of the structure height on the power of the laser employed Nd:Yag microlaser (532 nm) was reported. In contrast to photomediated NMP, Ir-mediated photoATRP required significantly higher laser powers and structure heights were limited to less than 350 nm.

**Insert Figure 43.**

## **9. Summary and perspective**

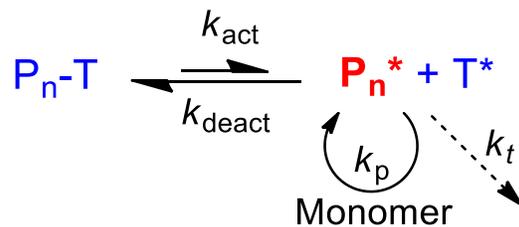
Photomediated controlled radical polymerizations and other processes are versatile methods that allow synthesis of well-defined polymers with predesigned structures. Additionally, the photomediated conditions allow creating specialized materials with interesting features. The photomediated processes can be conducted at ambient or subambient temperatures. Temporal and

spatial controls over the reaction could be easily controlled by switching on/off the irradiation source. Fine tuning of radical formation by controlling light intensity and wavelength leads to precisely controlled individual macromolecules and hybrid polymeric structures due to the exact control of radical concentration. In addition to these advantages of photomediated processes, the use of photochemistry also overcame challenges inherent in thermally mediated reactions. The excited molecules, which absorb the energy of photons, are active catalysts but they could not be generated under thermal conditions. Photoredox-mediated and metal-free ATRP could only be accessed under photochemical conditions.

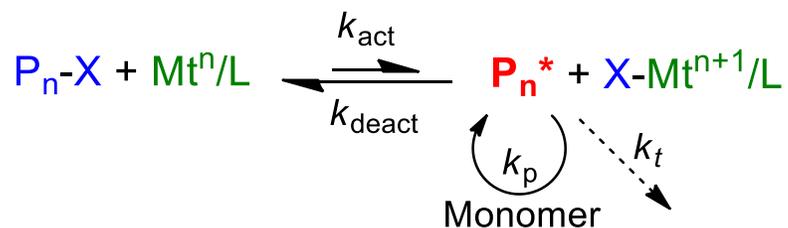
While significant progress has been achieved, future research should target more environmentally friendly and refined systems. For instance, development of catalyst and processes using milder irradiation and longer wavelength sources in the presence of photosensitizers is attractive, since most systems described in this review required high energy irradiation sources (350 – 500 nm). The design of novel catalysts containing more than one distinct catalytic center could provide a selective reactivity under different irradiation conditions. Furthermore, the different reactivity of more sophisticated catalysts might possibly target longer wavelengths and lower light intensity.

**Acknowledgments.** KM and XP acknowledge support from NSF (CHE 1400052). TJ and JL are grateful for support from the Flemish Fonds Wetenschappelijk Onderzoek (FWO).

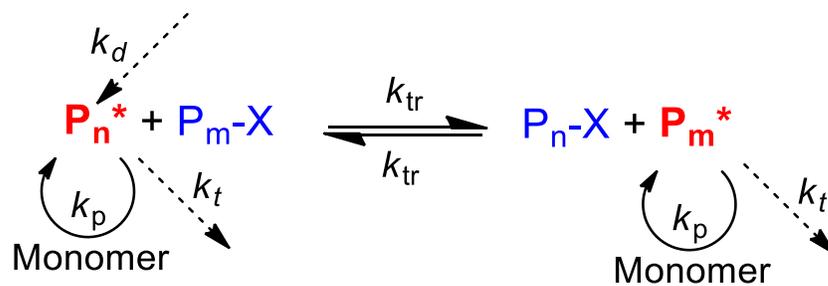
a) Reversible deactivation by coupling



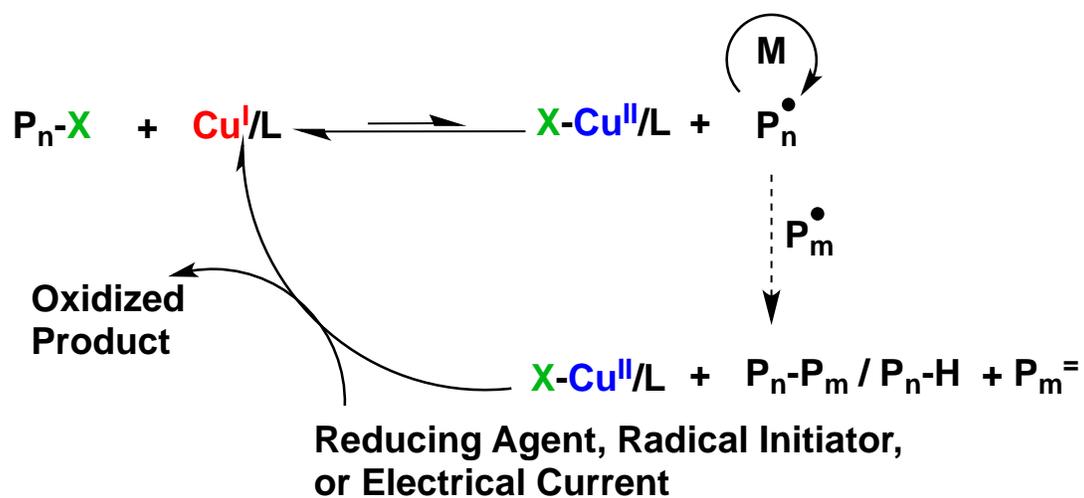
b) Catalytic reversible deactivation by atom transfer



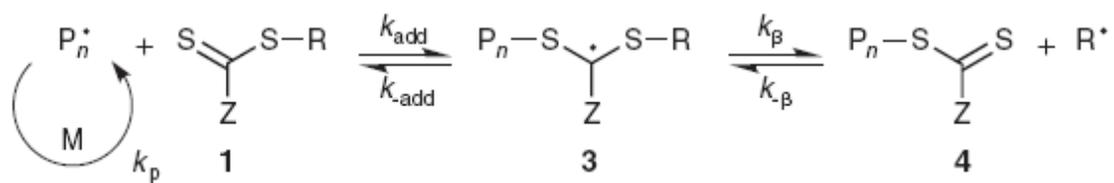
c) Degenerative transfer



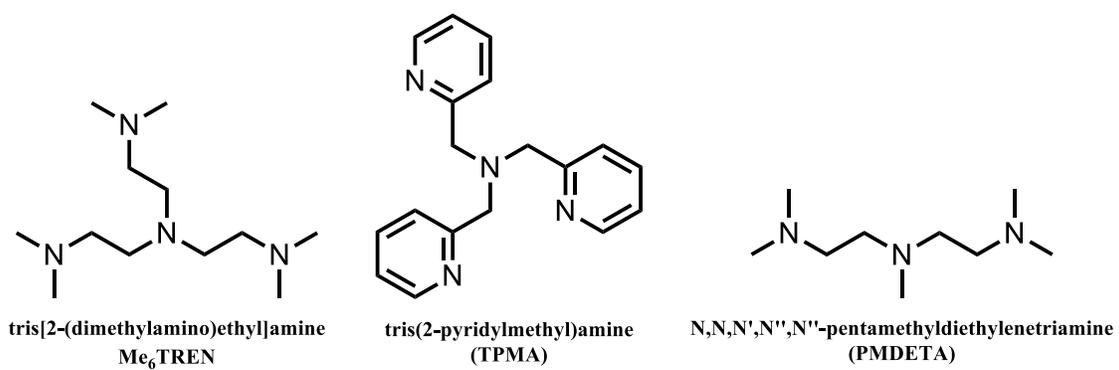
**Scheme 1.** Three general types of reversible-deactivation radical polymerization mechanism.



Scheme 2. Representation of ATRP Equilibrium.

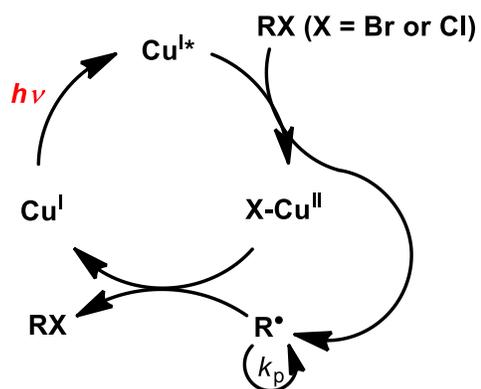


**Scheme 3.** Representation of RAFT Equilibrium.

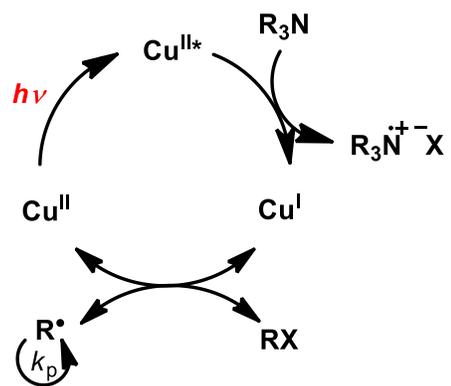


**Scheme 4.** The most commonly used ligands in Cu-mediated photoATRP.

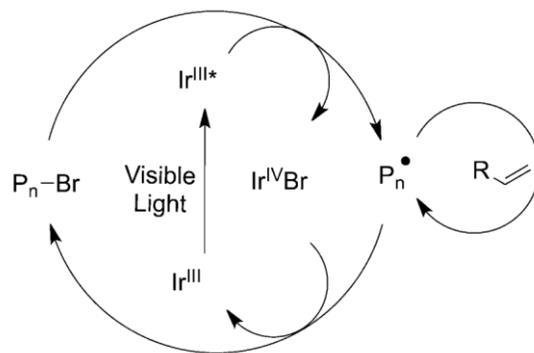
### Oxidative quenching



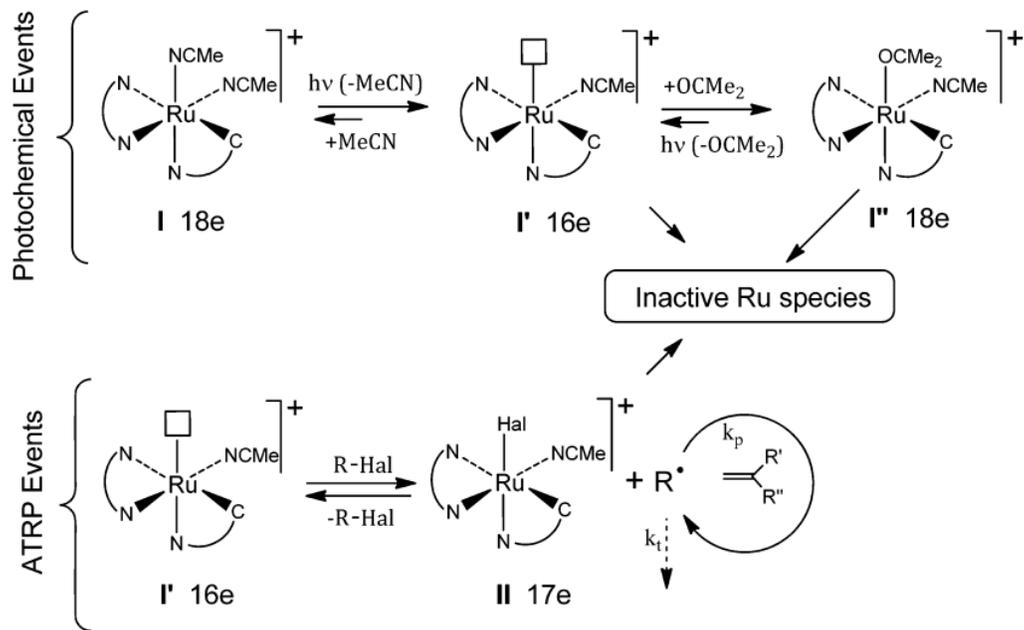
### Reductive quenching



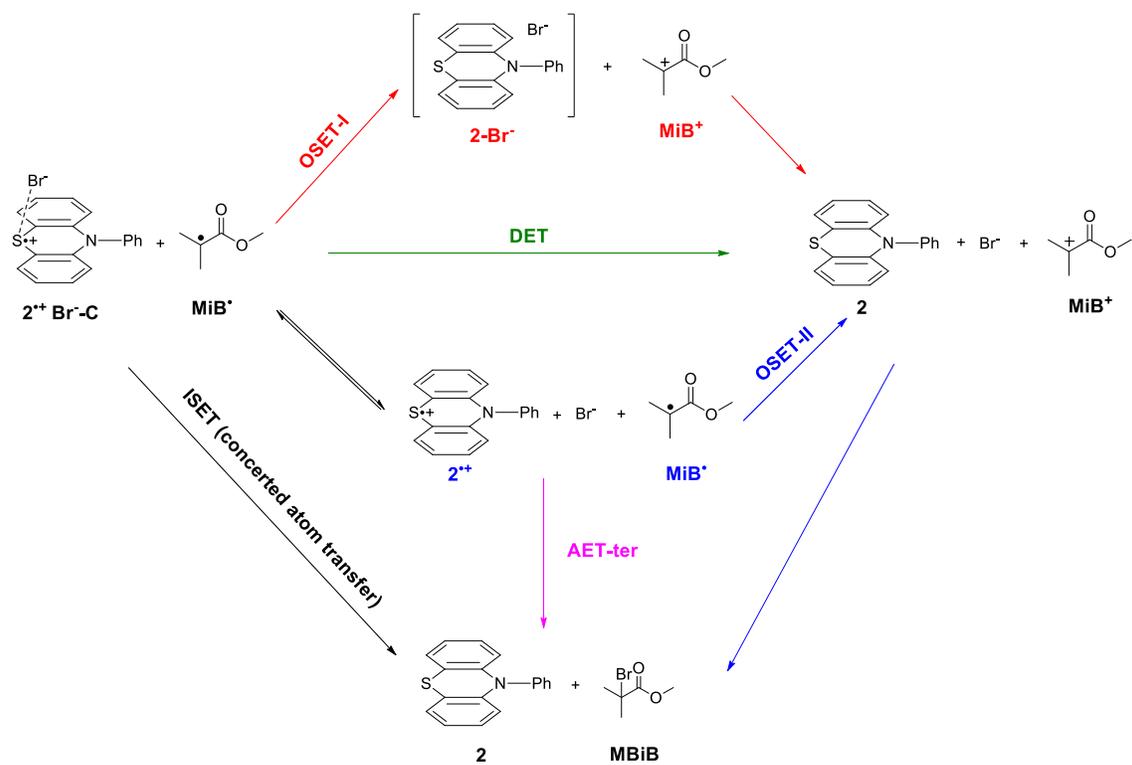
**Scheme 5.** Simplified mechanism for Cu-mediated photoATRP involved in an oxidative quenching cycle ( $[\text{Cu}(\text{phen})_2]\text{Br}$ ) or a reductive quenching cycle ( $\text{Cu}(\text{TPMA})\text{Br}_2$  or  $\text{Cu}(\text{Me}_6\text{TREN})\text{Br}_2$ ).



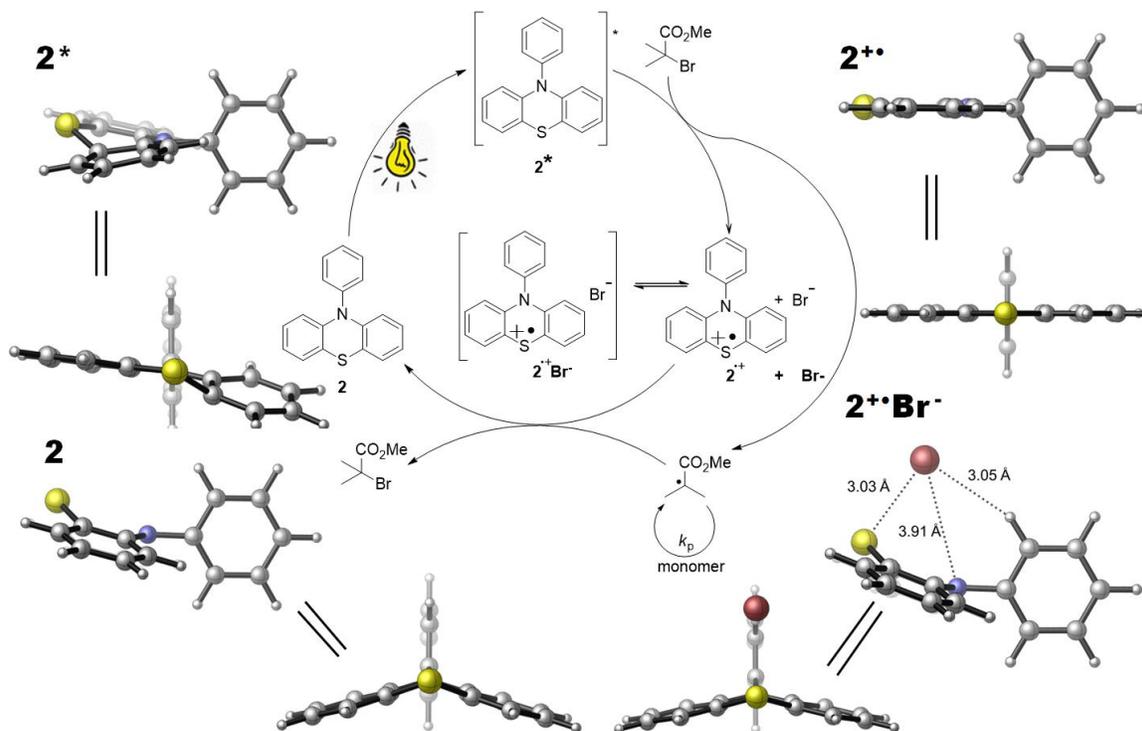
**Scheme 6.** A proposed mechanism of Ir-photoredox catalyzed ATRP by visible light. (Reproduced from Ref. [119], with permission from Wiley.)



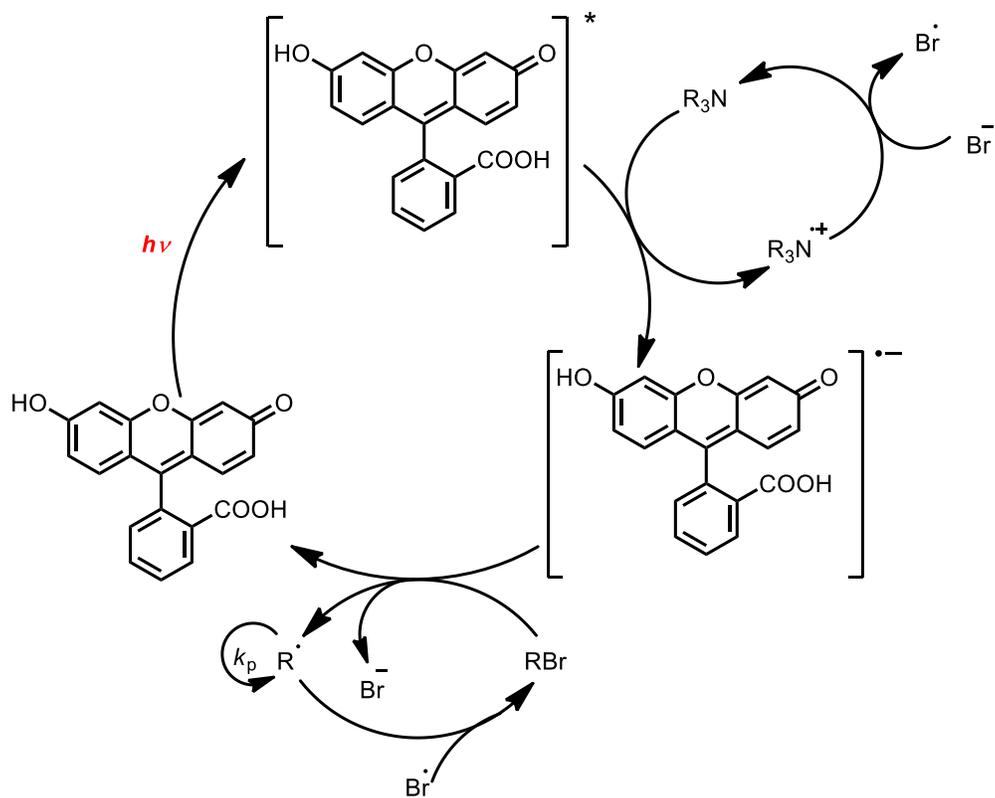
**Scheme 7.** Mechanism proposed for photolabile cyclometalated Ru(II) complex in photoATRP. (Reproduced from Ref. [129], with permission from the American Chemical Society.)



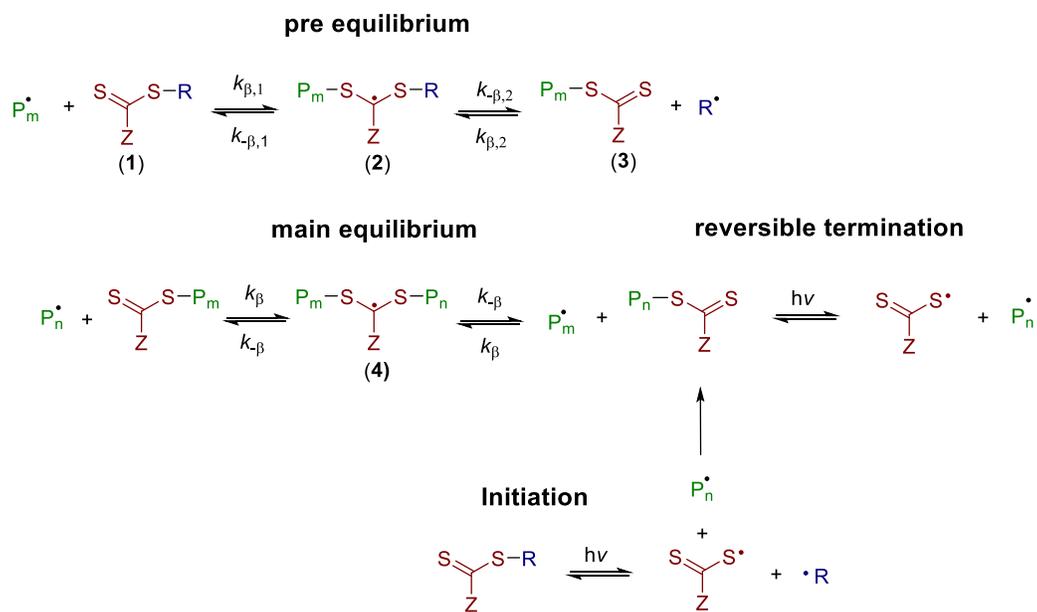
**Scheme 8.** Possible ISET and OSET mechanisms in the deactivation pathway. (Reproduced from Ref. [141], with permission from the American Chemical Society.)



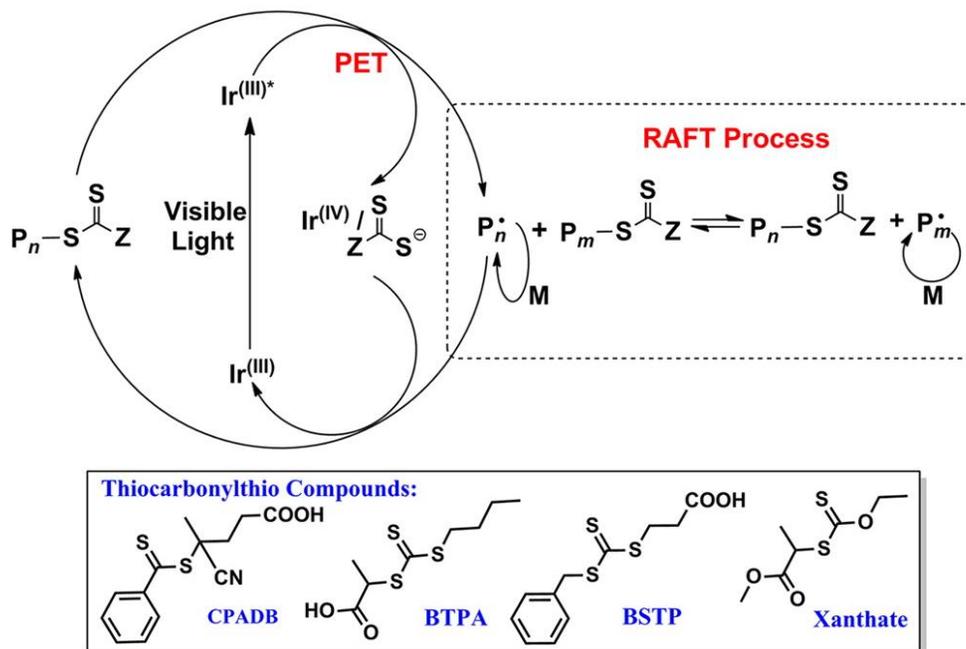
**Scheme 9.** Proposed mechanism for PTZ-catalyzed photoATRP and optimized geometries of related intermediates. (Reproduced from Ref. [141], with permission from the American Chemical Society.)



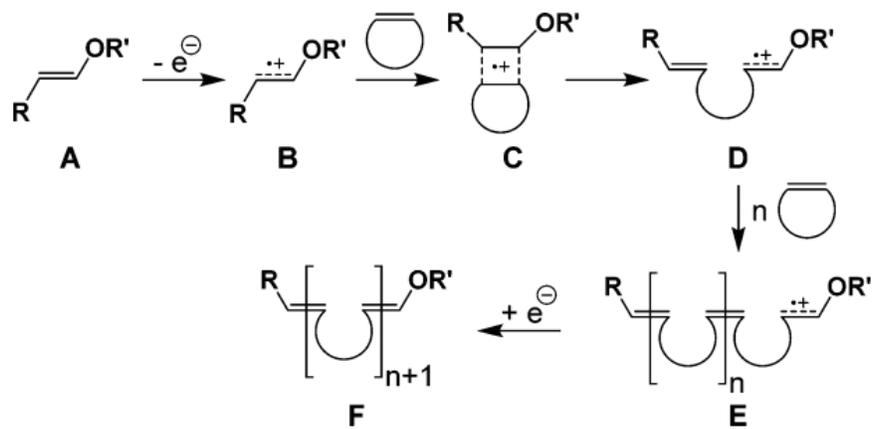
**Scheme 10.** Proposed mechanism of fluorescein-catalyzed photoinduced metal-free ATRP.



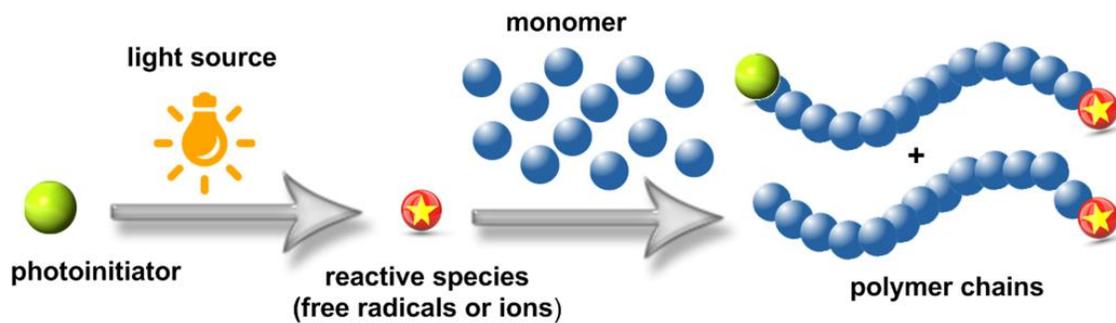
**Scheme 11.** Mechanism of initiator-free photoRAFT polymerization; the iniferter mechanism is superimposed on a classical RAFT process.



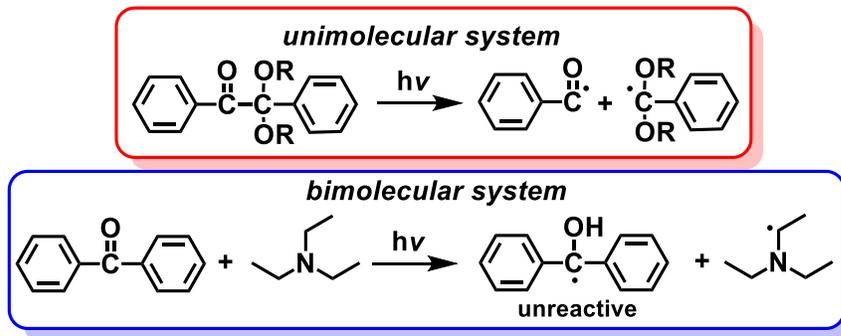
**Scheme 12.** Mechanism of PET-RAFT using a *fac*-[Ir(ppy)<sub>3</sub>] catalyst. Reproduced from Ref. [214], with permission from the American Chemical Society.



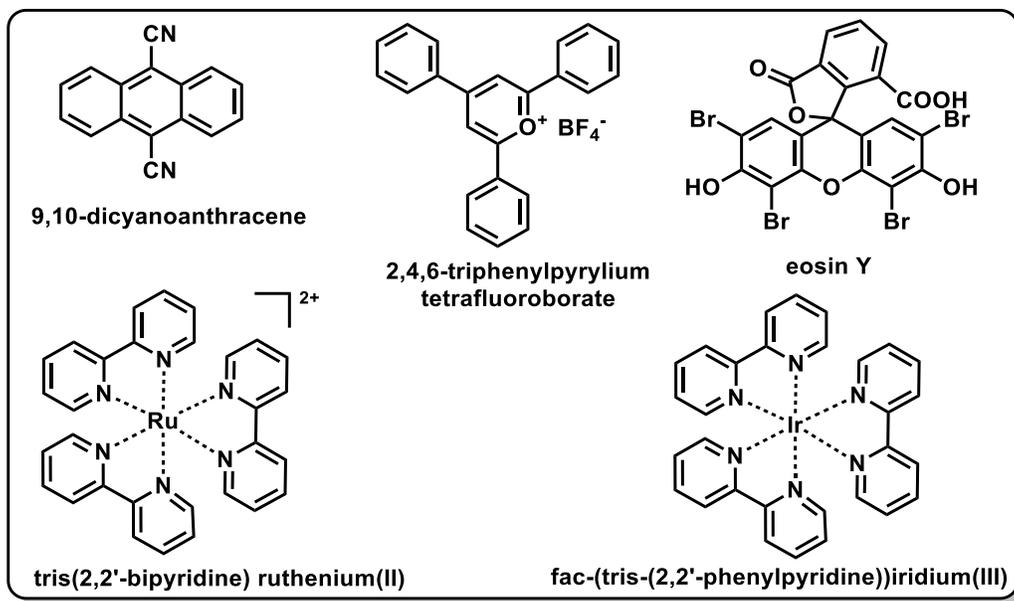
**Scheme 13.** Proposed mechanism of photoredox-mediated metal-free ROMP. (Reproduced from Ref. [284], with permission from the American Chemical Society.)



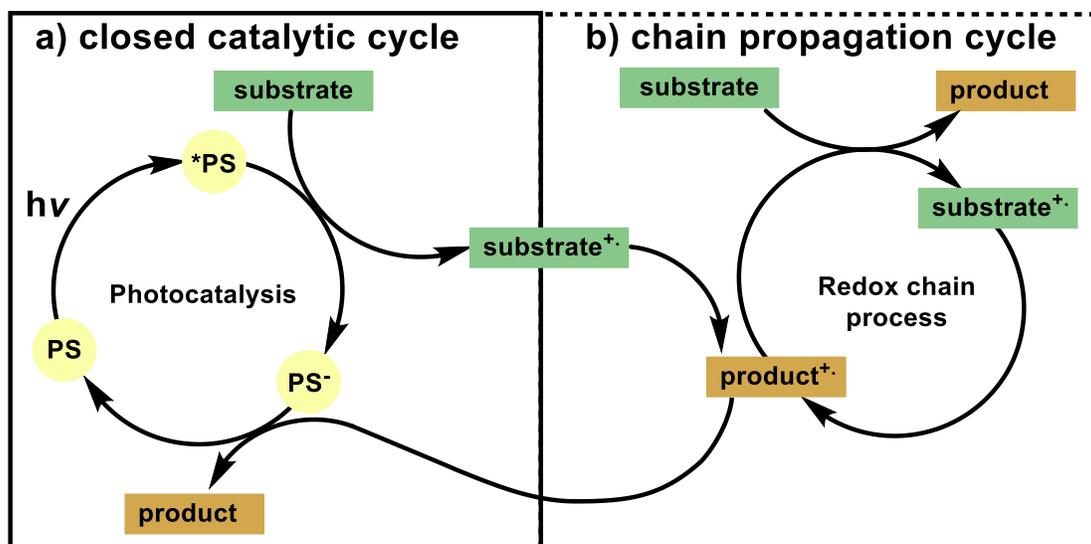
**Figure 1.** General presentation of photoinitiated polymerization.



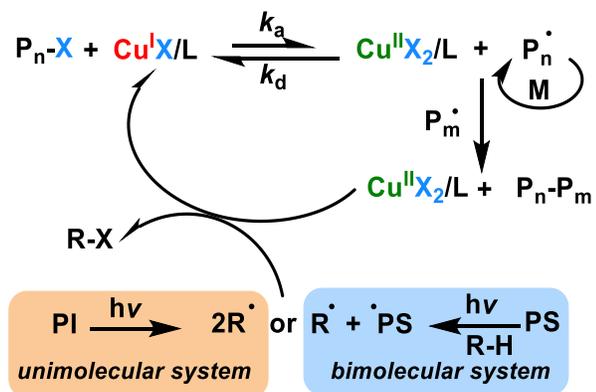
**Figure 2.** Mechanism of unimolecular (upper) or bimolecular (lower) photoinitiated systems.



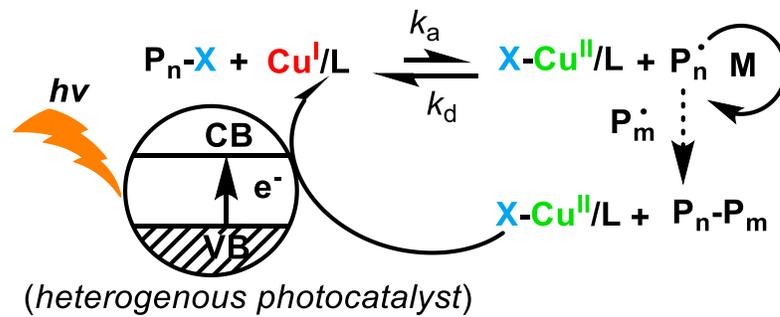
**Figure 3.** Commonly used photoredox catalysts.



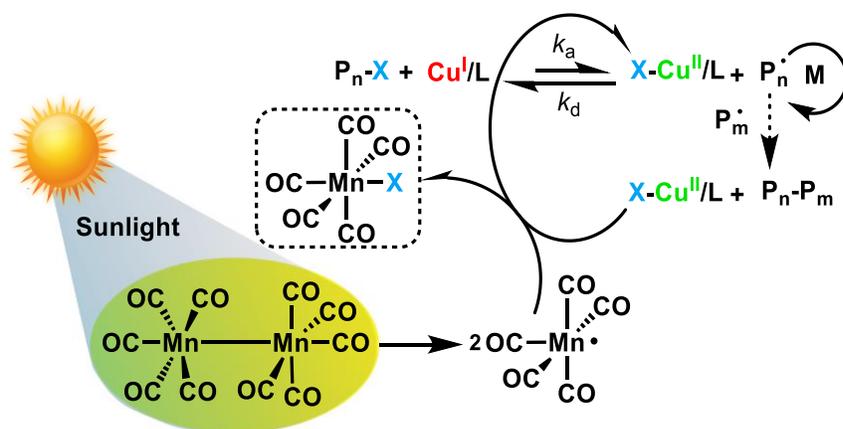
**Figure 4.** General mechanism for oxidative photoredox reactions.



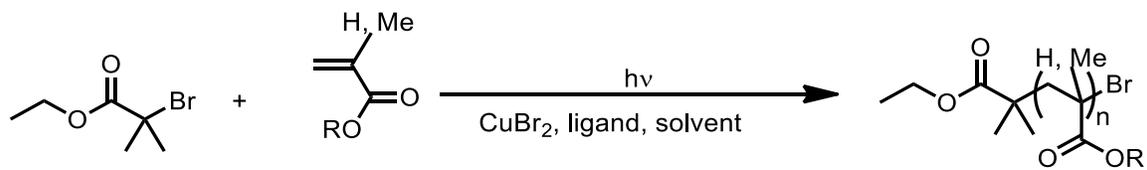
**Figure 5.** Proposed mechanisms for photomediated ATRP using unimolecular and bimolecular photoinitiated systems (PI: photoinitiator; PS: photosensitizer and R-H: hydrogen donor).



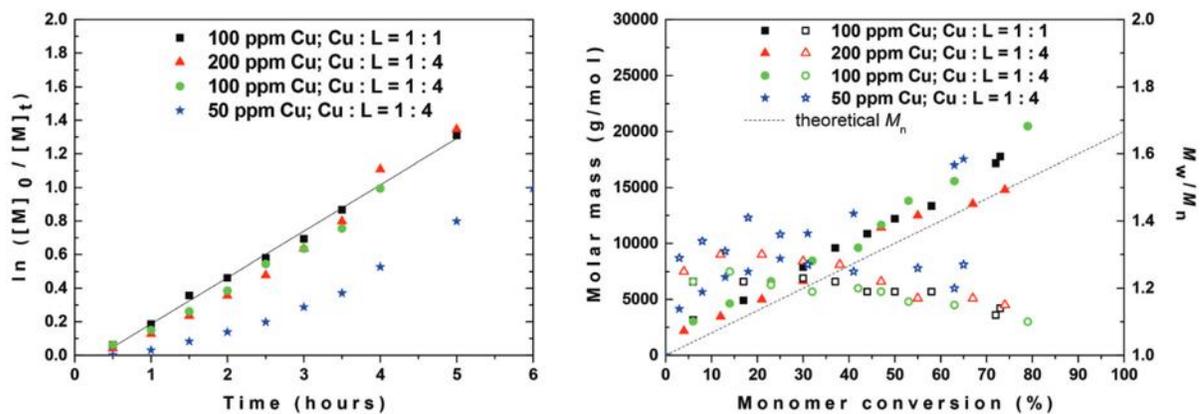
**Figure 6.** Postulated mechanism of photomediated ATRP using a heterogeneous photocatalyst.



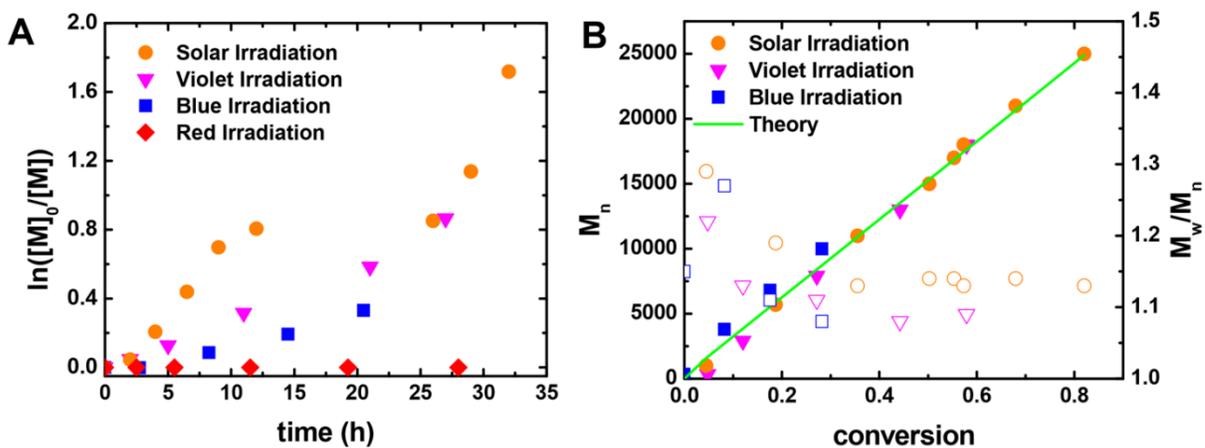
**Figure 7.** Sun-light mediated ATRP using  $\text{Mn}_2(\text{CO})_{10}/\text{Cu}(\text{II})\text{Br}_2/\text{alkyl halide}$  system. Reproduced from Ref. [61], with permission from the Royal Society of Chemistry.)



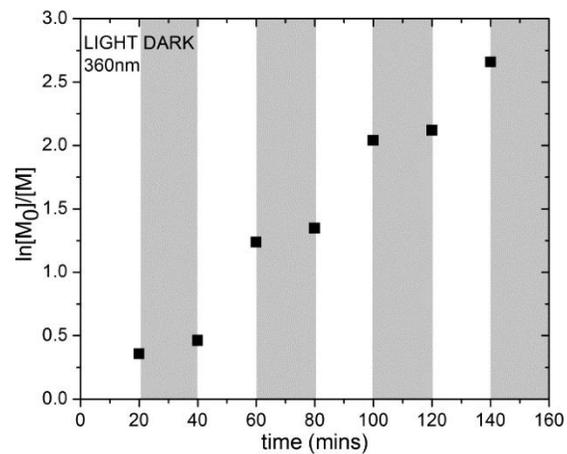
**Figure 8.** General reaction scheme for a photoATRP reaction using copper(II) as catalyst.



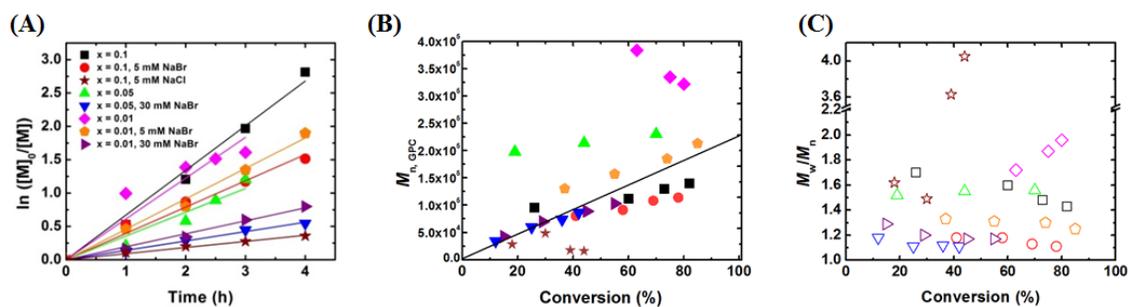
**Figure 9.** Comparison of various copper and ligand concentrations: (l.h.s.) kinetic plot, (r.h.s.) evolution of molar mass and dispersity with conversion. Reactions were carried out in 26 vol% DMSO with  $[MMA]:[BPN]:[CuBr_2]:[TPMA] = 200:1:0.02:0.20$ ;  $[MMA] = 6.9$  M;  $T = 35$  °C;  $\lambda < 350$  nm. (Reproduced from Ref. [74], with permission from The Royal Society of Chemistry.)



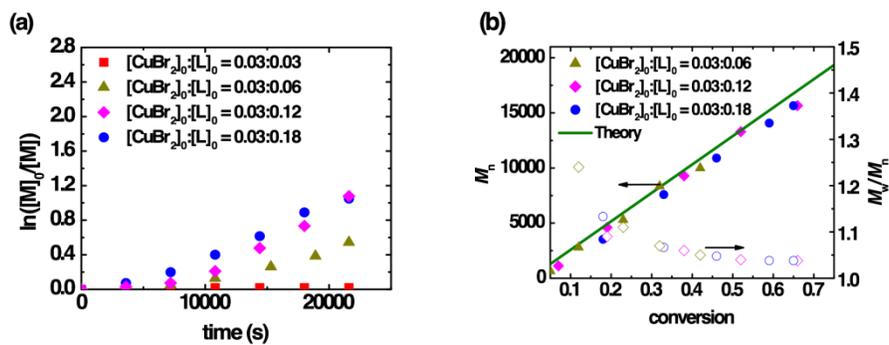
**Figure 10.** (A) Kinetics and (B)  $M_n$  (solid points) and  $M_w/M_n$  (open points) evolution in the polymerization of MA using different radiation sources. Conditions:  $[MA]:[EBiB]:[CuBr_2]:[TPMA^*] = 300:1:0.03:0.135$  in 50 vol% DMF at room temperature. (Reproduced from Ref. [23], with permission from the American Chemical Society.)



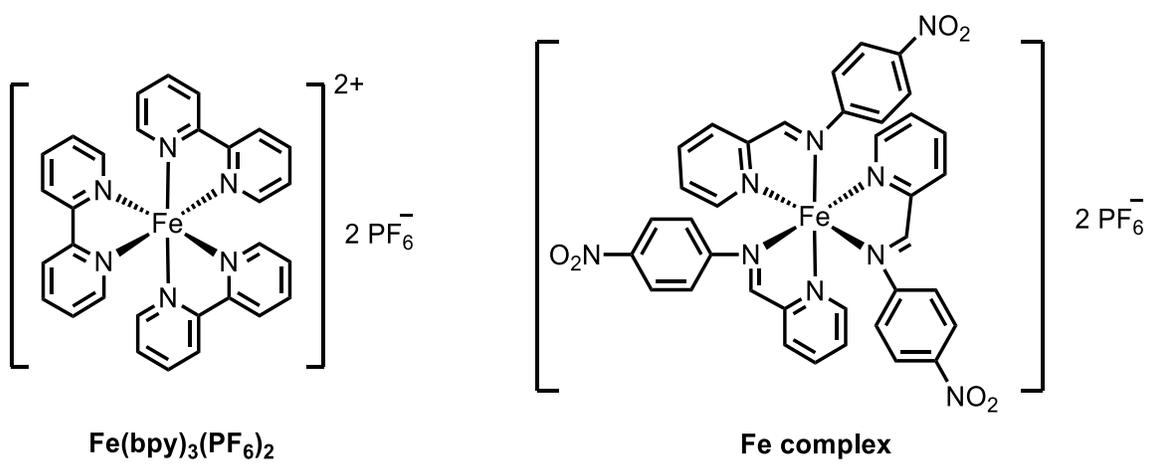
**Figure 11.** Evidence of temporal control via consecutive light (white area) and dark (shaded area) exposure. [M]:[EBiB]:[CuBr<sub>2</sub>]:[Me<sub>6</sub>TREN] = 50:1:0.02:0.12 in DMSO (50%, v/v) solvent. (Reproduced from Ref. [25], with permission from the American Chemical Society.)



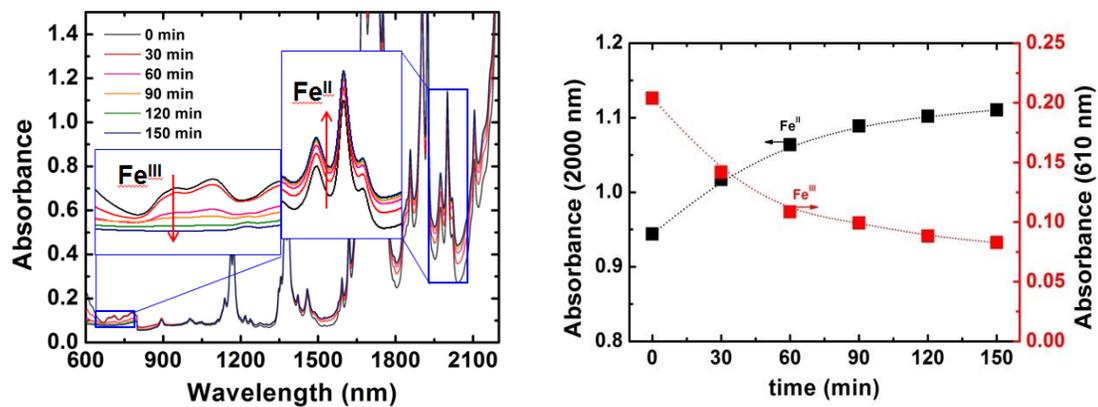
**Figure 12.** Photoinduced ATRP of OEOMA<sub>500</sub> in water with ppm of Cu with or without added salt. (A) Semilogarithmic kinetic plots (B) evolution of  $M_n$  (filled symbols) and (C)  $M_w/M_n$  (open symbols) with conversion, conditions:  $[OEOMA_{500}]_0/[PEO_{2000}BPA]_0/[CuBr_2]_0/[TPMA]_0 = 450/1/x/4x$  in 90% water with 392 nm at 0.9 mW/cm<sup>2</sup> at room temperature. (Reproduced from Ref. [87], with permission from the American Chemical Society.)



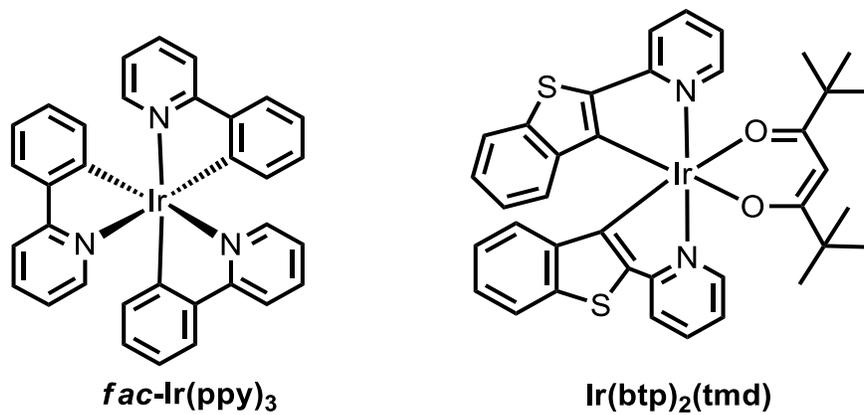
**Figure 13.** (a) Kinetics and (b) evolution of  $M_n$  (filled symbols) and  $M_w/M_n$  (opened symbols) with conversion for photoATRP of MA in DMSO under the conditions:  $[\text{MA}]_0:[\text{EBiB}]_0:[\text{CuBr}_2]_0:[\text{Me}_6\text{TREN}]_0 = 300:1:0.03:0.03\text{-}0.18$ ,  $[\text{MA}] = 7.4 \text{ M}$ , irradiated by 392 nm light ( $0.9 \text{ mW/cm}^2$ ) at  $25 \text{ }^\circ\text{C}$ . (Reproduced from Ref. [66], with permission from the American Chemical Society.)



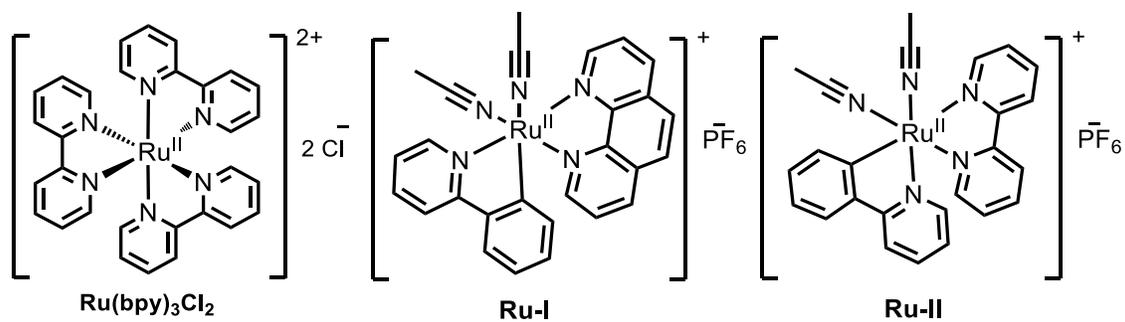
**Figure 14.** Structures of selected iron catalysts in photoATRP.



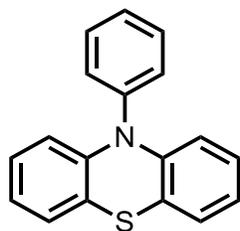
**Figure 15.** Photoreduction of Fe(III) to Fe(II), confirmed by UV-vis-NIR spectroscopic studies of 3 mM FeBr<sub>3</sub> in MeCN in the presence of 100 equiv of MMA at rt with irradiation. (Reproduced from Ref. [100], with permission from the American Chemical Society.)



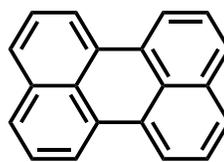
**Figure 16.** Structures of selected Ir catalysts in photoATRP.



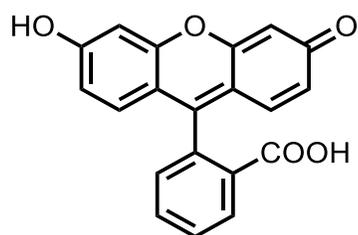
**Figure 17.** Structures of selected Ru catalysts examined for photoATRP.



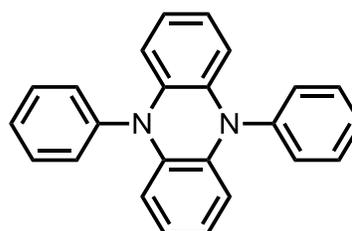
10-phenylphenothiazine



perylene

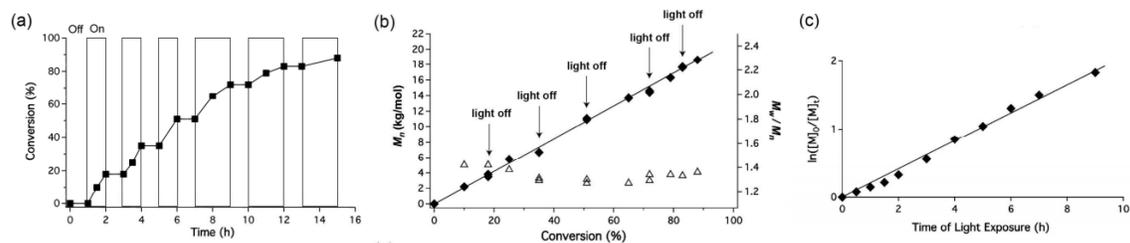


fluorescein

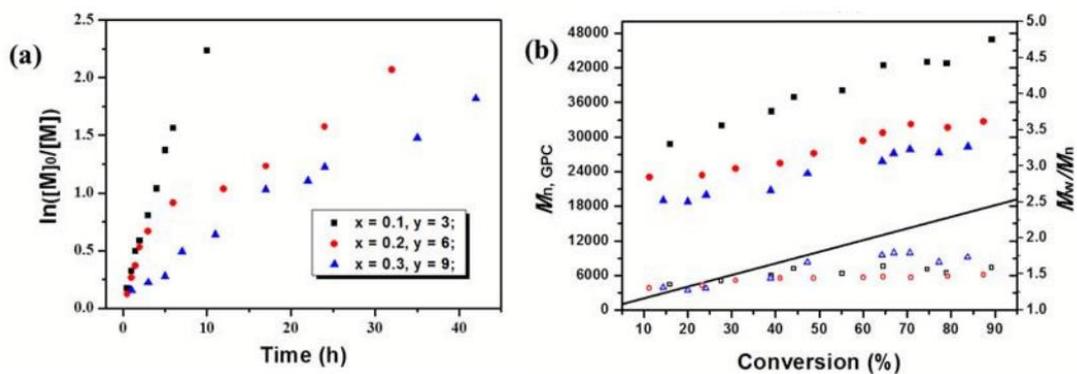


5,10-diphenyl-5,10-dihydrophenazine

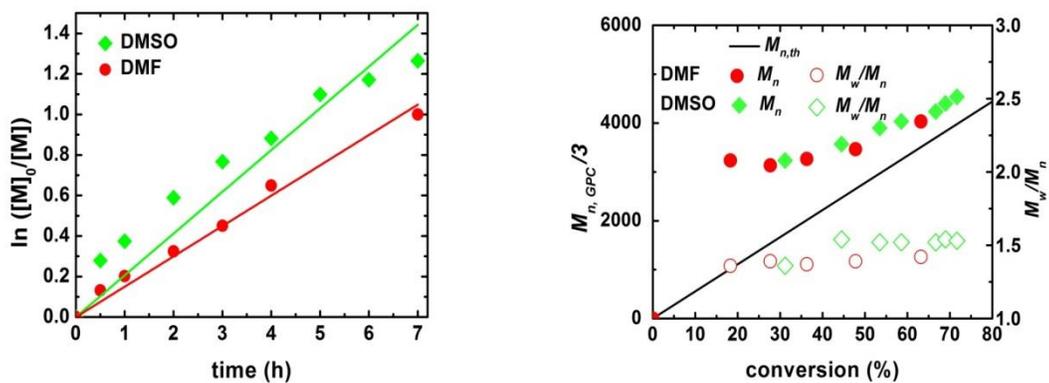
**Figure 18.** Structures of metal-free catalysts in photoATRP.



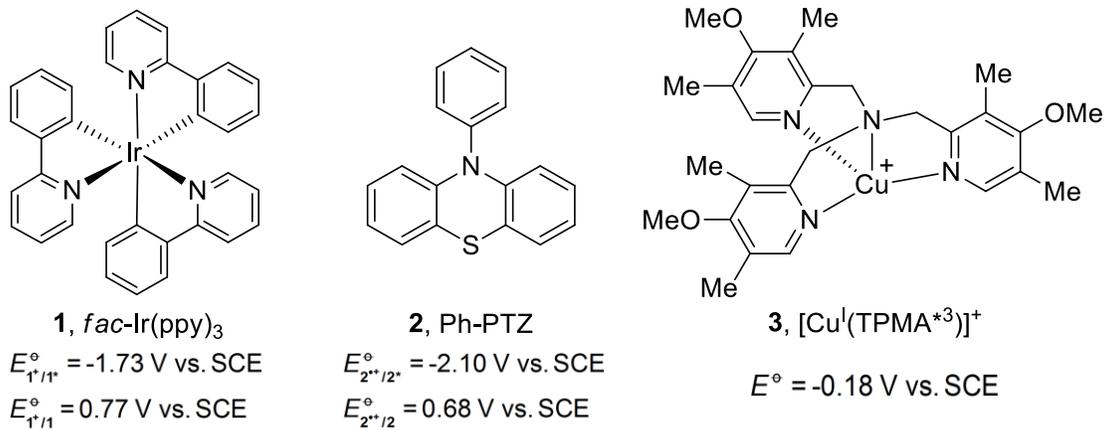
**Figure 19.** PTZ-catalyzed photoATRP of benzyl methacrylate. (a) Conversion during a series of “on-off” cycles of the polymerization to light. (b) Evolution of  $M_n$  (filled symbols) and  $M_w/M_n$  (opened symbols) with conversion. (c) Semilogarithmic kinetic plots of polymerization with irradiation. (Reproduced from Ref. [26], with permission from the American Chemical Society.)



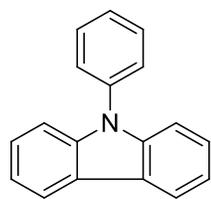
**Figure 20.** (a) Semilogarithmic kinetic plots of polymerization under conditions  $[MMA]_0/[EBPA]_0/[FL]_0/[TEA]_0 = 200/1/x/y$  ( $x = 0.1, y = 3$ , square;  $x = 0.2, y = 6$ , circle;  $x = 0.3, y = 9$ , triangle),  $V_{MMA} = 1.0$  mL,  $V_{DMSO} = 2.0$  mL; irradiation by a blue LED at rt. (b) Evolution of  $M_n$  (filled symbols) and  $M_w/M_n$  (opened symbols) with conversion. (Reproduced from Ref. [133], with permission from The Royal Society of Chemistry.)



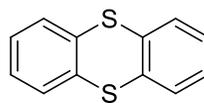
**Figure 21.** (a) Semilogarithmic kinetic plots of polymerization under conditions  $[AN]_0/[EBPA]_0/[Ph-PTZ]_0 = 100/1/0.1$ , in DMSO or DMF; irradiation by a 365 nm ( $4.9 \text{ mW/cm}^2$ ) at rt. (b) Evolution of  $M_n$  (filled symbols) and  $M_w/M_n$  (opened symbols) with conversion. (Reproduced from Ref. [27], with permission from the American Chemical Society.)



**Figure 22.** Structures of photoredox catalysts *fac*-[Ir(ppy)<sub>3</sub>], Ph-PTZ and a traditional Cu-based ATRP catalyst [Cu<sup>I</sup>(TPMA\*<sub>3</sub>)]<sup>+</sup>. (Reproduced from Ref. [141], with permission from the American Chemical Society.)

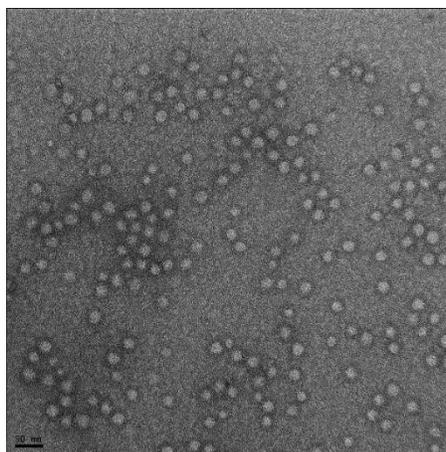


**9-phenylcarbazole**

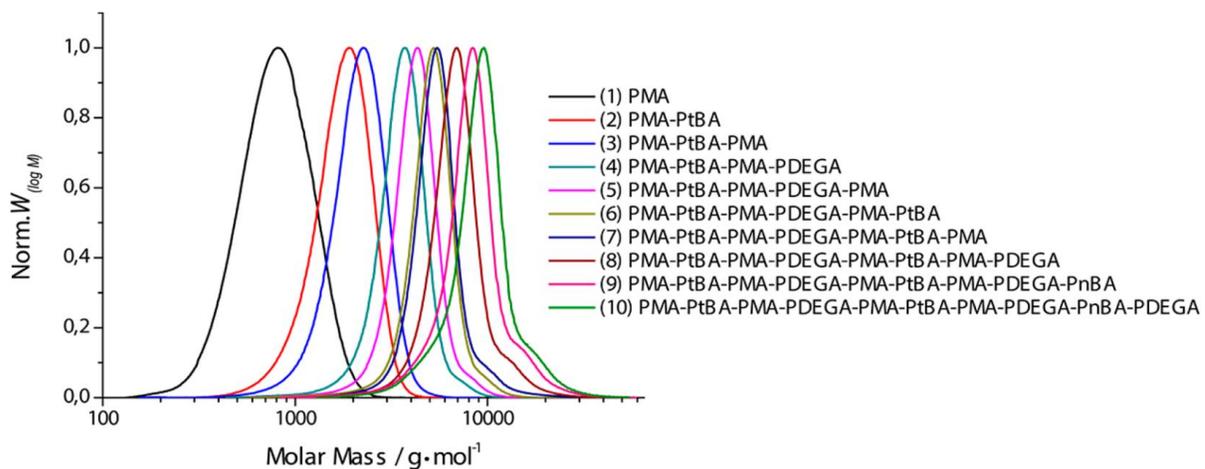


**thianthrene**

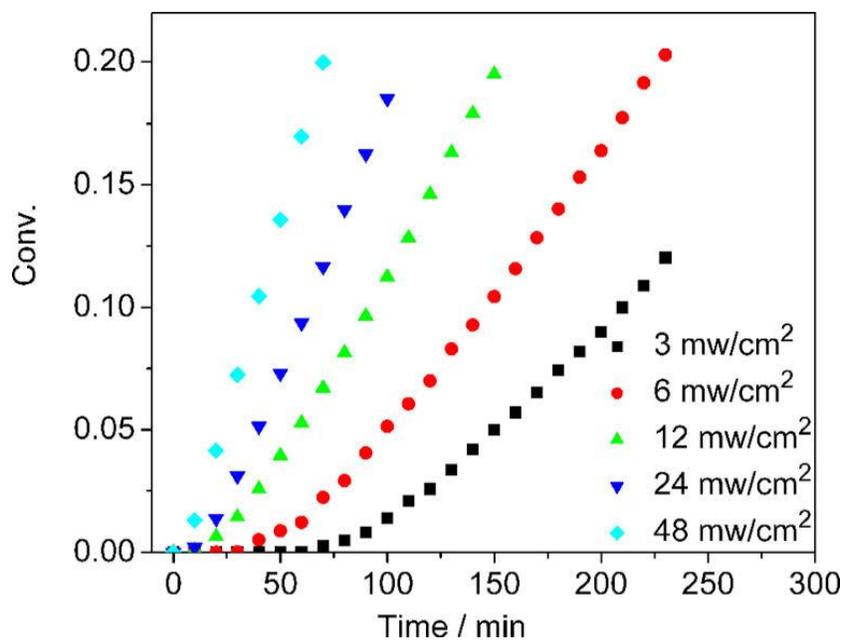
**Figure 23.** Structures of 9-phenylcarbazole and thianthrene.



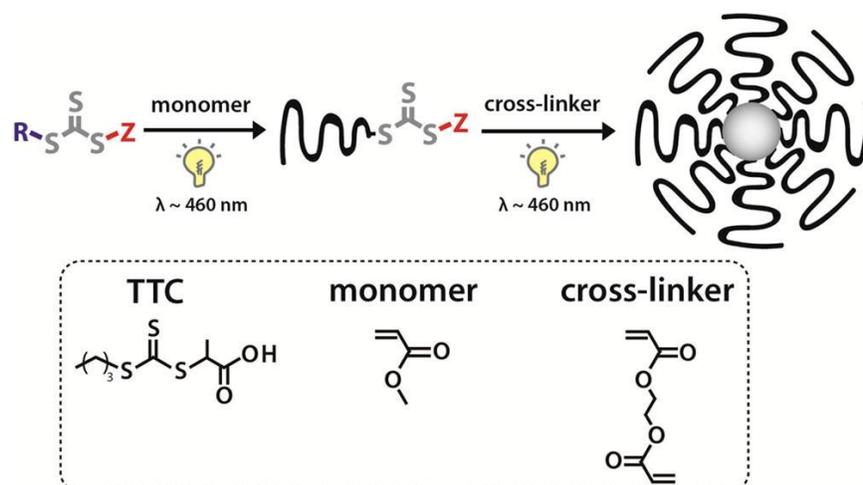
**Figure 24.** TEM images of micellized amphiphilic poly(methyl acrylate)<sub>50</sub>-b-(glycidyl acrylate)<sub>100</sub> block copolymer. Scale bar = 50 nm. (Reproduced from Ref. [69], with permission from the American Chemical Society.)



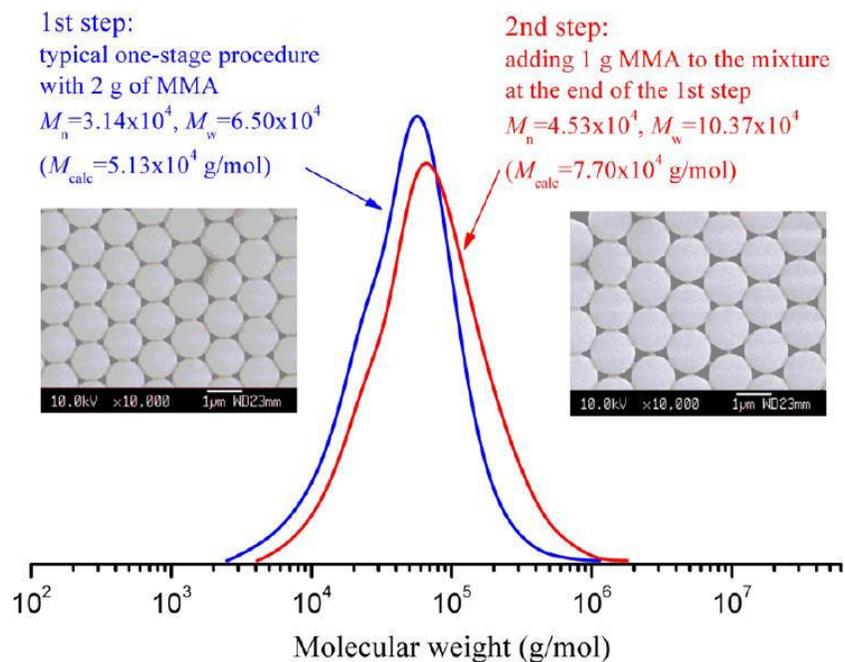
**Figure 25.** Molecular weight distributions of the intermediate block copolymer products in the synthesis of decablock copolymers from sequential monomer addition. (Reproduced from Ref. [70], with permission from the American Chemical Society.)



**Figure 26.** Effect of long-wave length irradiation ( $\lambda > 320$  nm) irradiation intensity on polymerization of methyl acrylate mediated by DDMAT at 25 °C. Reproduced from Ref. [206], with permission from the American Chemical Society.



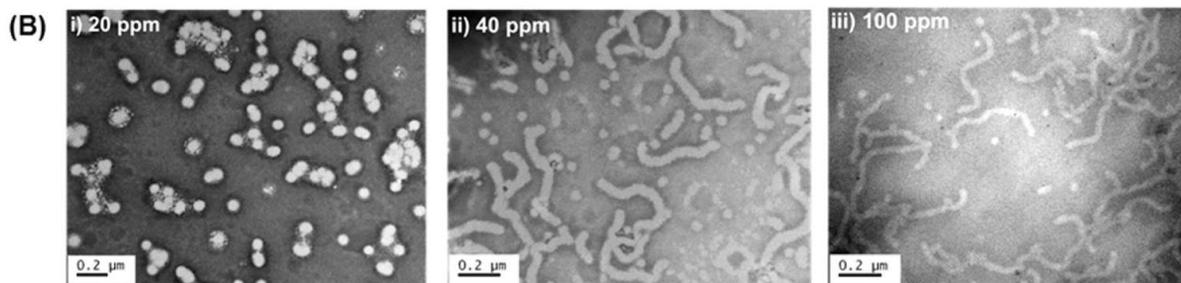
**Figure 27.** Formation of star polymers by an arm-first approach, using initiator-free photoRAFT to activate chains. Reproduced from Ref. [209], with permission from the American Chemical Society.



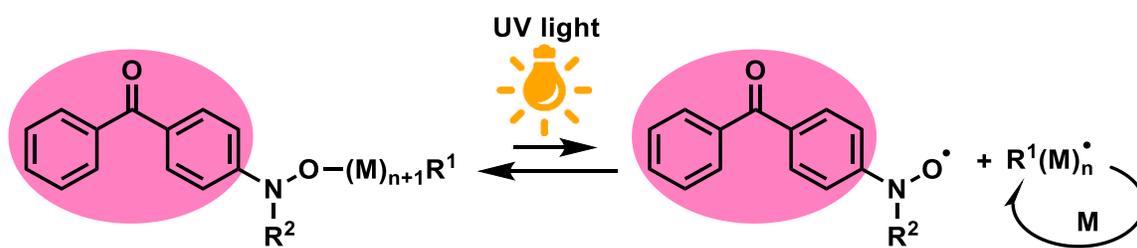
**Figure 28.** GPC curves for PMMA produced in a two-step process: The first step proceeded via a one-stage photoinitiated RAFT dispersion polymerization with 2 g of MMA. After 3 h irradiation, 1 g of MMA was added to the mixture and the reaction continued for further 3 h. Reproduced from Ref. [237], with permission from the American Chemical Society.

(A)

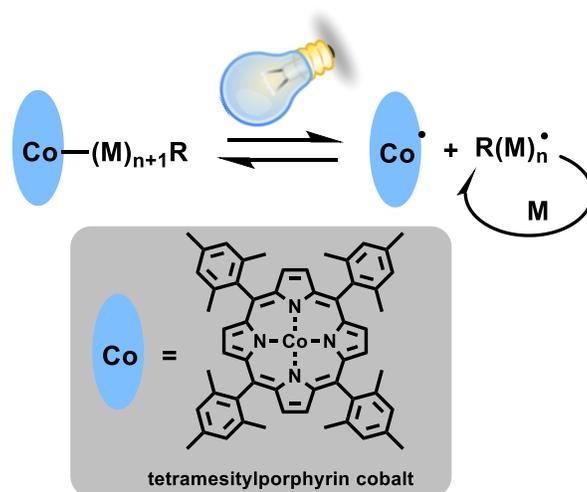
Exp.	Catalyst (ppm)	Conversion (%)	$M_{n, \text{theo}}$ (g/mol)	$M_{n, \text{NMR}}$ (g/mol)	$M_{n, \text{GPC}}$ (g/mol)	$\bar{D}$	$D_h$ (DLS) (nm)	PDI (DLS)
i	20	41	23 400	24 100	16 800	1.19	46	0.13
ii	40	50	26 600	25 900	19 600	1.21	134	0.20
iii	100	61	30 500	28 600	22 600	1.24	140	0.17



**Figure 29.** (A) Characterization of POEGMA-*b*-PBzMA diblock copolymers synthesized using (i) 20, (ii) 40, and (iii) 100 ppm catalyst relative to monomer concentration and a [BzMA]/[POEGMA] ratio of 200:1; (B) TEM images of self-assembled POEGMA-*b*-PBnMA nanoparticles (Exp. i, ii, and iii). Reproduced from Ref. [245], with permission from the American Chemical Society.

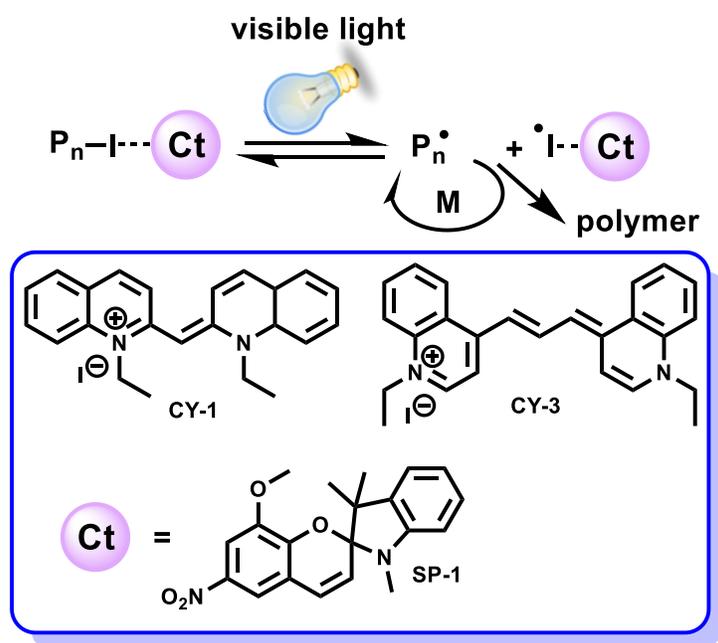


**Figure 30.** Photomediated NMP using a photosensitive alkoxyamine.

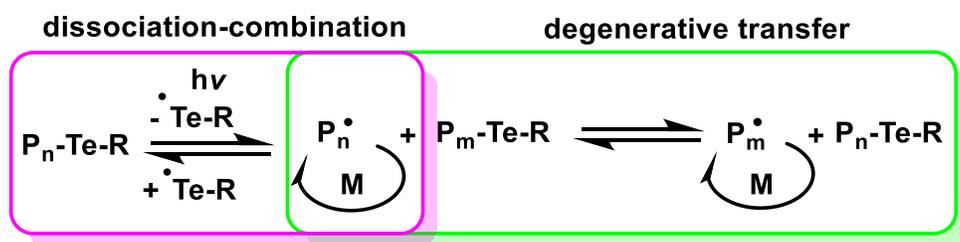


**Figure 31.** General mechanism of photomediated CMRP using organocobalt porphyrin as a control agent.

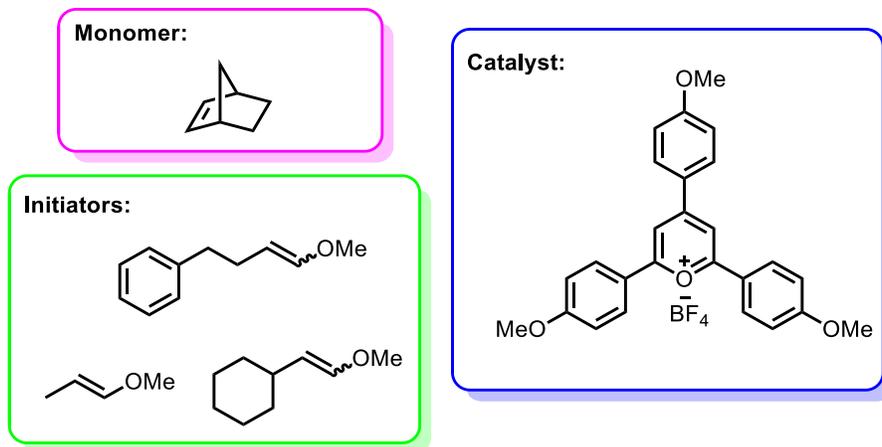




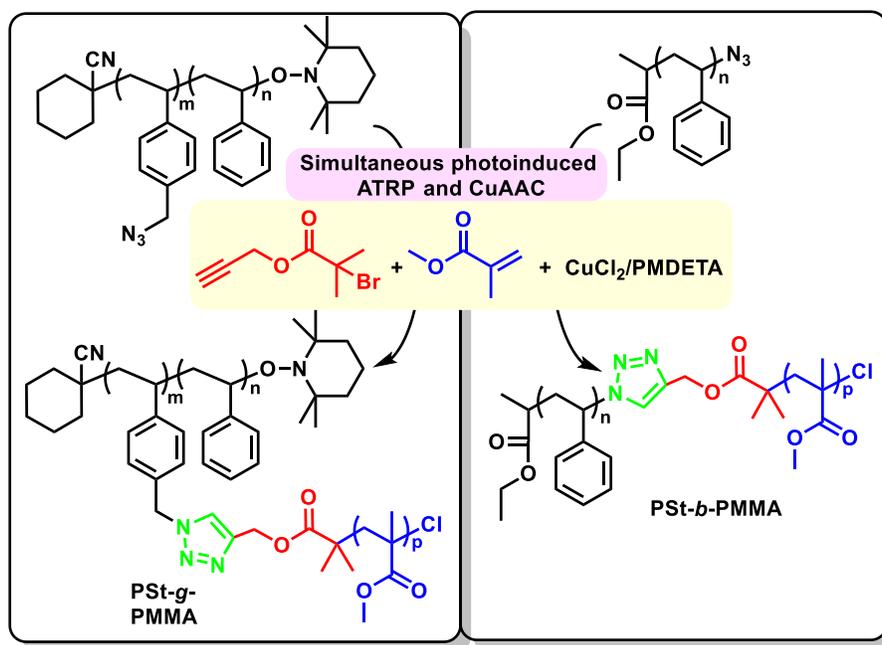
**Figure 33.** Proposed mechanism of photomediated ITP based on reversible complexation of alkyl iodides with various organic dyes.



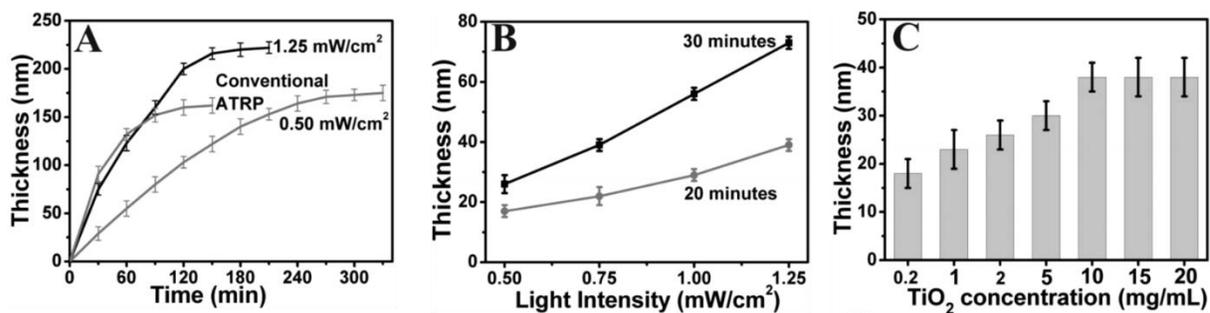
**Figure 34.** General scheme of photomediated TERP based on direct homolysis of C-Te bond.



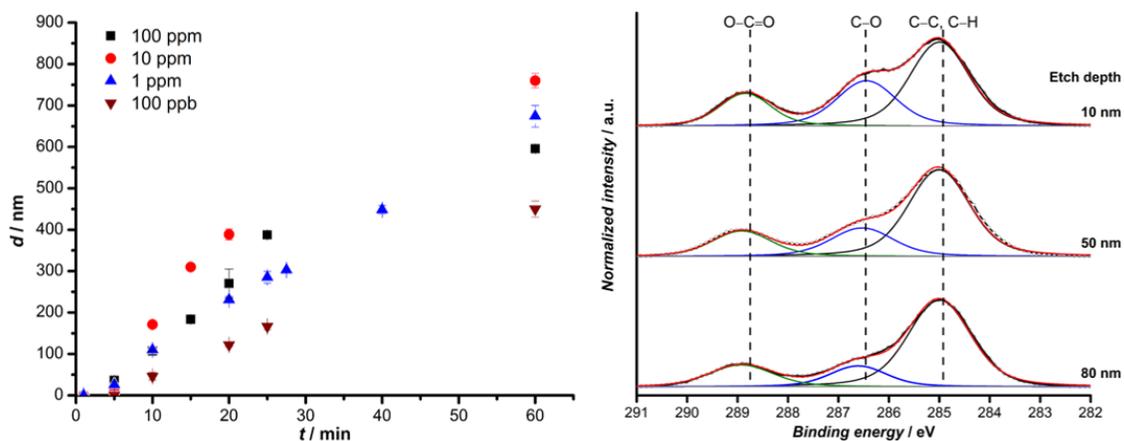
**Figure 35.** Structure of monomers, initiators and catalyst used in photoinduced metal-free ROMP.



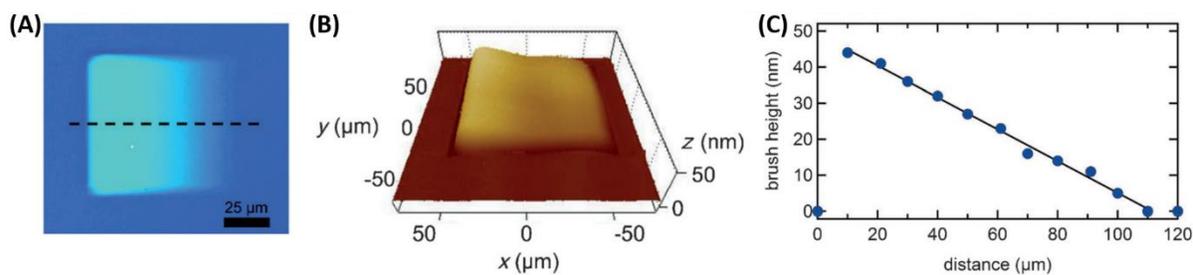
**Figure 36.** Synthesis of block and graft copolymers via simultaneous photoinduced ATRP and CuAAC reactions.



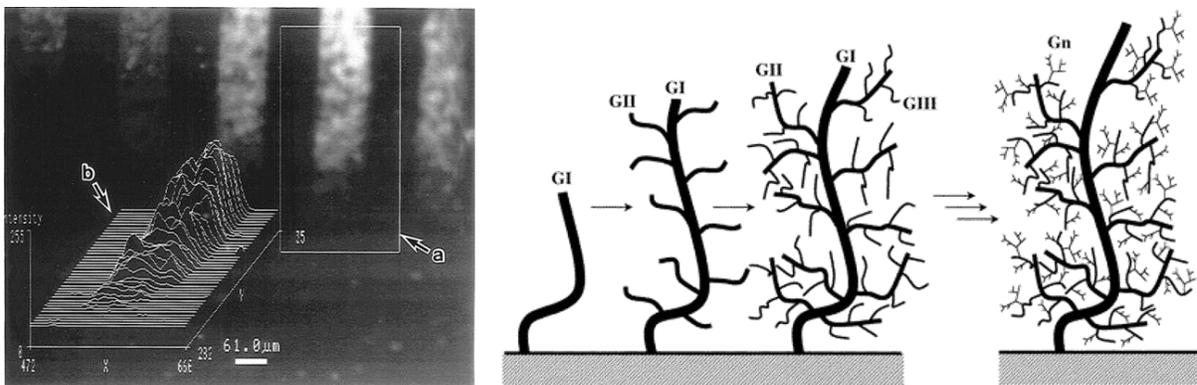
**Figure 37.** PhotoATRP grafting of 3-sulfopropyl methacrylate from thiol-patterned gold substrate. (A) Film thickness vs. time evolution at different light intensities in comparison with conventional ATRP. (B) Comparison of film thicknesses for various light intensities for exposure times of 20 and 30 min, (C) Evolution of film thickness with TiO<sub>2</sub> particle concentration at 1.25 mW/cm<sup>2</sup> illumination for 20 min. (Reproduced from Ref. [60], with permission from the American Chemical Society.)



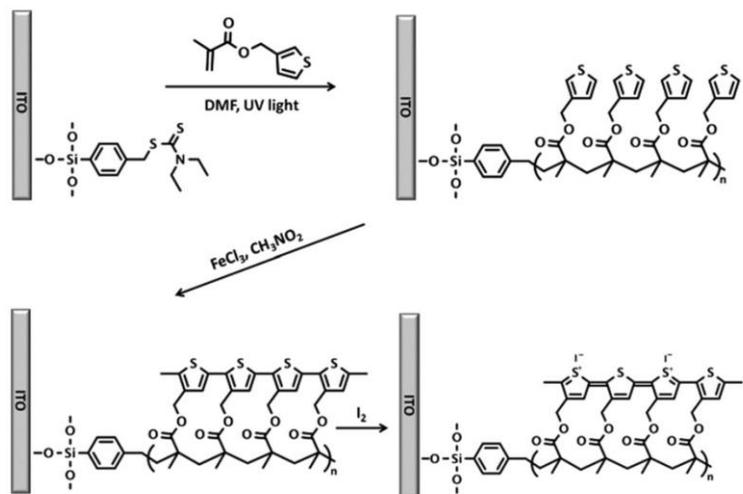
**Figure 38.** (l.h.s) Film thickness vs. time evolution of copper-mediated photoATRP grafting of poly(*t*-butyl acrylate) brushes at various copper catalyst loadings (ppm with respect to monomer). (r.h.s.) XPS depth profiling of poly(*t*-butyl acrylate)-*b*-poly(2-hydroxyethyl acrylate) brushes. A composition drift from preferentially poly(2-hydroxyethyl acrylate) at the polymer/air interface to almost pure poly(*t*-butyl acrylate) at the polymer silicon interface. (Reproduced from Ref. [76], with permission from Wiley.)



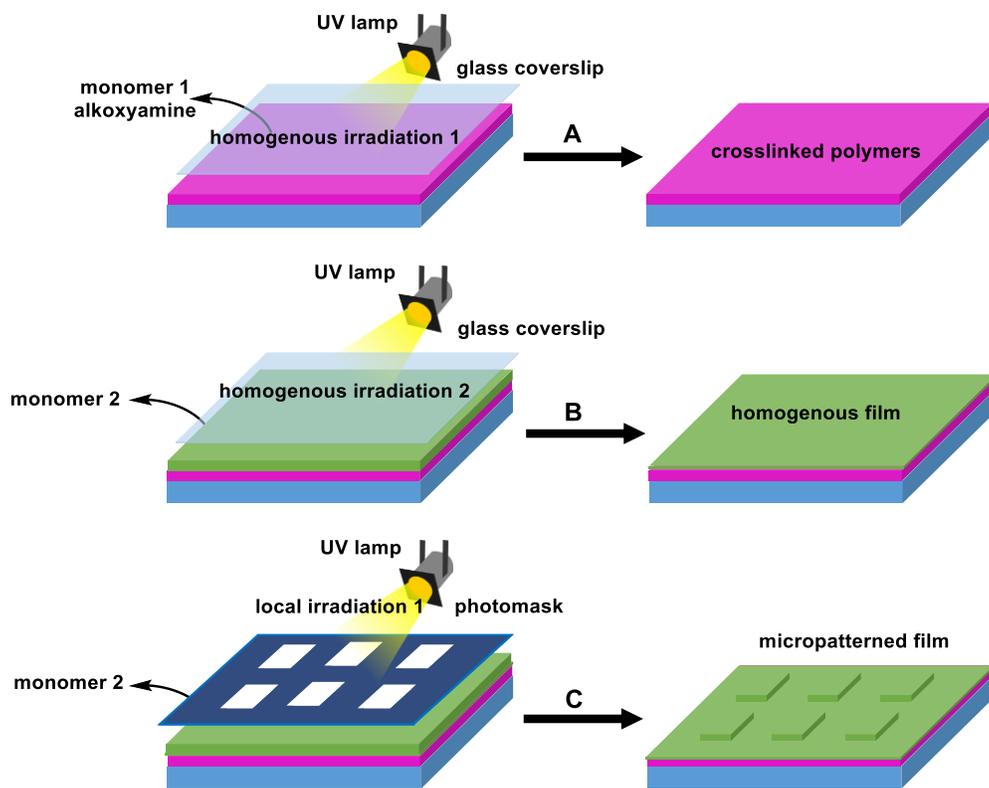
**Figure 39.** (A) Optical micrograph and (B) AMF image of nanoscale-inclined plane formed from polymer brushes upon light exposure through a shadow mask with optical density gradient. (C) Height along dashed line across feature as shown in (A). (Reproduced from Ref. [317], with permission from Wiley.)



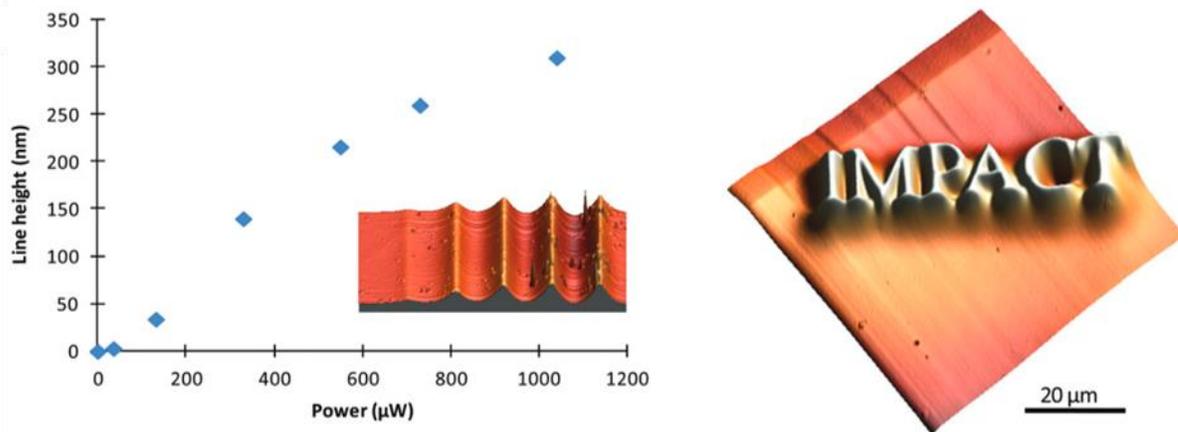
**Figure 40.** (l.h.s.) Fluorescence micrograph of photoiniferter grafting by UV irradiation through the stripe-patterned projection metal mask and the neutral-density filter and subsequently stained with rose bengal, and the three-dimensional image (b) of the distribution of the fluorescence intensity in the area shown in (a). (Reproduced from Ref. [323], with permission from the American Chemical Society.) (r.h.s) Schematic drawing of first generation to nth generation graft architectures. (Reproduced from Ref. [330], with permission from the American Chemical Society.)



**Figure 41.** Synthetic route to conductive, iodine-doped polythiophene brushes. (Reproduced from Ref. [347], with permission from The Royal Society of Chemistry.)

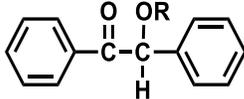
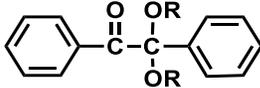
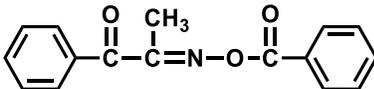
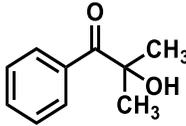
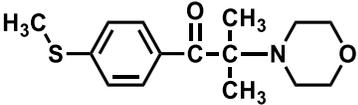
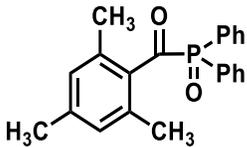
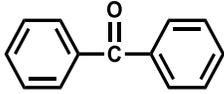
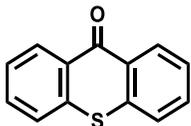


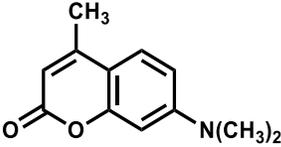
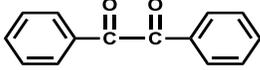
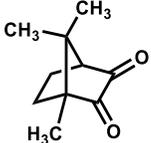
**Figure 42.** Schematic representation of fabrication micropatterned films via photomediated NMP process by combination homogeneous and local irradiations.



**Figure 43.** Direct laser writing with Ir-mediated photoATRP using a Nd:Yag microlaser (523 nm, 0.6 ns) with a  $\times 40$  microscope objective. (l.h.s.) Control of line height with laser power of 5 ms irradiation. (r.h.s.) Logo created with 130  $\mu\text{W}$ , 10 ms. (Reproduced from Ref. [126], with permission from The Royal Society of Chemistry.)

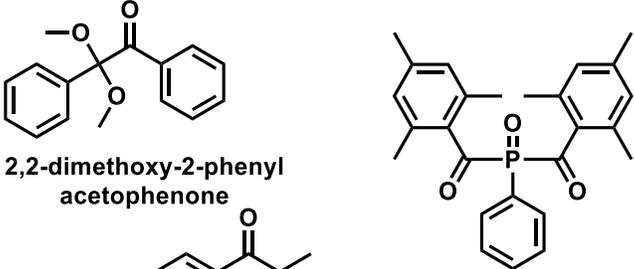
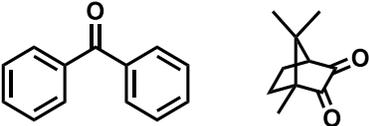
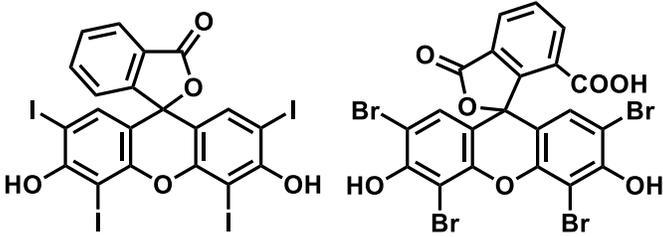
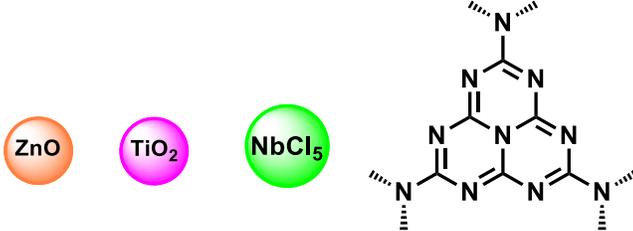
**Table 1.** Structure of some commercially available free radical photoinitiators.

Compounds	Type	Structure	$\lambda_{\text{max}}$ (nm)
Benzoin and benzoin ethers	Unimolecular		323
Benzil ketals	Unimolecular		365
Oxime esters	Unimolecular		335
Acetophenones	Unimolecular		340
Aminoalkyl phenones	Unimolecular		320
Acylphosphine oxides	Unimolecular		380
Benzophenones	Bimolecular		335
Thioxanthenes	Bimolecular		390

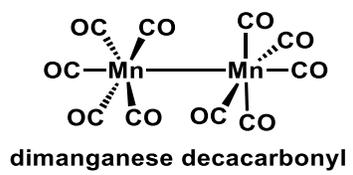
Coumarins	Bimolecular		370
Benzils	Bimolecular		340
Camphorquinones	Bimolecular		470

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**Table 2.** UV and visible light sensitive compounds used in photoinduced ATRP.

Types of photoactivators	Structures	References
Unimolecular photoinitiators	 <p data-bbox="475 558 760 615">2,2-dimethoxy-2-phenyl acetophenone</p> <p data-bbox="833 680 1154 737">bis (2,4,6-trimethylbenzoyl) phenylphosphine oxide</p> <p data-bbox="605 762 769 789">Irgacure 2959</p>	[50-53]
Bimolecular photoinitiators	 <p data-bbox="610 999 784 1024">benzophenone</p> <p data-bbox="821 999 1024 1024">camphorquinone</p>	[50, 54]
Dyes	 <p data-bbox="561 1329 719 1356">erythrosin B</p> <p data-bbox="946 1329 1039 1356">eosin Y</p>	[52]
Heterogeneous catalysts	 <p data-bbox="500 1644 565 1696">zinc oxide</p> <p data-bbox="602 1644 699 1696">titanium dioxide</p> <p data-bbox="740 1644 865 1696">niobium(V) chloride</p> <p data-bbox="911 1644 1089 1696">mesoporous carbon nitride</p>	[55-60]

Metal carbonyls



[61]

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**Table 3.** Examples for iron-mediated photoATRP of various methacrylates.

Polymer	Initiator	Catalyst system	Irradiation	$M_n$	$M_w/M_n$	[ref]
PMMA	CCl <sub>4</sub>	FeCl <sub>3</sub> , TMEDA, AIBN	Hg lamp (500 W)	17200	1.26	[95]
PMMA	CCl <sub>4</sub>	FeCl <sub>3</sub> , TMEDA, BPO	Hg lamp (500 W)	15800	1.25	[95]
PMMA	CCl <sub>4</sub>	FeCl <sub>3</sub> , bpy, ethanol	Hg lamp (500 W)	4200	1.22	[96]
PMMA	EBiB	FeCl <sub>3</sub> , PPh <sub>3</sub> , Fe <sub>2</sub> O <sub>3</sub>	Hg lamp (500 W)	11100	1.39	[97]
PMMA	EBiB	Fe(bpy) <sub>3</sub> (PF <sub>6</sub> ) <sub>2</sub>	LED (20 W, 75 mW/cm <sup>2</sup> )	20200	1.27	[98]
PGMA	EBiB	Fe(bpy) <sub>3</sub> (PF <sub>6</sub> ) <sub>2</sub>	LED (20 W, 75 mW/cm <sup>2</sup> )	22400	1.33	[98]
PMMA	EBPA	Fe complex, EDB	LED (405 nm)	41000	1.80	[99]
PMMA	EBPA	FeBr <sub>3</sub>	365 nm, 4.9 mW/cm <sup>2</sup>	4300	1.18	[100]
PBuMA	EBPA	FeBr <sub>3</sub>	365 nm, 4.9 mW/cm <sup>2</sup>	5800	1.19	[100]
PBnMA	EBPA	FeBr <sub>3</sub>	365 nm, 4.9 mW/cm <sup>2</sup>	12700	1.21	[100]
PMMA- <i>b</i> - PEMA	PMMA-Br	FeBr <sub>3</sub>	365 nm, 4.9 mW/cm <sup>2</sup>	9690	1.11	[100]

**Table 4.** Examples for Ir-mediated photoATRP of various methacrylates

Polymer	Initiator	Catalyst system	Irradiation	$M_n$	$M_w/M_n$	[ref]
PMMA	EBPA	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	435 nm LED (0.5 mW/cm <sup>2</sup> )	22900	1.25	[119]
PMAA	EBPA	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	435 nm LED (0.5 mW/cm <sup>2</sup> )	28000	1.61	[119]
PEGMA	EBiB	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	420 nm LED (4.5 mW/cm <sup>2</sup> )	6000- 19000	1.11- 1.26	[121]
PTFEMA	C <sub>6</sub> F <sub>13</sub> I	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	420 nm LED (3.2 mW/cm <sup>2</sup> )	22000	1.23	[122]
PPMA-b- PMMA	EBPA	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	420 nm LED (4.5 mW/cm <sup>2</sup> )	8300	1.24	[123]
PMA	BnBiB	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	380 nm LED (0.65 μ W/cm <sup>2</sup> )	8200	1.19	[124]
PEA-b-PAA	BnBiB	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	380 nm LED (0.65 μ W/cm <sup>2</sup> )	16300	1.65	[124]
PMMA	EBPA	Ir(btp) <sub>2</sub> (tmd)	462 nm LED (10 mW/cm <sup>2</sup> )	25140	1.2	[126]

**Table 5.** Examples for PI-initiated photoRAFT polymerizations of various monomers.

Polymer	PI	RAFT agent	$M_n$	$\bar{D}$	[ref]
PS	TPO	BDMAT	2 700	1.26	[191]
PS-b-PVBA	TPO	BDMAT	15 500	1.15	[191]
PNIPAAm	Irgacure 2959	BDMAT	13 500	1.17	[195]
PGMA	TPO	CPFDB	36 100	1.18	[193]
PBMA	TPO	CPFDB	6 800	1.20	[193]
PBMA-b-PGMA	TPO	CPFDB	27 800	1.17	[193]
PBA	TPO	BDMAT	80 000	1.3	[192]
PBA-b-PMA-b-PBA	TPO	BDMAT	-- <sup>a</sup>	-- <sup>a</sup>	[192]
PNAP	TPO	EDMAT	6 450	1.10	[194]
PNAP-b-PPEGA	TPO	EDMAT	-- <sup>a</sup>	-- <sup>a</sup>	[194]
PVAc	BAPO	MESA	10 500	1.23	[196]

<sup>a</sup>Good control over the block extension is seen, however, no  $M_n$  data was given in the original publication

**Table 6.** Examples for initiator-free photoRAFT polymerizations of various monomers.

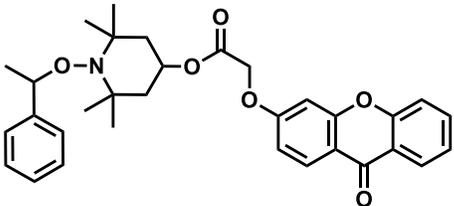
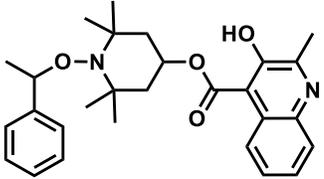
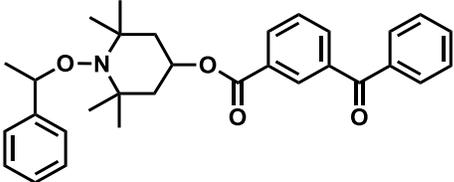
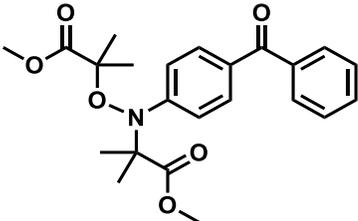
Polymer	RAFT agent	$M_n$	$\bar{D}$	[ref]
PAA	BDMAT	39 000	1.08	[201]
PBA	BDMAT	37 700	1.05	[203]
PS	BDMAT	15 600	1.12	[203]
PS- <i>b</i> -PBA- <i>b</i> -PS	BDMAT	210 000	1.15	[203]
PVAc	CPEC	16 900	1.38	[205]
PVAc	MESA	10 600	1.24	[205]
PMA	DDMAT	53 700	1.19	[206]
PMA	DDMAT	9 300	1.06	[208]
PNIPAAm	DDMAT	55 800	1.12	[208]
PMA <sup>a</sup>	DDMAT	~80 000	<1.15	[208]

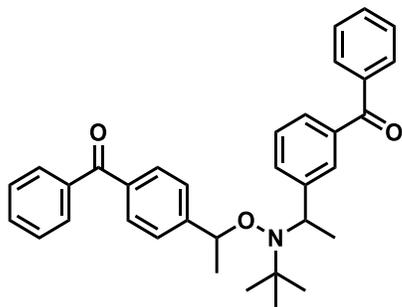
<sup>a</sup>Polymer was obtained by adding fresh monomer five times at >90% monomer conversion. The polymer can hence be defined as a pseudohexablock copolymer.

**Table 7.** Examples for PET-RAFT polymerizations of various monomers.

Polymer	PET catalyst	RAFT agent	$M_n$	$\bar{D}$	[ref]
PMA	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	BTPA	17 500	1.05	[214]
PMMA	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	CPADB	335 000	1.23	[214]
PMMA- <i>b</i> -PHPMA	<i>fac</i> -[Ir(ppy) <sub>3</sub> ]	BTPA	31 330	1.16	[214]
PDMA	Ru(bpy) <sub>3</sub>	BTPA	15 900	1.09	[218]
PMA	Ru(bpy) <sub>3</sub>	BTPA	28 900	1.11	[218]
PMMA	Eosin Y	CPADB	14 900	1.21	[220]
PMMA	Chlorophyll	CPADB	20 400	1.16	[222]
PHEMA	Chlorophyll	CPADB	22 700	1.08	[222]
PMA	porphyrin	BTPA	10 500	1.07	[221]
PNIPAAM	curcumin	Not given	43 500	1.12	[224]
PNIPAAM	PTZ	Symmetrical TTC	23 200	1.02	[227]
PNIPAAM	NbBA	CETP	45 500	1.18	[226]
PAA	NbBA	CETP	16 800	1.20	[226]
PPEGA	NbBA	CETP	89 000	1.22	[226]
PMMA	NbBA	CETP	18 200	1.16	[226]
PNVC	Mn <sub>2</sub> (CO) <sub>10</sub>	CPDN	2 600	1.22	[225]
PMA	Mn <sub>2</sub> (CO) <sub>10</sub>	CPDN	15 900	1.10	[225]
PDMAEMA	Mn <sub>2</sub> (CO) <sub>10</sub>	CPDN	19 600	1.37	[225]

**Table 8.** Examples of photosensitive alkoxyamines used in the NMP process.

Structures	Conditions	Results and References
	Methyl methacrylate, UV light (Rayonet reactor) 12 h	Molecular weight increase with monomer concentration [253]
	Styrene, UV light (300–400 nm) 175 W xenon lamp with cutoff filters 2h	10% conversion, $M_n = 2700$ and $M_w/M_n = 1.60$ [252]
	n-Butyl acrylate UV light (22 mW cm <sup>-2</sup> ) 500 s	30% conversion, $M_n = 250,000$ and $M_w/M_n = 2.50$ [172]
	n-Butyl acrylate UV light (22 mW cm <sup>-2</sup> ) 500 s	30-80% conversion, $M_n = 25,000-40,000$ and $M_w/M_n = 2.20-3.90$ [172]



n-Butyl acrylate, UV light

(44 mW cm<sup>-2</sup>)

1000 s

Isobornyl acrylate

UV light

Rayonet reactor with 16  
lamps @ 365 nm

5 min

45% conversion,

$M_n = 25,000-40,000$

and  $M_w/M_n = 2.20-3.90$

[255]

50,000-90,000

1.5-2.1 40% conversion

[258]

DMPA, MMA,

16-77% conversion

UV light

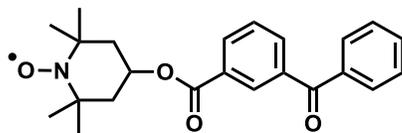
$M_n = 14,000-33,000$

(3 mW cm<sup>-2</sup>)

and  $M_w/M_n = 1.35-2.19$

45 min

[58]



DMPA, MMA,

14-79% conversion

UV light

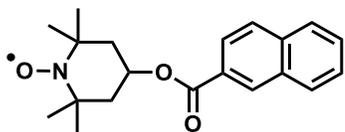
$M_n = 1,800-10,000$

(3 mW cm<sup>-2</sup>)

and  $M_w/M_n = 1.34-2.39$

45 min

[58]



DMPA, MMA,

15-58% conversion

UV light

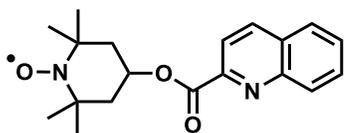
$M_n = 2,000-7,500$

(3 mW cm<sup>-2</sup>)

and  $M_w/M_n = 1.33-1.37$

45 min

[58]



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